

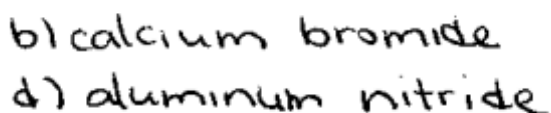
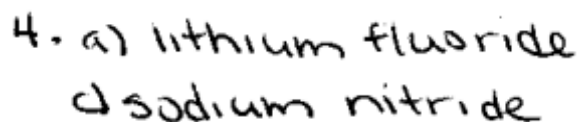
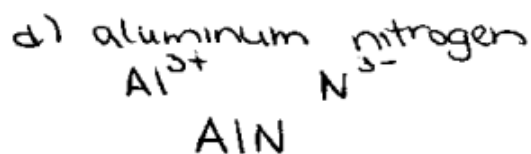
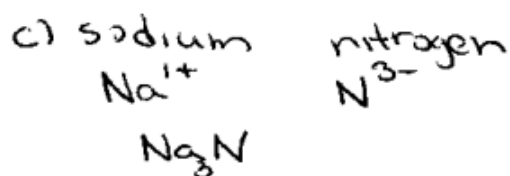
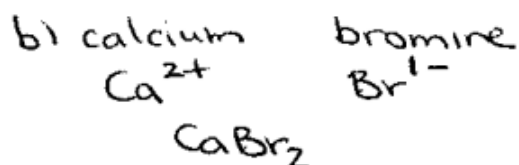
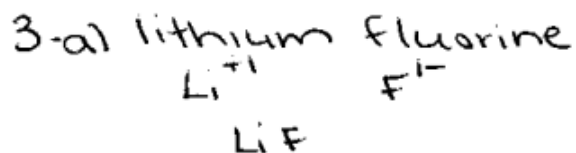
Tuesday Nov 29

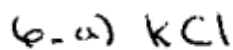
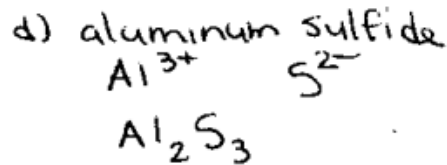
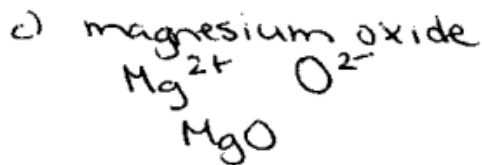
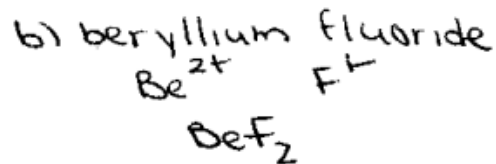
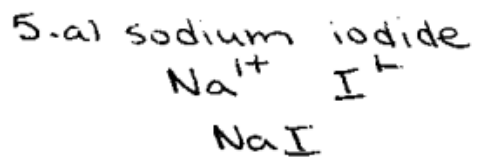
Answers pg 195 #3-6
Multivalent Compounds

Warm-Up

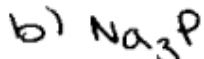
Give the formula and the names for the following compounds:

- a) sodium and iodine
- b) beryllium and fluorine
- c) magnesium and sulfur
- d) aluminum and phosphorous





potassium chloride



sodium phosphide



calcium fluoride

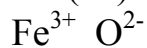
Multi-Valent Ions

some metals have more than one charge
they are called **multi-valent ions**

- these elements are found in the middle block of the periodic table i.e. Fe, Ni, Sn, Hg, Cu, Au etc (Table 2 p. 195)

the charge that is to be used is indicated in brackets with a Roman numeral

Ex. iron(III) oxide



copper (II) chloride



Try These

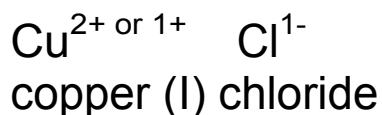
Write the formula for each of the following:

1. nickel (II) phosphide
2. copper (I) chloride
3. calcium oxide
4. Iron (III) oxide

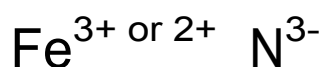
Naming Ionic Compounds from Formula (multivalent ions):

- Identify positive ion (metal) and negative ion (nonmetal)
- If metal is multivalent, determine its charge from the formula (balance total positives and negatives) and include in name

Ex. CuCl



Ex. Fe₃N₂



iron (II) nitride

pg 195 #7-10