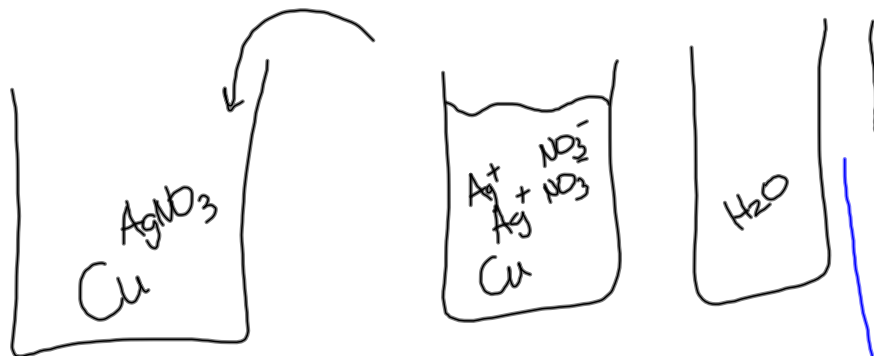


Case Study 7.4

Check Thursday's work

- ① a) Compound + compound \rightarrow D.R.
- b) element + compound \rightarrow S.R.
- c) element + element \rightarrow SYNTHESIS
- d) compound \rightarrow DECOMP.
- _____ + O₂ \rightarrow

2.a) copper + silver nitrate \rightarrow silver + copper(II) nitrate



Kinetic Molecular Theory

Principles of KMT:

- smallest particles of a substance are in constant, random motion
- these particles can be atoms, ions or molecules
- these particles collide with one another
- in order for a reaction to occur, a collision must take place
- the faster they are moving, the more likely a reaction can occur
(*some move faster than others*)

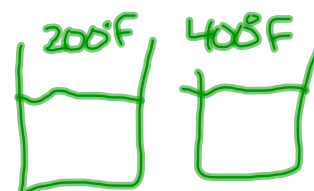
Factors affecting a Reaction

Rate of reaction - the speed at which a reaction occurs.
The greater the rate of reaction, the less time the reaction will take.

There are four factors affecting the rate of a chemical reaction:

1. Temperature

more collisions



2. Concentration of reactants

more collisions

3. Surface area

more collisions

4. Catalysts

less energy

Homework

p. 247 #1-6