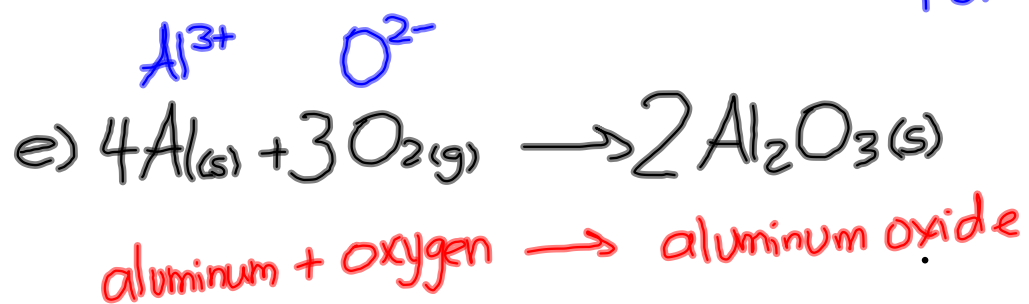
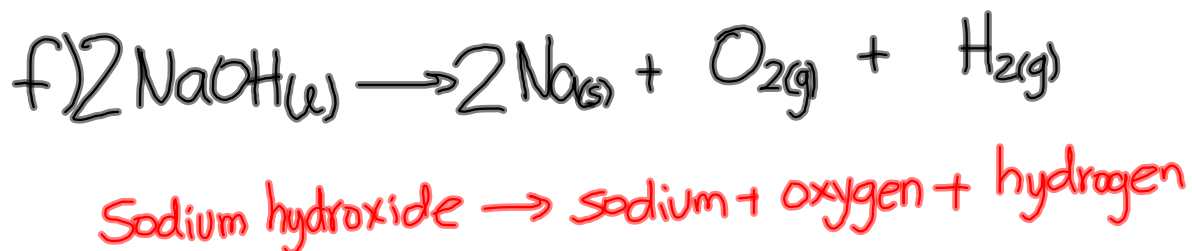


## Check Homework - Worksheet

FORMATION



DECOMP.



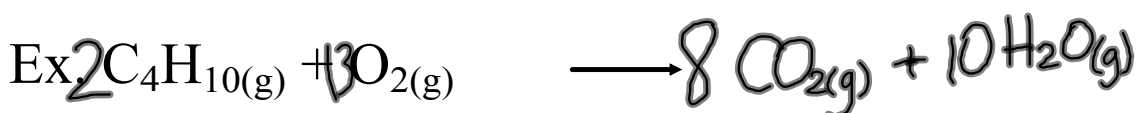
# Chemical Reactions

## III. Combustion Reaction

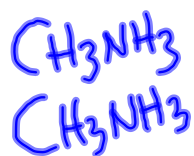
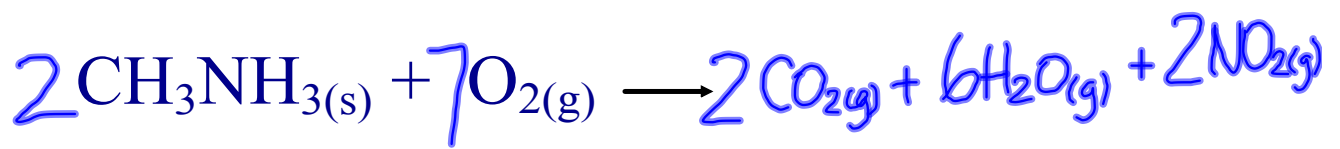
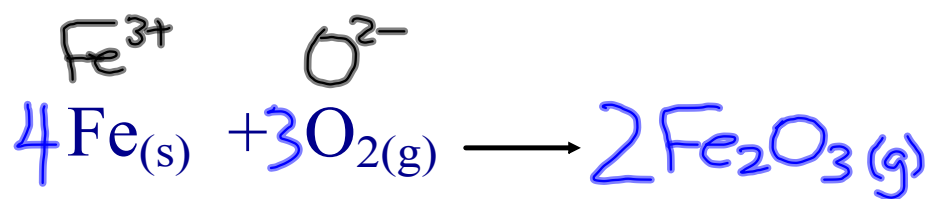
A complete combustion reaction is the burning of a substance with oxygen to produce the most common oxides of the elements in the substance being burned.

### Most Common Oxides:

- Carbon :  $\text{CO}_{2(g)}$
- Hydrogen:  $\text{H}_2\text{O}_{(g)}$
- Sulfur:  $\text{SO}_{2(g)}$
- Nitrogen:  $\text{NO}_{2(g)}$
- A metal: Oxide of metal with most common ion charge

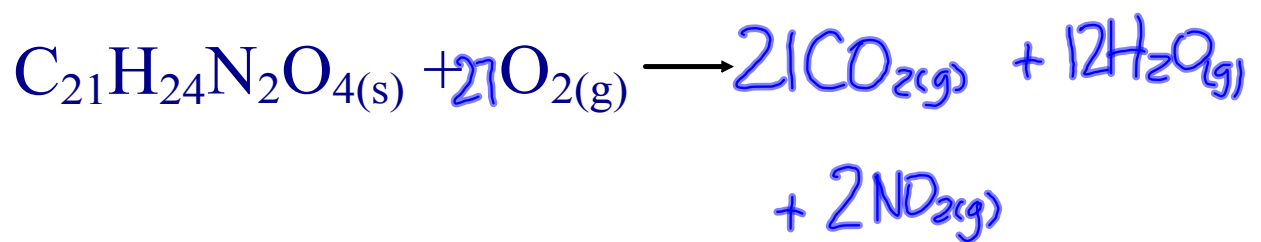
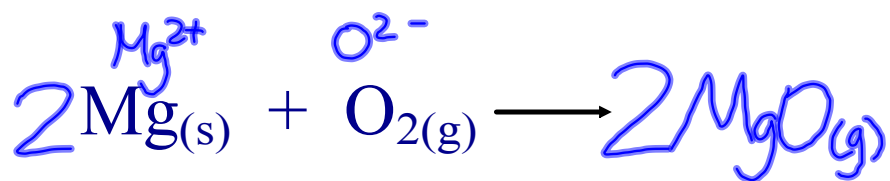


element/compound +  $\text{O}_2 \longrightarrow$  most common oxides



## Combustion Reactions

Write a balanced chemical equation for the following combustion reactions:



# Homework

**p. 331 #13, 14**

**p. 332 #15, 16**

**p. 337 #20, 21**