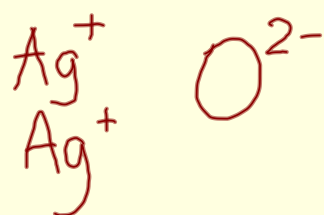
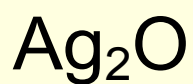




Quiz - Tomorrow

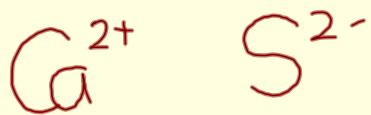
Ionic Compounds Review

Type I Binary Ionic Compounds



silver oxide

calcium sulfide



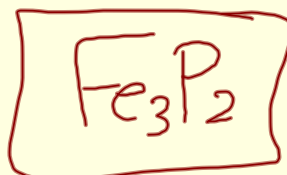
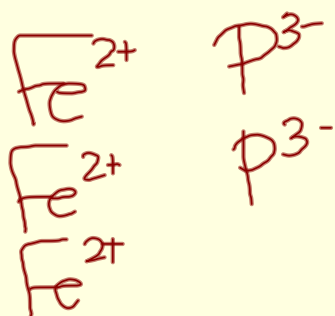
Ionic Compounds Review

Type II Binary Ionic Compounds



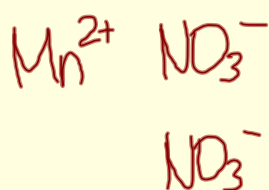
Copper(II) chloride

iron (II) phosphide



Ionic Compounds Review

Type III Ionic Compounds



manganese (II) nitrate

zinc oxalate



Molecular Compounds

MOLECULAR THEORY - **nonmetal atoms share electrons** in a **covalent bond** to attain a maximum number of valence electrons (complete outer shell) rather than gaining electrons from metal atoms.

Ex. CO₂

Molecular elements - although the chemical formula of metals are frequently shown alone as a single atom (Na), nonmetals frequently form **diatomic molecules**.

Ex. H₂, N₂, O₂, F₂, Cl₂, Br₂, I₂ *

Also: O₃, P₄, S₈

Naming *Binary molecular compounds*

As outlined by IUPAC rules, some molecular compounds signify the number of atoms in the molecular formula by using the same prefixes as hydrates.

Ex. CS₂

carbon disulfide

N₂O₅

dinitrogen pentoxide

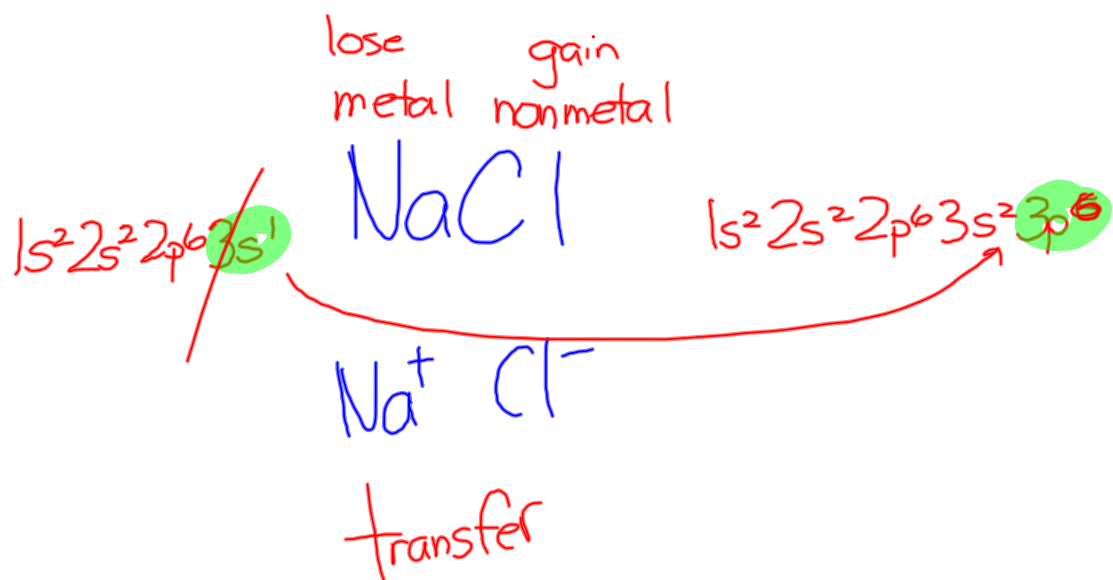
see Table 9.4 p. 269

The prefix system is usually not used for hydrogen molecular compounds

Ex. water - H₂O

Br HONClIF

Br₂ H₂ O₂ N₂ Cl₂ I₂ F₂



Worksheet