

Homework - #17

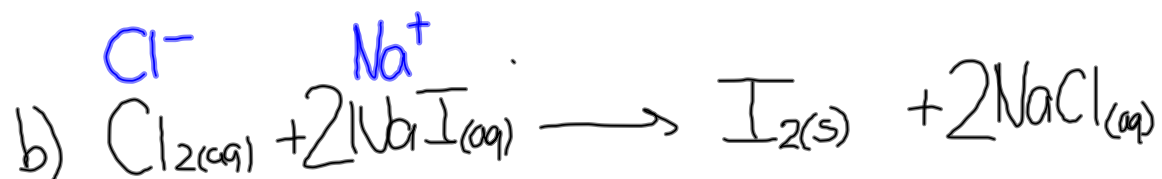
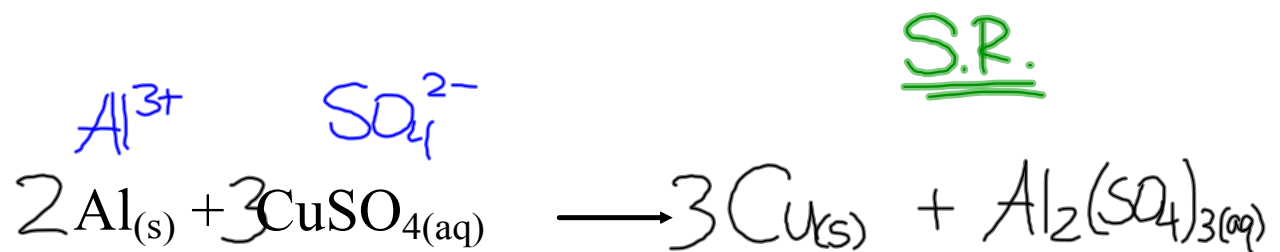


Table 11.2**Activity Series of Metals**

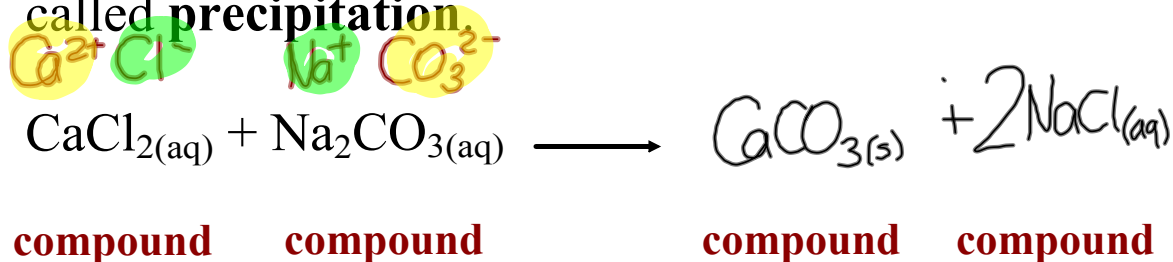
	Name	Symbol
Decreasing reactivity ↓	Lithium	Li
	Potassium	K
	Calcium	Ca
	Sodium	Na
	Magnesium	Mg
	Aluminum	Al
	Zinc	Zn
	Iron	Fe
	Lead	Pb
	(Hydrogen)	(H) ⁺
	Copper	Cu
	Mercury	Hg
	Silver	Ag

Chemical Reactions

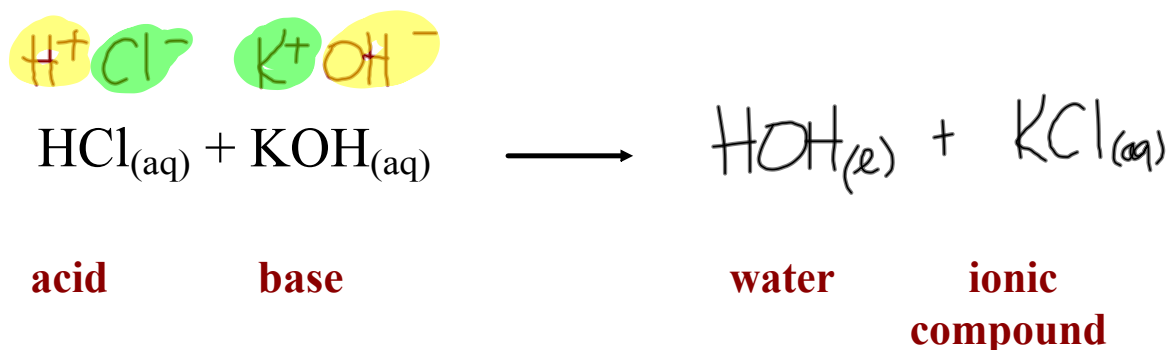
V. Double Replacement Reaction

Reaction that occurs between two ionic compounds in solution. Ions will "change partners".

⇒ if one of the products has low solubility, it may form a precipitate (solid). This double replacement reaction is called **precipitation**.

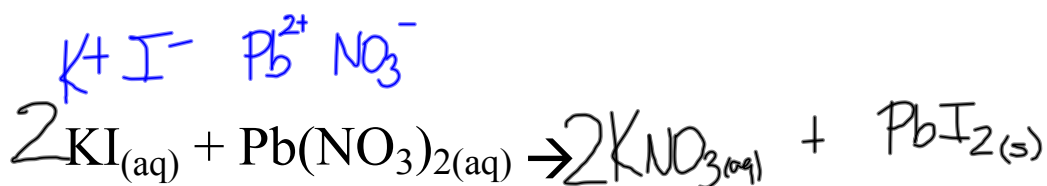
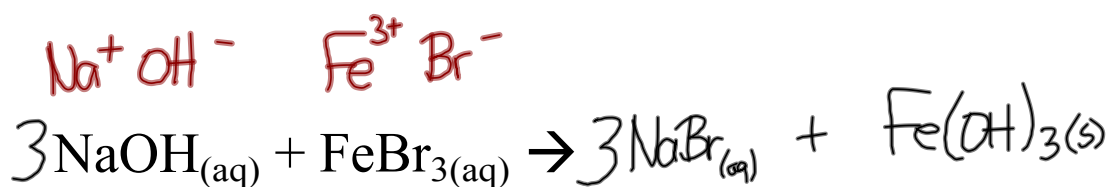
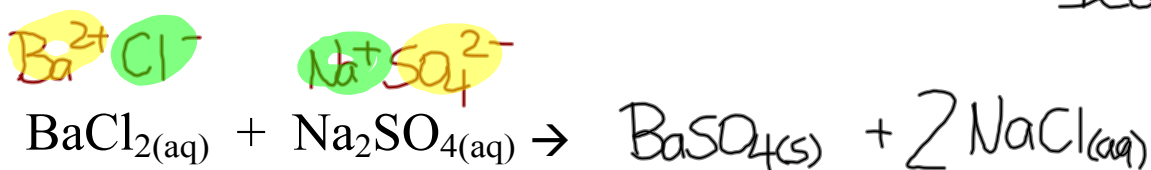


A second type of double replacement reaction is a **neutralization** reaction, which is a reaction between an acid and a base, to form water and an ionic compound.



Practice Problems

~~Fe(OH)₃~~



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