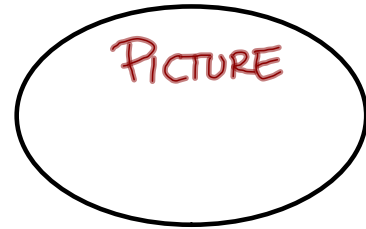


Diagramming

CONTENT



CAPTION

SUB-HEADING

CONTENT
(EQUATION)

SUB-HEADING

CONTENT

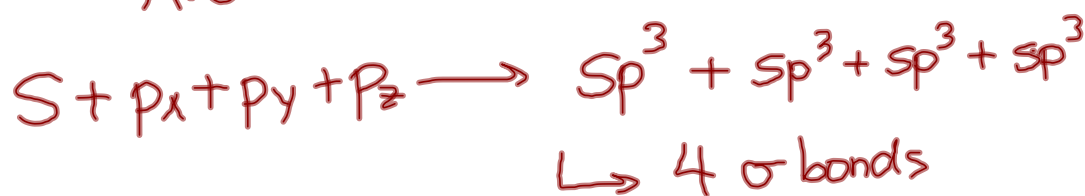
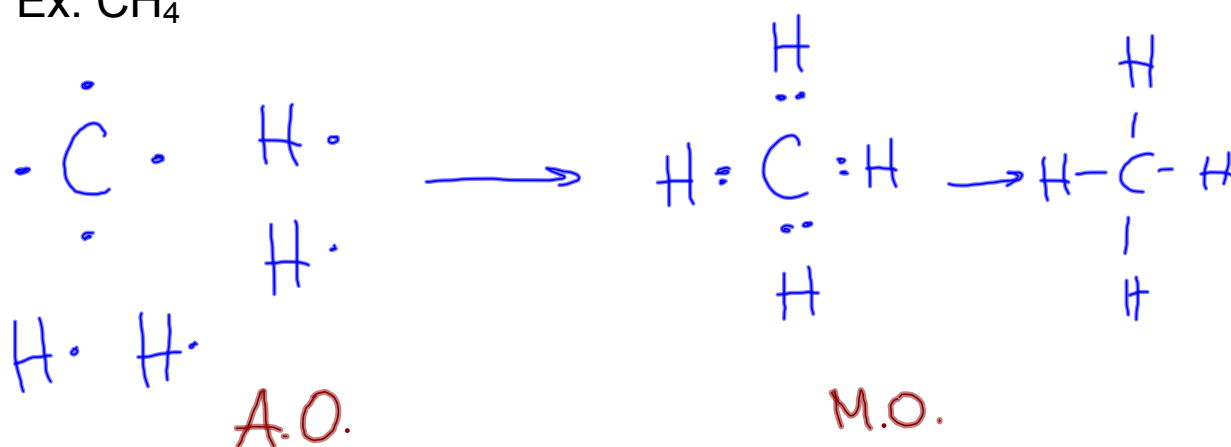
SUB-HEADING

CONTENT

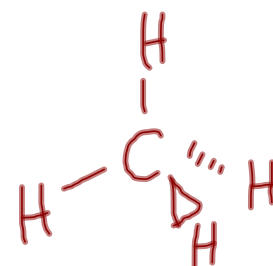
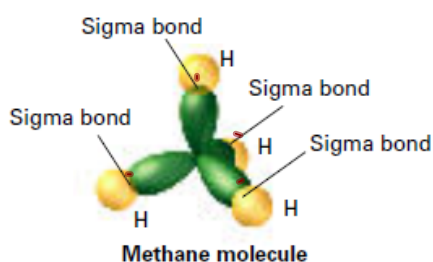
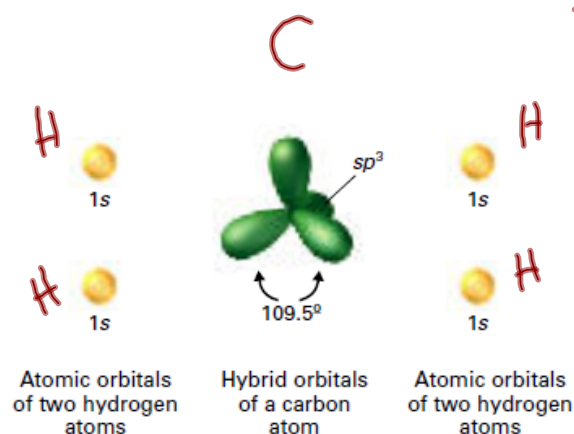
Hybridization Involving Single Bonds

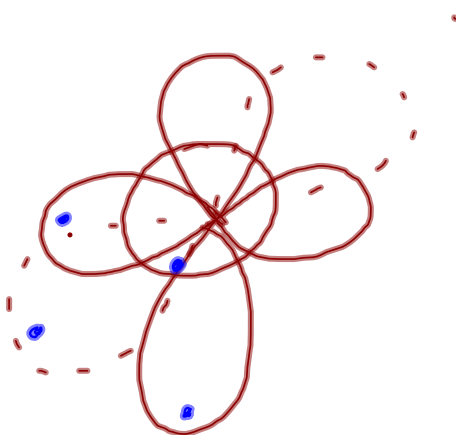
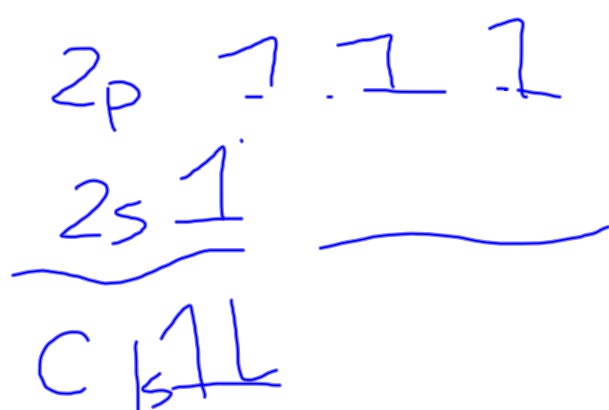
In hybridization, atomic orbitals mix to form the same total number of equivalent hybrid orbitals.

Ex. CH₄



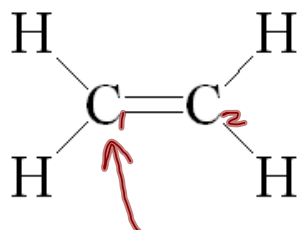
tetrahedral





Hybridization Involving Double Bonds

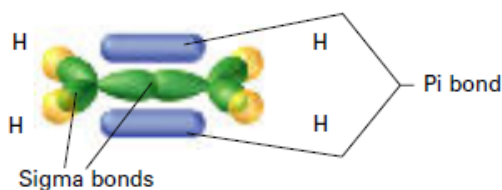
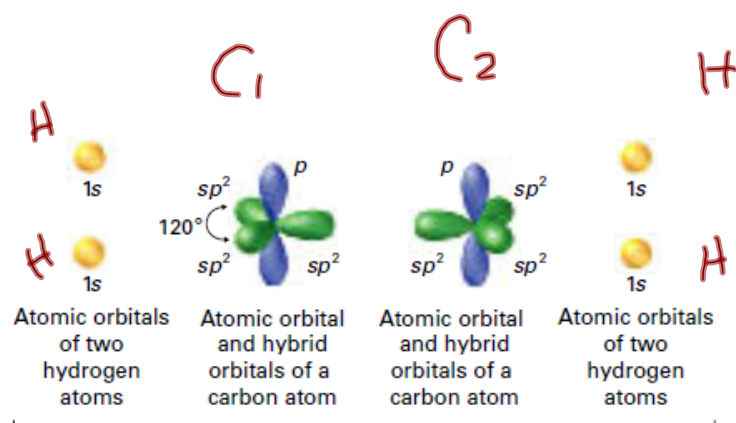
Ex. C_2H_4



$\hookrightarrow 3 \sigma$ bonds

$p_y \rightarrow p_y$ trigonal planar

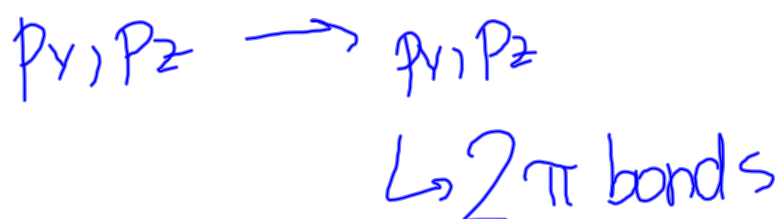
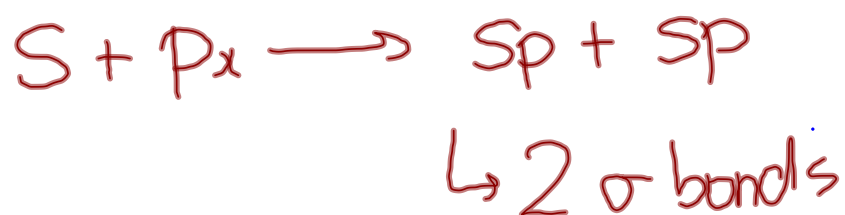
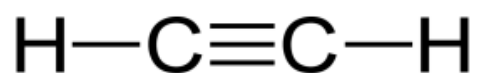
$\hookrightarrow \pi$ bond



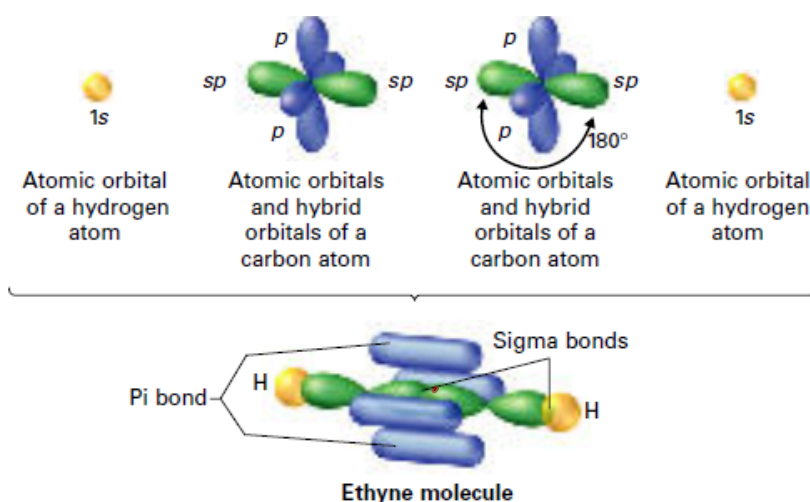
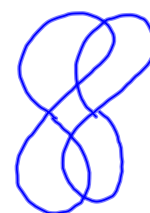
Ethene molecule

Hybridization Involving Triple Bonds

Ex. C_2H_2



LINEAR



Homework

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