

Labs

Submitted in 5 minutes

Unit 4 - Equilibrium

- What is an equilibrium?
- Rate of reaction
 - Transition state, Activation energy
- Factors affecting Reaction Rate
- Percent Reaction / Percent Yield
- Equilibrium Law
- Le Chatelier's Principle



$$K = \frac{[\text{NH}_3(\text{g})]^2}{[\text{N}_2(\text{g})][\text{H}_2(\text{g})]^3}$$

$$[\text{H}_2(\text{g})] = \left(\frac{[\text{NH}_3(\text{g})]^2}{[\text{N}_2(\text{g})]K} \right)^{1/3}$$

$$[\text{H}_2(\text{g})] = \sqrt[3]{\frac{[.20]^2}{[0.8]0.90}}$$

$$[\text{H}_2(\text{g})] = 1.3 \text{ M}$$

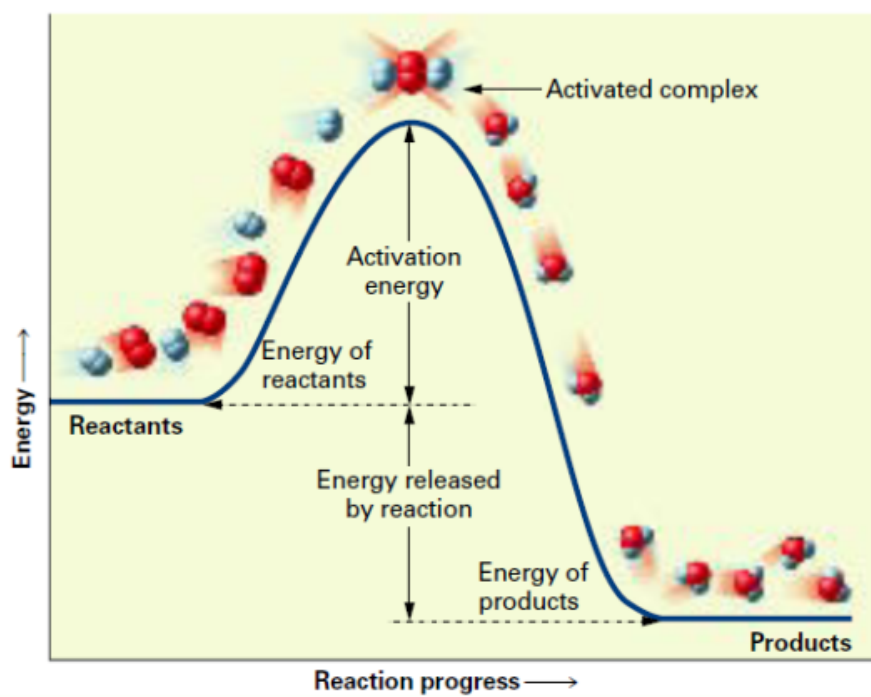


H₂ is L.R. ⇒ 283.38 g

Cl₂ is L.R. ⇒ 93.77 g

60.0 g

Energy Changes in a Reaction



Equilibrium Review Worksheet