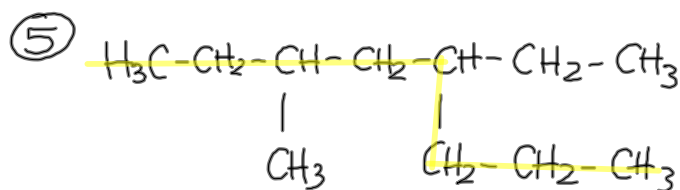
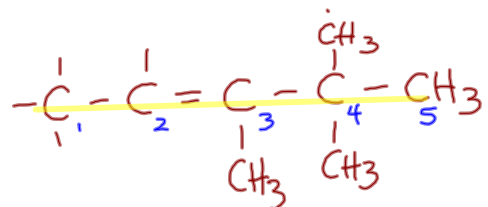
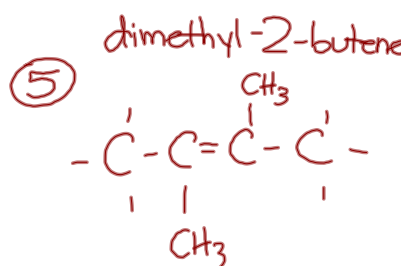
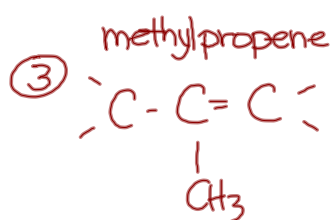


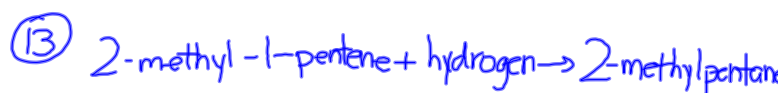
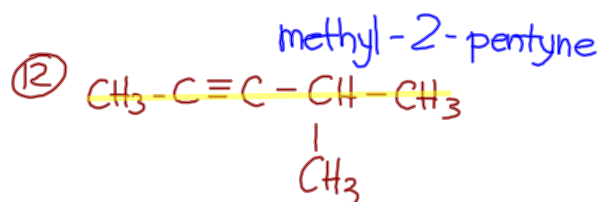
Check Homework

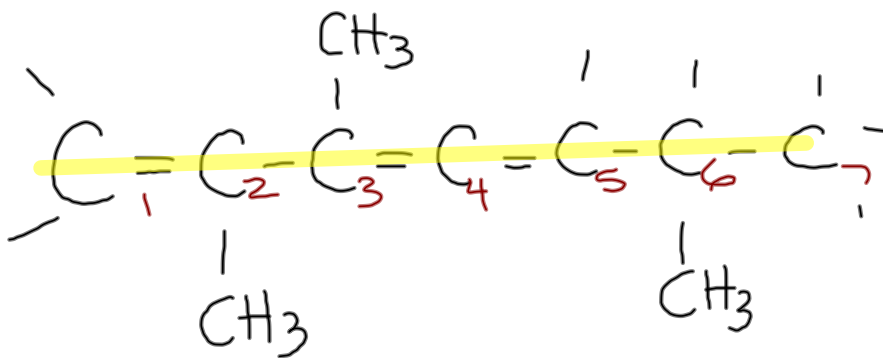


5-ethyl-3-methyloctane



3,4,4-trimethyl-2-pentene





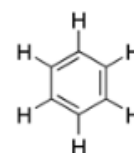
• 2,3,6-trimethyl-1,3,4-heptatriene

Aromatic Compounds

Historically aromatic compounds were organic compounds with an odour. Today aromatic compounds are defined as **benzene (C₆H₆)** and all carbon compounds that contain benzene-like structures.

Ex.

Although the molecular formula for benzene suggests 3 double bonds between three single bonds, empirical evidence shows:



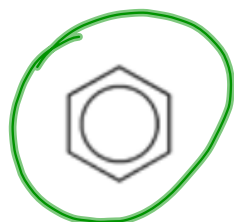
(i) the ring is relatively unreactive

← we know multiple bonds are reactive

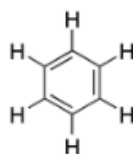
(ii) The C--C bonds are of equal length and strength

[EMPIRICAL EVIDENCE DOES NOT MATCH THEORY]

The evidence can only be explained if the pi electrons are delocalized (do not stay with any one carbon) and circle in a donut shaped cloud above and below the plane of the sp² C-C bonds.



or

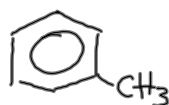


think multiple bonds

Substituted Benzenes

Mono- substituted benzene structures

Ex.

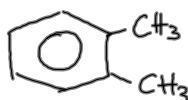


methylbenzene

No number is needed for mono-substituted benzenes because all ring positions are identical.

Simple Di - substituted benzenes

Ex.



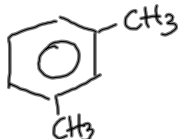
When two groups are attached to benzene, the ring is numbered to give the lower numbers to the branches.



1,2-dimethylbenzene or ortho-dimethylbenzene

The prefix meta is used for 1,3 di-substituted benzenes.

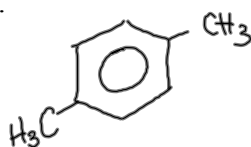
Ex.



1,3-dimethylbenzene or meta-dimethylbenzene

The prefix para is used for 1,4 di- substituted benzenes.

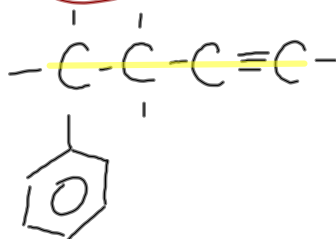
Ex.



1,4-dimethylbenzene or para-dimethylbenzene

When the benzene ring itself is considered as a branch, it is given the name *phenyl*

Ex.



4-phenyl-1-butyne

Worksheet 47 - Naming Aromatics

