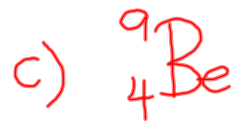
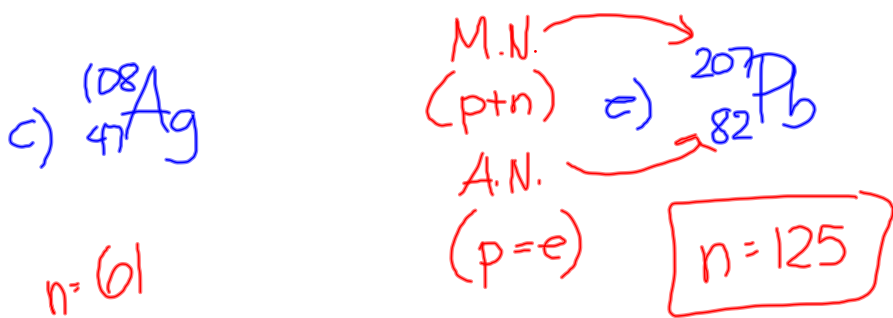
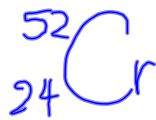


Homework #17-20



$4p^+$
 $4e^-$
 $5n$



$26n$

$28n$

$29n$

Isotope	protons	neutrons	electrons
copper - 64			
gold - 108			
sulfur - 33			

Sample Problem

Element X has two natural isotopes. The isotope with a mass of 10.012 amu (^{10}X) has a relative abundance of 19.91%. The isotope with a mass of 11.009 amu (^{11}X) has a relative abundance of 80.09%. Calculate the atomic mass of this element.

$$^{10}\text{X} \rightarrow 10.012 \text{ amu } (19.91\%)$$

$$^{11}\text{X} \rightarrow 11.009 \text{ amu } (80.09\%)$$

$$(10.012 \text{ amu})(0.1991) + (11.009 \text{ amu})(0.8009)$$

$$= \boxed{10.81}$$

$$\text{Course: } 27 (70\%)$$

$$\text{Exam: } 96 (30\%)$$

$$27(0.70) + 96(0.3)$$

$$= \boxed{47.7\%}$$

12.48

- ① ¹²C
- ② ¹³C
- ③ ¹⁴C

24.01

12.01

Homework

Isotope worksheet

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Homework - Isotopes Worksheet

Isotope Name	Atomic Number	Mass Number	Symbol	# of Protons	# of Neutrons
carbon - 14	6	14	$^{14}_6\text{C}$	6	8
hydrogen - 2					
lawrencium - 257					