Homework #1-5

Chemical Equilibrium

Reversible Reaction

Reaction in which both the forward and reverse processes are occuring simultaneously.

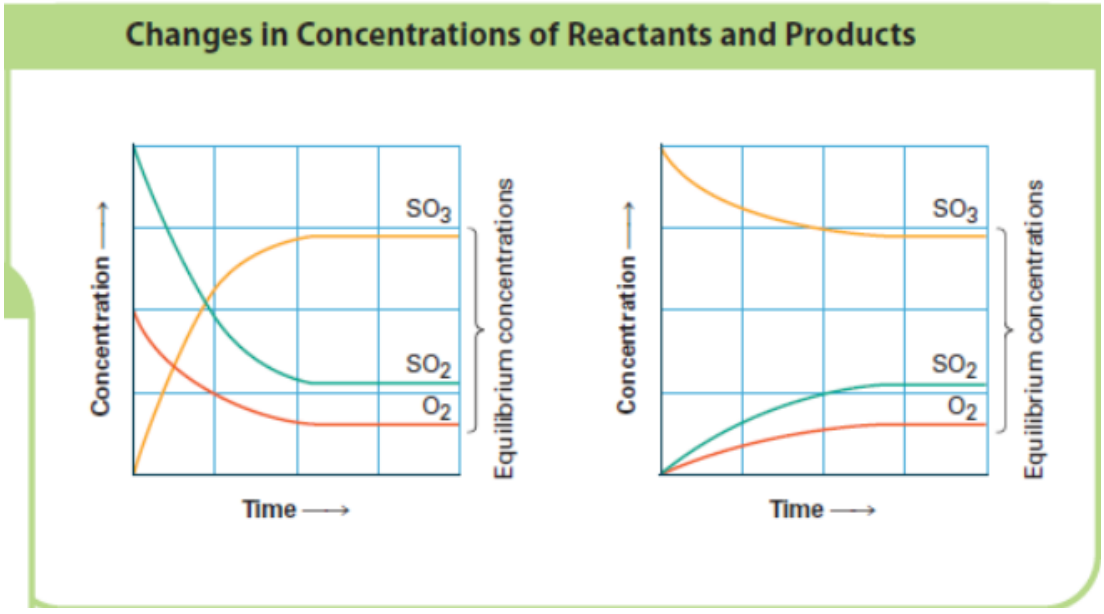
$$A + B \rightleftharpoons C$$

Chemical equilibrium

A system is said to have reached chemical equilibrium when the forward and reverse reactions are occuring at the same rate.

⟶ no net change occurs in the concentration of components of the system

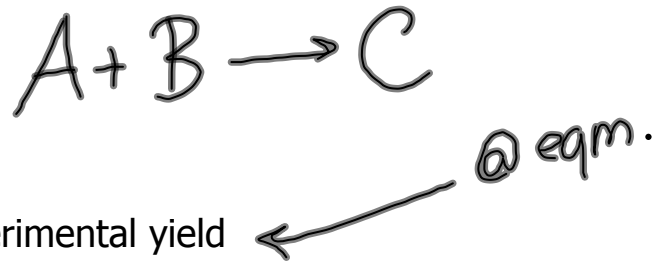
http://www.chm.davidson.edu/ronutt/che115/EquKin/EquKin.htm



Percent Reaction

Percent Reaction (percent yield) - is the amount of product measured at equilibrium compared with the maximum possible amount of product.

Equilibrium position
relative concentration of reactants and products at equilibrium
⇒0 % indicates no product formed
⇒100 % indicates the maximum possible product formed



$$\% \text{ reaction} = \frac{\text{Experimental yield}}{\text{Theoretical yield (maximum)}} \times 100\ \%$$

Classifying Chemical Equilibria

< 50 % - reactants favored
> 50 % - products favored
> 99 % - quantitative

The equilibrium position of the reaction is indicated in the following manner :

$$\text{Ex.} H_{2(g)} + I_{2(g)} \overset{78\%}{<==>} 2HI_{(g)} \qquad t = 448^{o}$$

Indicates that 78 % of the total amount of HI possible is produced at 448°C. Therefore this is a **product** favored reaction.

Limiting Reagent

In a chemical reaction, the reactant that will "run out" first is called the **limiting reagent**.
The other reactant is called the **excess reagent**.

$$2NH_3 + 3CuO \longrightarrow N_2 + 3Cu + 3H_2O$$

18.1 g **90.4 g**

If NH_3 is L.R.

$$18.1\,g\,NH_3 \times \frac{1\,mol\,NH_3}{17.04\,g\,NH_3} \times \frac{1\,mol\,N_2}{2\,mol\,NH_3} \times \frac{28.02\,g\,N_2}{1\,mol\,N_2} =$$

$$14.882\,g\,N_2$$

If CuO is L.R.

$$90.4\,g\,CuO \times \frac{1\,mol\,CuO}{79.54\,g\,CuO} \times \frac{1\,mol\,N_2}{3\,mol\,CuO} \times \frac{28.02\,g\,N_2}{1\,mol\,N_2} =$$

$$10.615\,g\,N_2$$

∴ CuO is L.R.

SAMPLE PROBLEM : % REACTION

Find the % reaction and write the expression if 6.90 g of $H_{2(g)}$ and 32.0 g of $O_{2(g)}$ react to form 25.75 g of ice at - 70 C°.

$$2H_{2(g)} + O_{2(g)} \rightleftharpoons 2H_2O_{(s)}$$

init.	6.90 g	32.0 g	—
eqm.			25.75 g

$$\% \text{ rxn} = \frac{\text{exp.}}{\text{theor.}} \times 100\%$$