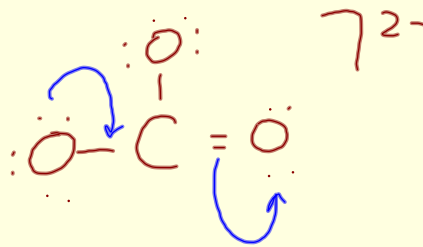
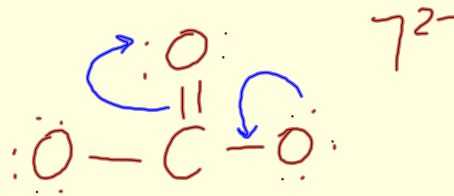
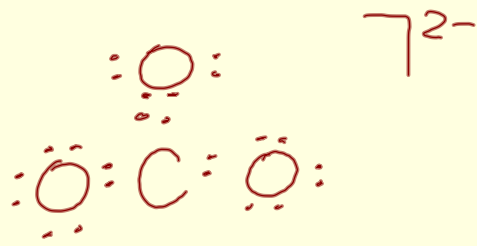
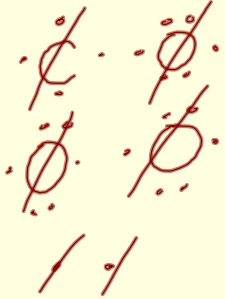
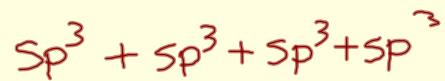
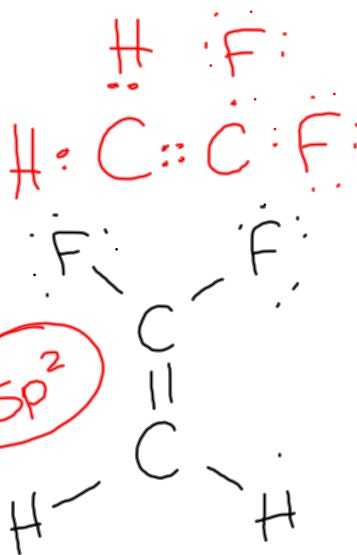
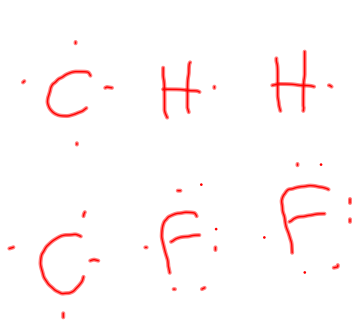
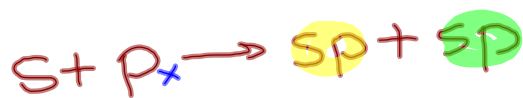
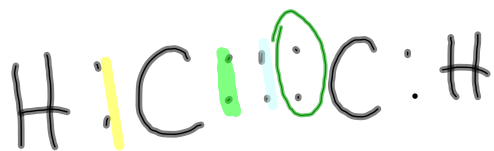
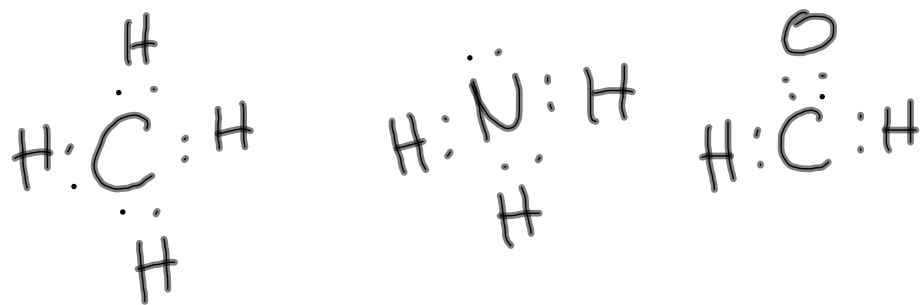


Worksheet



8.3

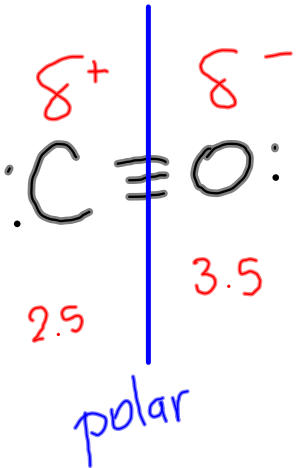
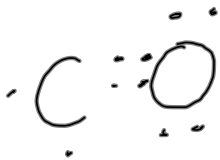




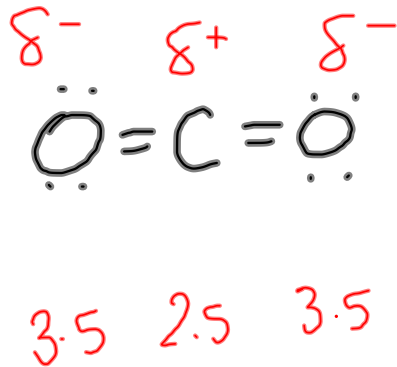
Trigonal planar = sp^2

8.4

③



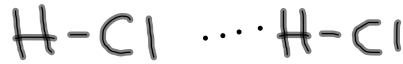
nonpolar



Attraction Between Molecules

Intermolecular forces are weaker than both ionic and covalent bonds.

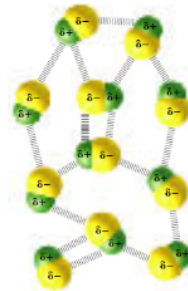
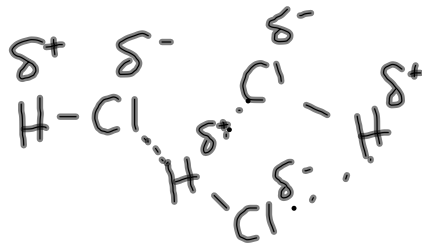
Van der Waals Forces



- Weakest attractions between molecules.
- Can be separated into two categories:

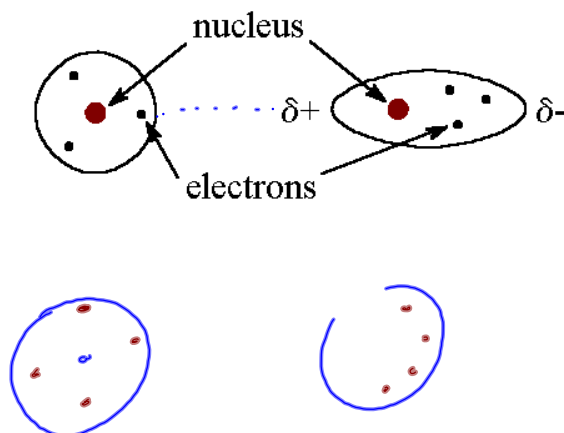
Dipole Interactions

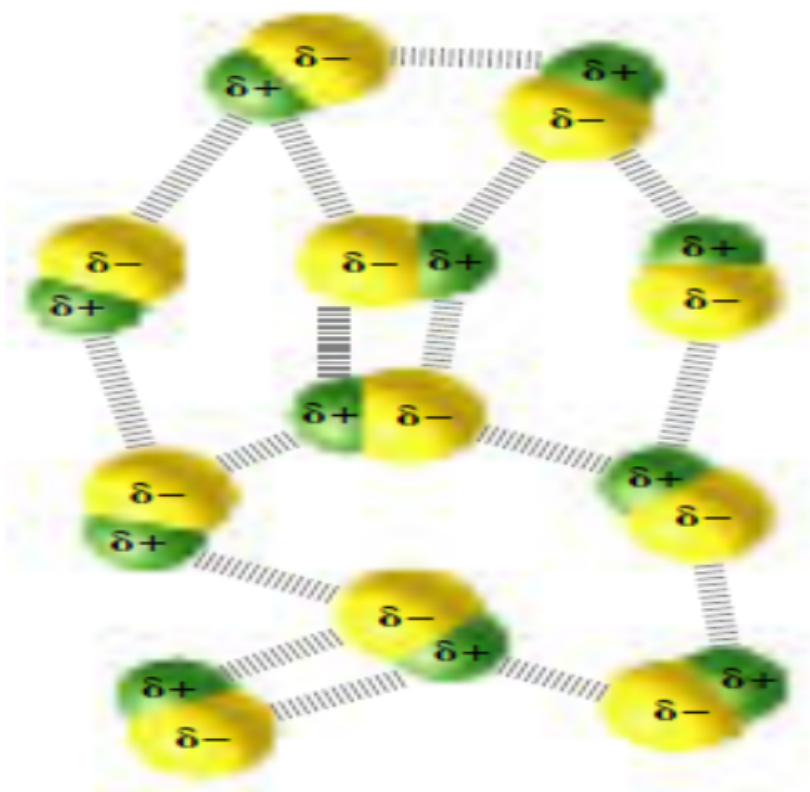
Electrical attraction between oppositely charged regions of polar molecules.

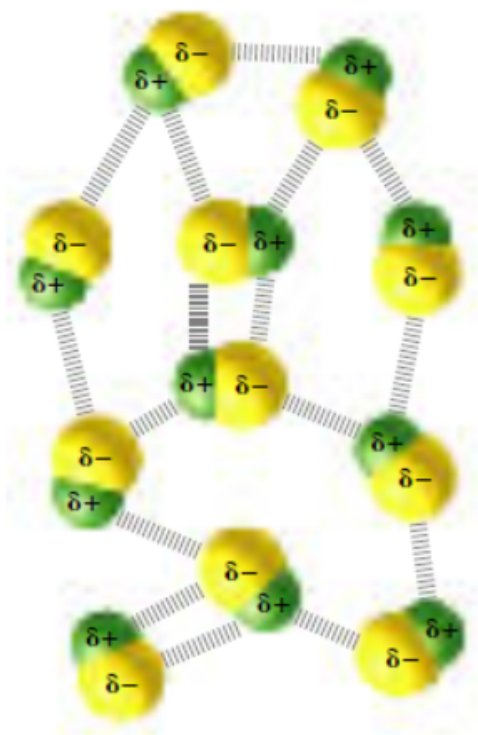


Dispersion Forces (London Dispersion Forces)

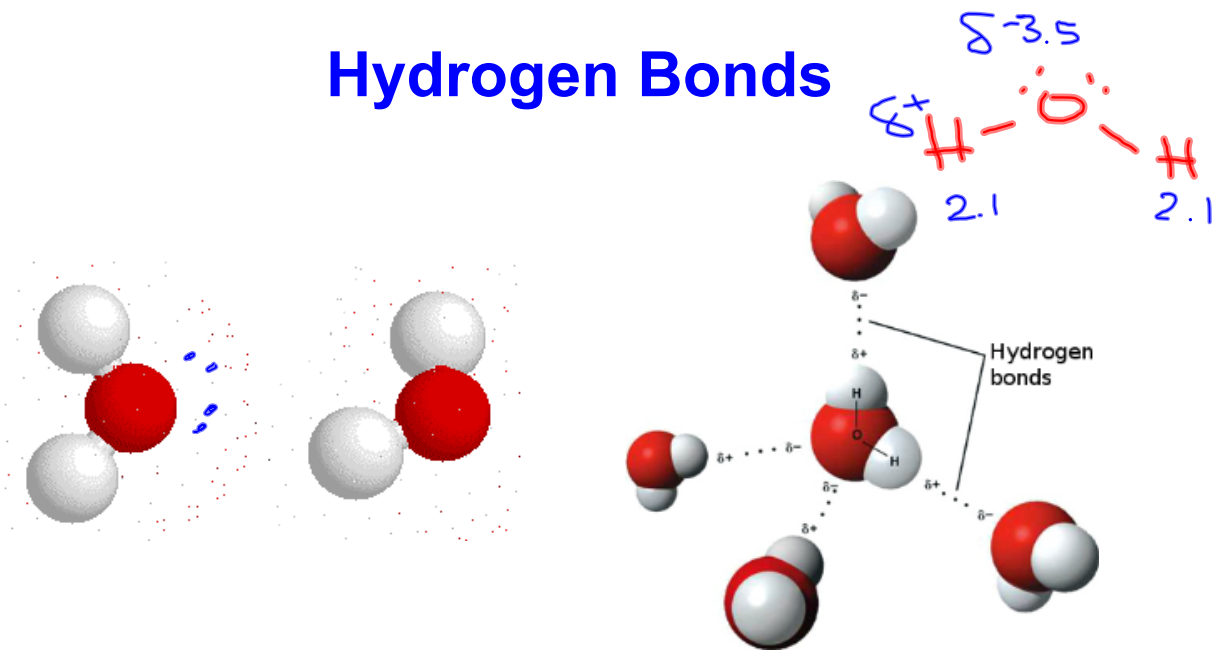
- weakest of all molecular interactions
- occur between even non-polar molecules
- caused by the motion of electrons
when moving electrons are momentarily on one side of a molecule, the electrons of the neighbouring molecule will move to the opposite side, causing a weak attraction.







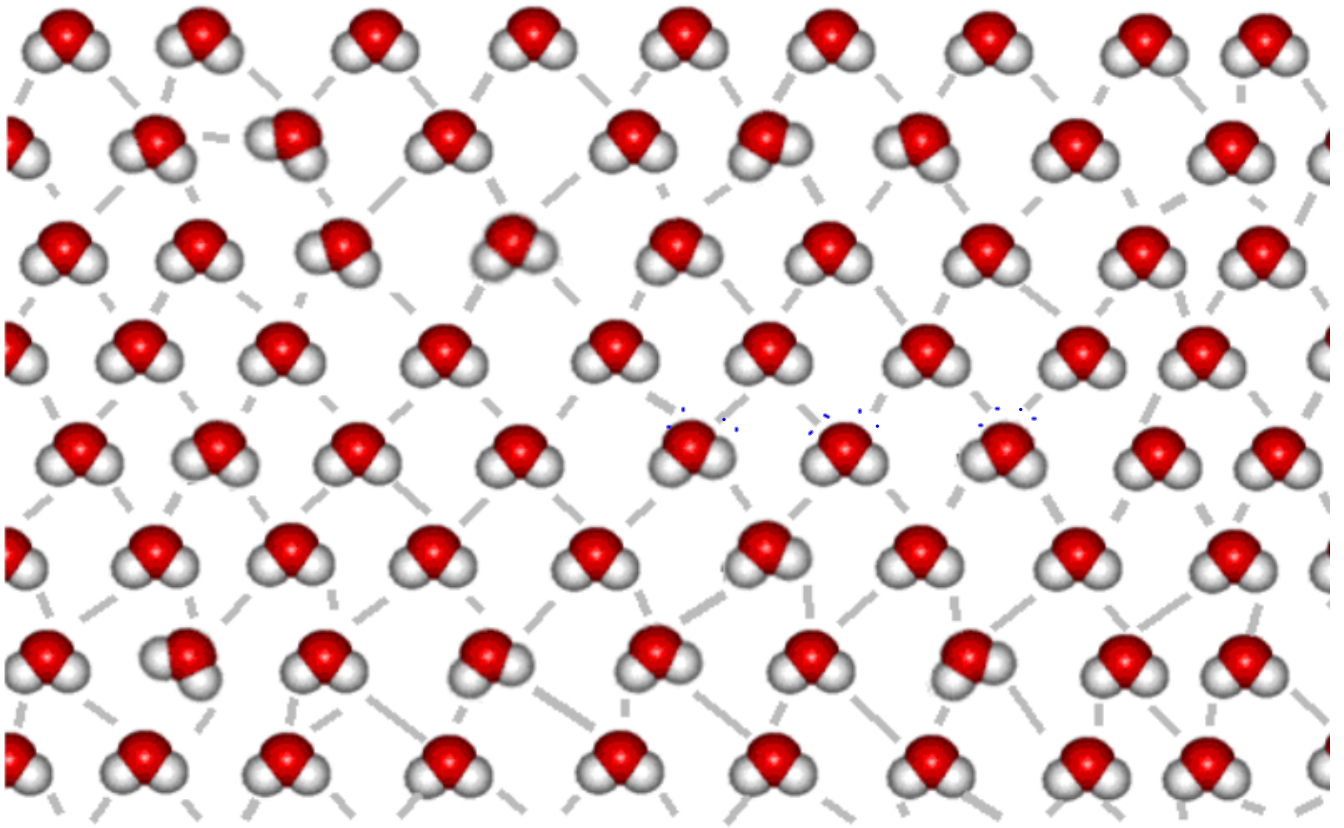
Hydrogen Bonds



Hydrogen Bonds

Strong attractive forces in which a hydrogen covalently bonded to a very electronegative atom (O, N, F), is weakly bonded to an unshared electron pair of another electronegative atom.

- strongest intermolecular force
- not as strong as an ionic or covalent bond



Homework

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