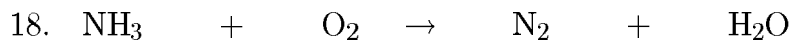
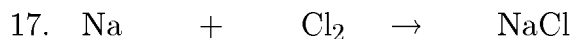
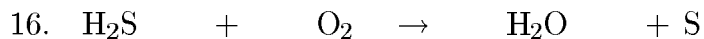
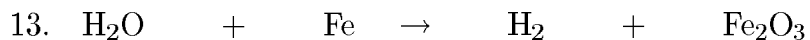
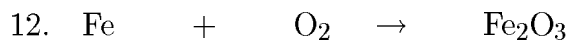
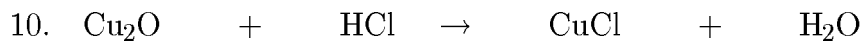
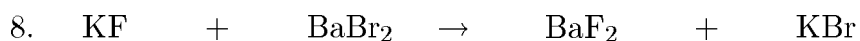
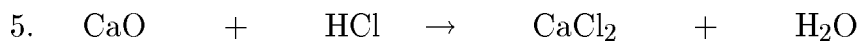
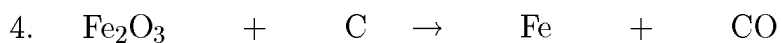
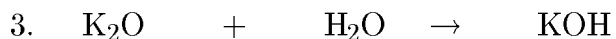
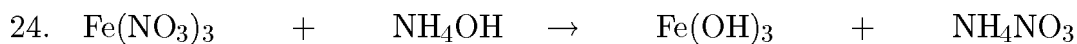
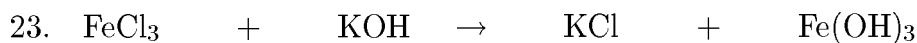
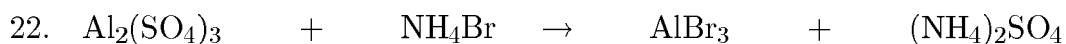
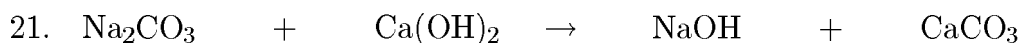


Balancing Equations Practice

Per/Sec. _____ Date _____

Part I. Balance the following chemical equations.





Part II. Write the skeleton equation and then balance the chemical reaction.

25. ammonium chloride + sodium nitrite \rightarrow sodium chloride + ammonium nitrite

26. ammonium sulfide + lead (II) nitrate \rightarrow ammonium nitrate + lead (II) sulfide

27. calcium + water \rightarrow calcium hydroxide + hydrogen

28. Chromic chloride and barium sulfate combine to form chromic sulfate and barium chloride.

29. aluminum sulfate + calcium hydroxide \rightarrow aluminum hydroxide + calcium sulfate

30. fluorine + sodium hydroxide \rightarrow sodium fluoride + oxygen + water

31. Scandium sulfate and sodium bicarbonate react to produce scandium hydroxide, sodium sulfate, and carbon dioxide.

Answer List

1. $2\text{Fe} + 3\text{S} \rightarrow \text{Fe}_2\text{S}_3$
2. $2\text{MgO} \rightarrow 2\text{Mg} + \text{O}_2$
3. $\text{K}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{KOH}$
4. $\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 2\text{Fe} + 3\text{CO}$
5. $\text{CaO} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O}$
6. $2\text{KI} + \text{Cl}_2 \rightarrow 2\text{KCl} + \text{I}_2$
7. $2\text{KI} + \text{H}_2\text{O}_2 \rightarrow 2\text{KOH} + \text{I}_2$
8. $2\text{KF} + \text{BaBr}_2 \rightarrow \text{BaF}_2 + 2\text{KBr}$
9. $\text{CO}_2 + \text{C} \rightarrow 2\text{CO}$
10. $\text{Cu}_2\text{O} + 2\text{HCl} \rightarrow 2\text{CuCl} + \text{H}_2\text{O}$
11. $2\text{W} + 3\text{Cl}_2 \rightarrow 2\text{WCl}_3$
12. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
13. $3\text{H}_2\text{O} + 4\text{Fe} \rightarrow 3\text{H}_2 + 2\text{Fe}_2\text{O}_3$
14. $\text{Al}_2\text{O}_3 + 3\text{C} + 3\text{Cl}_2 \rightarrow 3\text{CO} + 2\text{AlCl}_3$
15. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $2\text{H}_2\text{S} + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{S}$
17. $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
18. $4\text{NH}_3 + 3\text{O}_2 \rightarrow 2\text{N}_2 + 6\text{H}_2\text{O}$
19. $2\text{Hg}_2\text{O} + \text{O}_2 \rightarrow 4\text{HgO}$
20. $\text{CaSO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + \text{Na}_2\text{SO}_4$
21. $\text{Na}_2\text{CO}_3 + \text{Ca}(\text{OH})_2 \rightarrow 2\text{NaOH} + \text{CaCO}_3$
22. $\text{Al}_2(\text{SO}_4)_3 + 6\text{NH}_4\text{Br} \rightarrow 2\text{AlBr}_3 + 3(\text{NH}_4)_2\text{SO}_4$
23. $\text{FeCl}_3 + 3\text{KOH} \rightarrow 3\text{KCl} + \text{Fe}(\text{OH})_3$
24. $\text{Fe}(\text{NO}_3)_3 + 3\text{NH}_4\text{OH} \rightarrow \text{Fe}(\text{OH})_3 + 3\text{NH}_4\text{NO}_3$
25. $\text{NH}_4\text{Cl} + \text{NaNO}_2 \rightarrow \text{NaCl} + \text{NH}_4\text{NO}_2$
26. $(\text{NH}_4)_2\text{S} + \text{Pb}(\text{NO}_3)_2 \rightarrow 2\text{NH}_4\text{NO}_3 + \text{PbS}$
27. $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca}(\text{OH})_2 + \text{H}_2$
28. $2\text{CrCl}_3 + 3\text{BaSO}_4 \rightarrow \text{Cr}_2(\text{SO}_4)_3 + 3\text{BaCl}_3$
29. $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 2\text{Al}(\text{OH})_3 + 3\text{CaSO}_4$
30. $2\text{F}_2 + 4\text{NaOH} \rightarrow 4\text{NaF} + \text{O}_2 + 2\text{H}_2\text{O}$
31. $\text{Sc}_2(\text{SO}_4)_3 + 6\text{NaHCO}_3 \rightarrow 2\text{Sc}(\text{OH})_3 + 3\text{Na}_2\text{SO}_4 + 6\text{CO}_2$