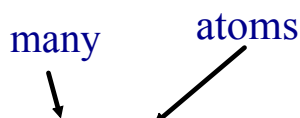


Monday Dec 5

- 1) Ionic Compounds Quiz
- 2) Notes on Polyatomic Compounds

Polyatomic Compounds

many atoms



Polyatomic Compounds

- **polyatomic ions** are groups of atoms that stay together and carry single charge

Ex. NO_3^- (nitrate ion) the whole ion has a charge of -1

PO_4^{-3} (phosphate) the whole ion has a charge of -3

- they are found on Table 2, pg. 196 (they will not be on the periodic table)

- you need to recognize polyatomic ions when they are in formulas so look over the names in the chart and become familiar with them (no need to memorize)

Writing Formulas for Polyatomic Compounds

Find the symbol for the first ion on the periodic table

Write the symbol with the charge above it

Find the symbol for the second ion in the chart on pg 196

Write the symbol with the charge above it

Cross the charges the same as a regular ionic compound (if you need to put a subscript with the polyatomic ion put it in brackets first)

Ex. sodium hydroxide



Ex. calcium sulphate



Ex. magnesium phosphate



Naming Polyatomic Ions

Combination of the name of the metal, and the name of the polyatomic ion.

Remember a compound has ONLY two names no more!!!!

Ex. K_2CO_3

1. Determine which ions make compound:



2. Name each ion:

pottasium carbonate

Ex. CaCO_3

calcium carbonate

Ex. Cu_3PO_4

copper (I) phosphate

p.198 #3,4

Attachments

answers pg 198 #1-7.notebook