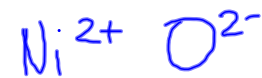
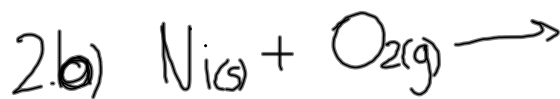
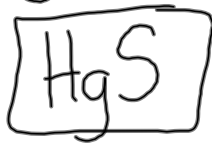
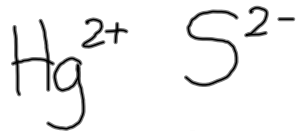
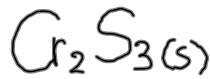
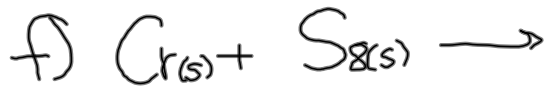
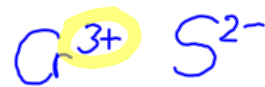


# Homework - Worksheet

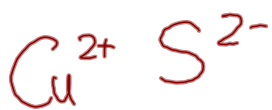
1. a) mercury(II) sulfide



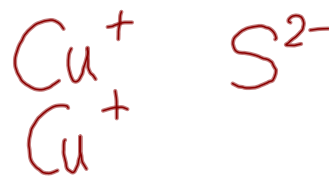
nickel(II) oxide



chromium(III) sulfide



copper(II) sulfide



copper(I) sulfide

# Writing Formulas for Polyatomic Ionic Compounds

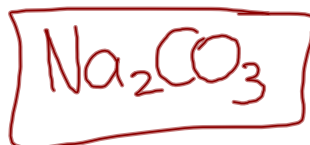
polyatomic ions - are ions made up of more than one atom.

Ex.  $\text{CO}_3^{-2}$  carbonate ion

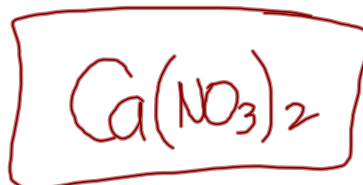
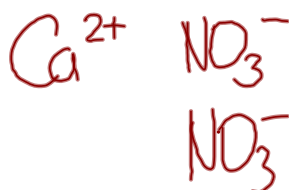
All polyatomic ions can be found with their charges on the back cover of the book.

Determining formulas of polyatomic ion compounds are done in the same manner as binary compounds.

Ex. sodium carbonate



Ex. calcium nitrate



## Naming Polyatomic Ionic Compounds

Ex.  $\text{CaCO}_3$

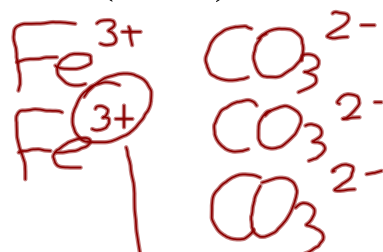
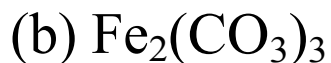


Calcium carbonate

Name the following polyatomic compounds:



Sodium bicarbonate



iron(III) carbonate

Write the chemical formula for the following polyatomic compounds:



**Write the chemical formula for the following polyatomic compounds:**

(c) calcium hydroxide

(d) lithium phosphate

# Sheet - Ionic Compounds

I	II										III	IV	V	VI	VII	VIII	
H																He	
Li	Be										B	C	N	O	F	Ne	
Na	Mg										Al	Si	P	S	Cl	Ar	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Uun	Uuu	Uub						
			La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

