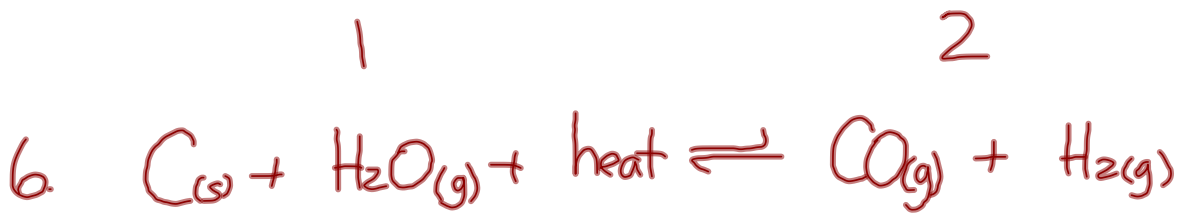


Check Homework



a) lower temp. \leftarrow

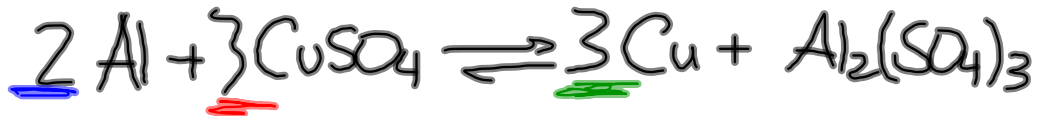
b) inc. pressure \leftarrow

c) remove H_2 \rightarrow

d) add $\text{H}_2\text{O(g)}$ \rightarrow



$$K = \frac{[\text{NH}_3\text{(g)}]^2}{[\text{N}_2\text{(g)}][\text{H}_2\text{(g)}]^3}$$



init 1.87g 9.65g

eqm.

3.65g

$$\% \text{rxn} = \frac{\text{exp.}}{\text{theor.}} \times 100\%$$

$$\% \text{rxn} = \frac{3.65\text{g}}{3.84\text{g}} \times 100\%$$

$$= 95.1\%$$

If Al is L.R.

$$1.87\text{g Al} \times \frac{1 \text{ mol Al}}{26.98\text{g Al}} \times \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \times \frac{63.54\text{g Cu}}{1 \text{ mol Cu}} = 6.606\text{g Cu}$$

If CuSO₄ is L.R.

$$9.65\text{g CuSO}_4 \times \frac{1 \text{ mol CuSO}_4}{159.60\text{g CuSO}_4} \times \frac{3 \text{ mol Cu}}{3 \text{ mol CuSO}_4} \times \frac{63.54\text{g Cu}}{1 \text{ mol Cu}} =$$

3.842g Cu

∴ CuSO₄ is L.R.

Worksheet