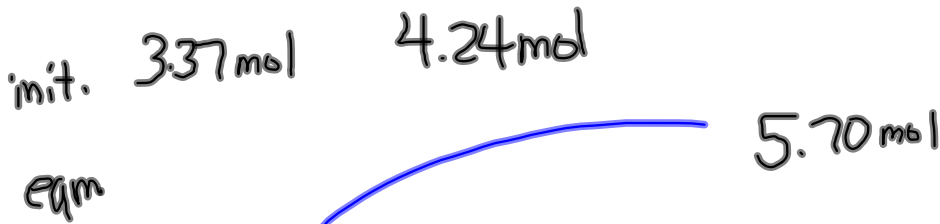


Homework - Worksheet



$$\% \text{ rxn} = \frac{\text{exp.}}{\text{theor.}} \times 100\%$$

$$\% \text{ rxn} = \frac{5.70 \text{ mol}}{6.74 \text{ mol}} \times 100\%$$

$$\% \text{ rxn} = 84.6\%$$

If H₂ is L.R.

$$3.37 \text{ mol H}_2 \times \frac{2 \text{ mol HI}}{1 \text{ mol H}_2}$$

$$= 6.74 \text{ mol HI}$$

If I₂ is L.R.

$$4.24 \text{ mol I}_2 \times \frac{2 \text{ mol HI}}{1 \text{ mol I}_2} = 8.48 \text{ mol HI}$$

∴ H₂ is L.R.

Equilibrium Law

Equilibrium Constant (K) - analysis of the experiments reveals a mathematical relationship that provides a constant value for an chem system over a range of concentrations.

This constant value is called the equilibrium constant (K) for the reac system

EQUILIBRIUM LAW

The Equilibrium Law is applied to calculate K_{eq} :

For the reaction: $aA + bB \rightleftharpoons cC + dD$

$$K_{eq} = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

A,B,C,D are chemical entities and a,b,c,d are their coefficients in the balanced equation.

The greater the value of the equilibrium constant, the more the system will favor the forward reaction.

The greater the value of K_{eq} , the greater the percent reaction.

$K_{eq} > 1$ is a product-favoured reaction

$K_{eq} < 1$ is a reactant-favoured reaction

Sample Problem

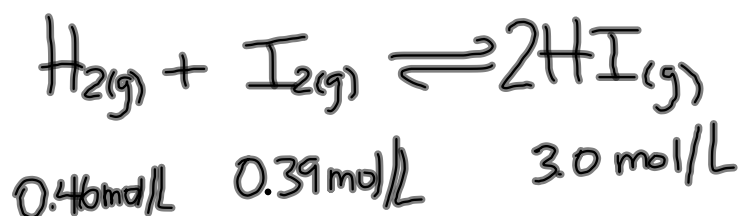
Write the equilibrium law for the reaction of nitrogen monoxide gas with oxygen gas to form nitrogen dioxide gas



$$K = \frac{[NO_{2(g)}]^2}{[NO_{(g)}]^2[O_{2(g)}]}$$

Sample Problem

A mixture of H_2 and I_2 is allowed to react at 448°C . When the equilibrium is established the concentrations of the participants are found to be $[\text{H}_2] = 0.46 \text{ mol/L}$, $[\text{I}_2] = 0.39 \text{ mol/L}$, and $[\text{HI}] = 3.0 \text{ mol/L}$. Calculate the value of K_{eq} at 448°C from these data.



$$K = \frac{[\text{HI}_{(g)}]^2}{[\text{H}_{2(g)}][\text{I}_{2(g)}]}$$

$$K = \frac{[3.0]^2}{[0.46][0.39]}$$

$$K = 50.$$

Product-favored

Worksheet

