Homework - #17

Practice Problems

$$Zn_{(s)} + Pb(NO_3)_{2(aq)} \longrightarrow$$

$$F_{2(g)} + HCl_{(aq)}$$
 —

$$Al_{(s)} + CuSO_{4(aq)} \longrightarrow$$

Chemical Reactions

V. Double Replacement Reaction

Reaction that occurs between two ionic compounds in solution. Ions will "change partners".

⇒if one of the products has low solubility, it may form a precipitate (solid). This double replacement reaction is called **precipitation**.

$$CaCl_{2(aq)} + Na_{2}CO_{3(aq)} \longrightarrow CaCO_{3G} + 2NaCl_{M}$$

$$compound + compound \longrightarrow compound + compound$$

A second type of double replacement reaction is a **neutralization** reaction, which is a reaction between an acid and a base, to form water and an ionic compound.

Practice Problems

BaCl_{2(aq)} + Na₂SO_{4(aq)}
$$\rightarrow$$
 BaSO_{4(s)} + 2NaCl_{aq)}

3NaOH_(aq) + FeBr_{3(aq)} \rightarrow 3NaBr_{bq)} + Fe(OH)_{3(s)}

2KI_(aq) + Pb(NO₃)_{2(aq)} \rightarrow KNO_{3|q|} + PbI_{2(s)}

polassium | lead (II) | rotussium | rotussium | rotuse | rotussium | rotuse | rotus

Worksheet

p. 335 #18,19p. 339 #22-27