

## Ionic and Molecular Compounds Review Worksheet

### Write the name for each of the following Ionic Compounds:

- |                                      |                       |                                       |                        |
|--------------------------------------|-----------------------|---------------------------------------|------------------------|
| 1. NaBr                              | sodium bromide        | 6. VCl <sub>4</sub>                   | vanadium (IV) chloride |
| 2. Ti(SO <sub>4</sub> ) <sub>2</sub> | titanium (IV) sulfate | 7. Mg <sub>3</sub> P <sub>2</sub>     | magnesium phosphide    |
| 3. FePO <sub>4</sub>                 | iron (III) phosphate  | 8. Sr(HCO <sub>3</sub> ) <sub>2</sub> | strontium bicarbonate  |
| 4. K <sub>3</sub> N                  | potassium nitride     | 9. Ba <sub>3</sub> N <sub>2</sub>     | barium nitride         |
| 5. CuOH                              | copper (I) hydroxide  | 10. Zn(NO <sub>3</sub> ) <sub>2</sub> | zinc nitrate           |

### Write the name for each of the following Molecular Compounds:

- |                                  |                               |                                   |                       |
|----------------------------------|-------------------------------|-----------------------------------|-----------------------|
| 1. CO                            | carbon monoxide               | 6. SO <sub>2</sub>                | sulfur dioxide        |
| 2. P <sub>2</sub> O <sub>5</sub> | diphosphorous pentaoxide      | 7. N <sub>2</sub> O               | dinitrogen monoxide   |
| 3. CO <sub>2</sub>               | carbon dioxide                | 8. N <sub>2</sub> O <sub>4</sub>  | dinitrogen tetraoxide |
| 4. CH <sub>4</sub>               | carbon tetrahydride / methane | 9. BCl <sub>3</sub>               | boron trichloride     |
| 5. CF <sub>4</sub>               | carbon tetrafluoride          | 10. C <sub>2</sub> S <sub>4</sub> | dicarbon tetrasulfide |

### Write the formula for each of the following Ionic Compounds:

- |                             |   |                        |                                 |
|-----------------------------|---|------------------------|---------------------------------|
| 1. nickel (III) sulfide     | Ni <sub>2</sub> S <sub>3</sub>                  | 6. potassium carbonate | K <sub>2</sub> CO <sub>3</sub>  |
| 2. copper (II) sulfate      | CuSO <sub>4</sub>                               | 7. silver oxide        | Ag <sub>2</sub> O               |
| 3. manganese (II) phosphate | Mn <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> | 8. tin (IV) selenide   | Sn <sub>3</sub> Se <sub>4</sub> |
| 4. aluminum phosphate       | AlPO <sub>4</sub>                               | 9. lithium nitride     | Li <sub>3</sub> N               |
| 5. magnesium hydroxide      | Mg(OH) <sub>2</sub>                             | 10. copper (I) sulfide | Cu <sub>2</sub> S               |

### Write the formula for each of the following Molecular Compounds:

- |                         |                                |                          |                               |
|-------------------------|--------------------------------|--------------------------|-------------------------------|
| 1. silicon dioxide      | SiO <sub>2</sub>               | 6. hydrogen peroxide     | H <sub>2</sub> O <sub>2</sub> |
| 2. diboron tetrabromide | B <sub>2</sub> Br <sub>4</sub> | 7. dicarbon tetrasulfide | C <sub>2</sub> S <sub>4</sub> |
| 3. carbon tetrachloride | CCl <sub>4</sub>               | 8. nitrogen tribromide   | NBr <sub>3</sub>              |
| 4. nitrogen monoxide    | NO                             | 9. methane               | CH <sub>4</sub>               |
| 5. sulfur dioxide       | SO <sub>2</sub>                | 10. oxygen               | O <sub>2</sub>                |