

Part I. Write the chemical name of the following ionic compounds.

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| 1. RbBr | 2. RaI ₂ | 3. Cs ₂ Se |
| 4. AlI ₃ | 5. AgBr | 6. Fr ₃ N |
| 7. SrO | 8. Au ₂ S | 9. Zn ₃ P ₂ |
| 10. NaBr | 11. Co ₃ N ₂ | 12. FeP |
| 13. Cu ₂ S | 14. Fe ₂ O ₃ | 15. PbO ₂ |
| 16. NiCl ₂ | 17. CrN | 18. HgO |
| 19. ZrI ₂ | 20. MnO | 21. Ti ₃ N |
| 22. FeO | 23. SnBr ₂ | 24. Cu ₃ P ₂ |
| 25. SnS ₂ | | |

Part II. Write the chemical formulas for the following ionic compounds.

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| 26. Sodium Iodide | 27. Barium Bromide | 28. Titanium (III) Oxide |
| 29. Potassium Selenide | 30. Strontium Hydride | 31. Calcium Phosphide |
| 32. Rubidium Sulfide | 33. Beryllium fluoride | 34. Cobalt (II) Bromide |
| 35. Chromium (II) Telluride | 36. Iron (II) Bromide | 37. Nickel Oxide |
| 38. Zinc Nitride | 39. Mercury (I) Sulfide | 40. Titanium Phosphide |
| 41. Silver Hydride | 42. Mercury (II) Nitride | 43. Tin (IV) Oxide |
| 44. Scandium Phosphide | | |

Answer List

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|------------------------------------|----------------------------|------------------------------------|
| 1. Rubidium Bromide | 2. Radium Iodide | 3. Cesium Selenide |
| 4. Aluminum Iodide | 5. Silver Bromide | 6. Francium Nitride |
| 7. Strontium Oxide | 8. Gold Sulfide | 9. Zinc Phosphide |
| 10. Sodium Bromide | 11. Cobalt (II) Nitride | 12. Iron (III) Phosphide |
| 13. Copper (I) Sulfide | 14. Iron (III) Oxide | 15. Lead (IV) Oxide |
| 16. Nickel (II) Chloride | 17. Chromium (III) Nitride | 18. Mercury (II) Oxide |
| 19. Zirconium (II) Iodide | 20. Manganese (II) Oxide | 21. Titanium (I) Nitride |
| 22. Iron (II) Oxide | 23. Tin (II) Bromide | 24. Copper (II) Phosphide |
| 25. Tin (IV) Sulfide | 26. NaCl | 27. BaBr ₂ |
| 28. Ti ₂ O ₃ | 29. K ₂ Te | 30. SrH ₂ |
| 31. Ca ₃ P ₂ | 32. Rb ₂ S | 33. BeF ₂ |
| 34. CoBr ₂ | 35. CrTe | 36. FeBr ₂ |
| 37. NiO | 38. Ca | 39. Hg ₂ S |
| 40. Ti ₃ P ₂ | 41. AgH | 42. Hg ₃ N ₂ |
| 43. SnO ₂ | 44. ScP | |