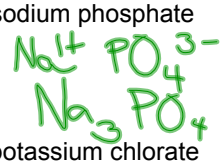
Tuesday Dec 6

1) answers pg 198 #3,4
2) oxyacids
3) Practice naming/formulas for polyatomic compounds

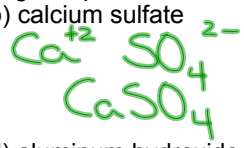
Answers pg 198

3. Write the formulas for the following compounds

a) sodium phosphate

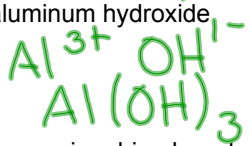


b) calcium sulfate

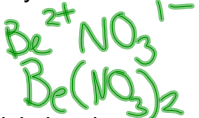


c) potassium chlorate

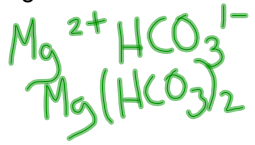
d) aluminum hydroxide



e) beryllium nitrate



f) magnesium bicarbonate



g) nickel carbonate

4. Write the names of the following compounds:

a) K_2CO_3 potassium carbonate

b) Na_2SO4 sodium sulphate

c) $Al(HCO_3)_3$ aluminum bicarbonate

d) $AgNO_3$ silver nitrate

<u>oxyacids</u> - compounds formed when hydrogen combines with polyatomic ions that contain oxygen.

Ex. H^+ NO_3^- (nitrate) Ex. H^{1+} SO_4^{2-}

HNO_3 H_2SO_4

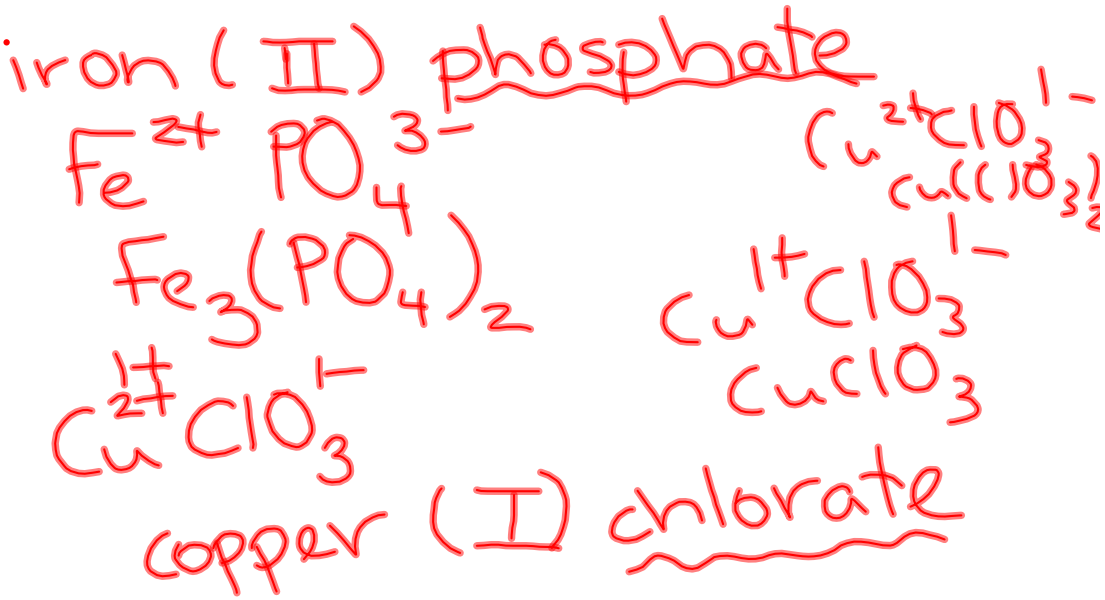
(nitric acid) (sulfuric acid)

common oxyacids on pg 198 table 3

H_3PO_4 phosphuric acid

Review Polyatomic Compounds

found on pg 196 table 2



Naming/Writing Formulas of Polyatomic Compounds
Practice WS

answers pg 198 #1-7.notebook