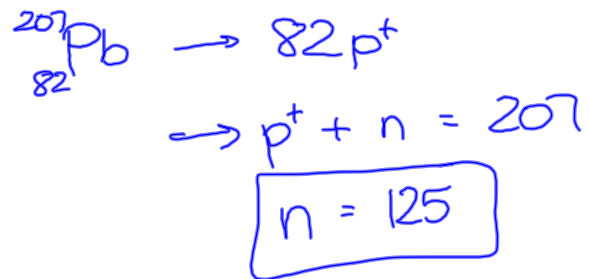


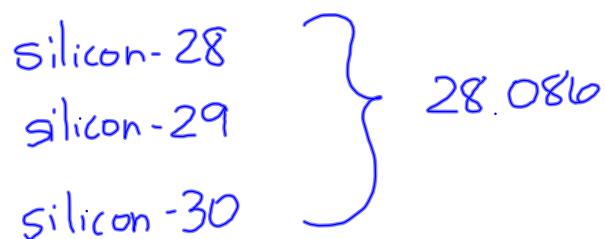
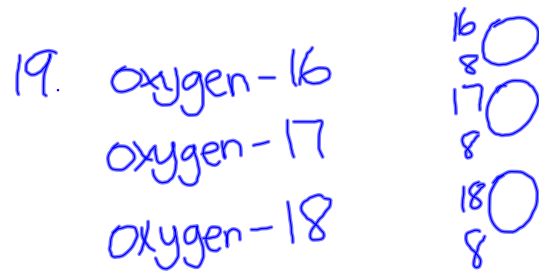
Warm Up

Isotope	protons	neutrons	electrons
copper - 64	29	35	29
chromium - 53	24	29	24
sulfur - 33	16	17	16
calcium - 41	20	21	20
gold - 108	79	29	79

Homework #17-22



18. b) fluorine-19



100
100
100
100

0

50%

85

10

70%

30%

Sample Problem

Element X has two natural isotopes. The isotope with a mass of 10.012 amu (^{10}X) has a relative abundance of 19.91%. The isotope with a mass of 11.009 amu (^{11}X) has a relative abundance of 80.09%. Calculate the atomic mass of this element.

Homework

p. 117 #23,24

Isotope worksheet

p. 119 #30-33