

Answers Pg 187 #1-4

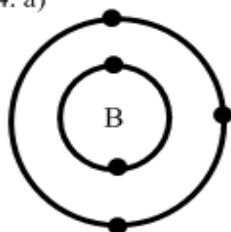
Properties	Metals	Non-Metals
Lustre	Shiny	Dull
Conductivity	Conduct electricity	Insulators
Location	Left side of table	Right side of table past staircase line
State	Mostly solids	s,l,g
e <sup>-</sup> in valence	1-9	3-8
Gain/lose	Lose	Gain
Charges of ions	+	-
Other properties	React with acids	No reaction with acid
Examples	Na,K,Li,Ca	C,O,F,Cl

2. Hydrogen is in the metallic area of the periodic table, but has non-metallic properties.

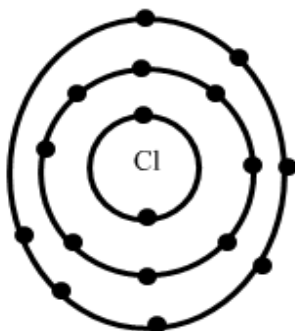
3.

Element #	e <sup>-</sup> in orbit 1	e <sup>-</sup> in orbit 2	e <sup>-</sup> in orbit 3
1	1		
2	2		
3	2	1	
4	2	2	
5	2	3	
6	2	4	
7	2	5	
8	2	6	
9	2	7	
10	2	8	
11	2	8	1
12	2	8	3
13	2	8	4
14	2	8	5
15	2	8	6
16	2	8	7
17	2	8	8
18	2	8	8
19	2	8	8
20	2	8	8

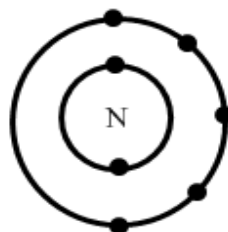
4. a)



b)



c)



d)

