

## Unit 2 - Compounds

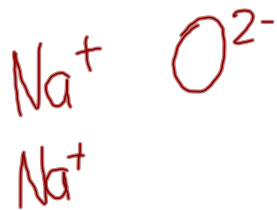
- ✓ Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- ✓ Naming Ionic Compounds
- ✓ Writing formulas for Ionic Compounds
- ✓ Ionic hydrates
- ✓ Naming Molecular Compounds
- ✓ Writing formulas for Molecular Compounds
  - Molecular Elements *Br HONClIF*
  - Drawing structural diagrams
- ✓ Naming and writing formulas for Acids and Bases
- ✓ Lab - Identifying Unknown Compounds

	Theoretical	Litmus	Conductivity
Ionic	Metal / Nonmetal + -	no change	✓
Molecular	2 or more nonmetals (share e <sup>-</sup> )	no change	✗
Acid	aqueous hydrogen compd. (H- or -COOH)	blue → red	✓
Base	ionic hydroxides CH <sub>3</sub> OH <del>CH<sub>3</sub></del> OH <sup>-</sup>	red → blue	✓

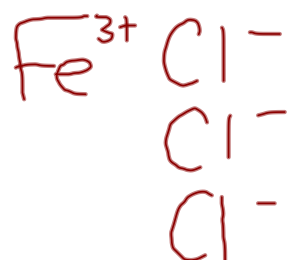


Calcium chloride

sodium oxide



$\text{Na}_2\text{O}$



iron (III) chloride

copper (II) sulfide

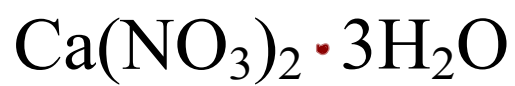




lithium nitrate

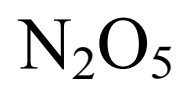
sodium sulfate





Calcium nitrate - 3-water

Calcium nitrate trihydrate



dinitrogen pentoxide

HClO

$H^+$   $ClO^-$  aqueous hydrogen hypochlorite  
hypochlorous acid

chromic acid  
chromate

$H^+$   $CrO_4^{2-}$   
 $H^+$



$CH_3COOH$

$H^+$   $CH_3COO^-$



**Review Questions p. 281-282**

**# 43-61, 65-71**