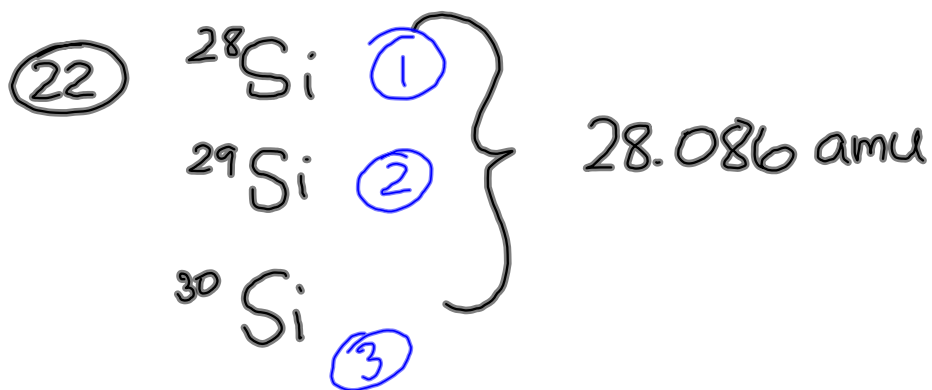


Homework #17-24



③ 62.93 amu
69.2%

64.93 amu
30.8%

$$(62.93)(0.692) + (64.93)(0.308) =$$

63.6 amu

Isotope	protons	neutrons	electrons
copper - 64	29	35	29
gold - 108	79	29	79
sulfur - 33	16	17	16

Sample Problem

Element X has two natural isotopes. The isotope with a mass of 10.012 amu (^{10}X) has a relative abundance of 19.91%. The isotope with a mass of 11.009 amu (^{11}X) has a relative abundance of 80.09%. Calculate the atomic mass of this element.

Homework

Isotope worksheet

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Homework - Isotopes Worksheet

Isotope Name	Atomic Number	Mass Number	Symbol	# of Protons	# of Neutrons
carbon - 14					
hydrogen - 2					
lawrencium - 257					