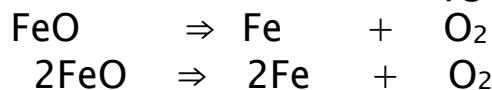


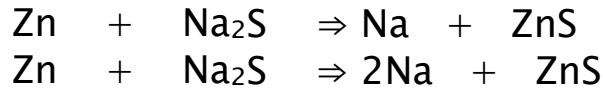
Answers
Physical Science 10
Review: Chemical Reactions

2. a) iron (II) oxide \Rightarrow iron + oxygen Type: Decomposition

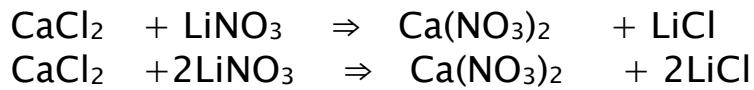


b) zinc + sodium sulphide \Rightarrow sodium + zinc sulfide

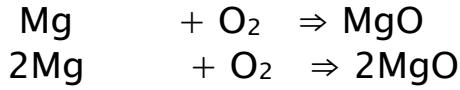
Type: Single Replacement



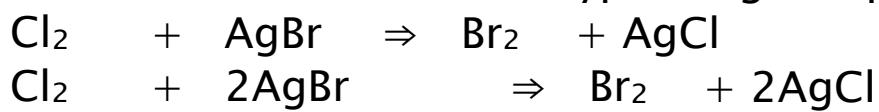
c) calcium chloride + lithium nitrate \Rightarrow calcium nitrate + lithium chloride
Type: Double Replacement



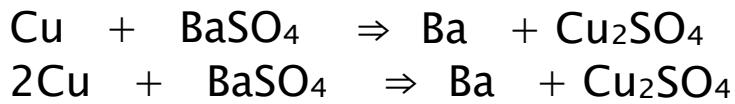
d) magnesium + oxygen \Rightarrow magnesium oxide Type: Synthesis



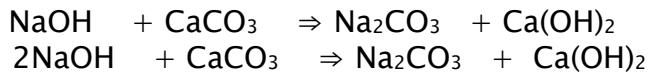
e) chlorine + silver bromide \Rightarrow bromine + silver chloride
Type: Single Replacement



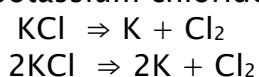
f) copper + barium sulphate \Rightarrow barium + copper (I) sulphate
Type: Single Replacement

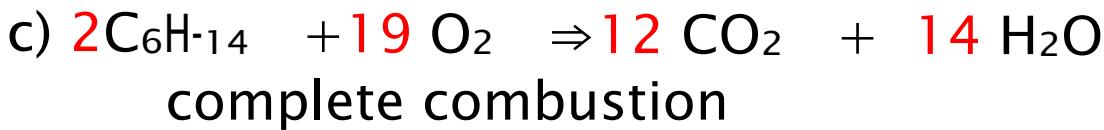
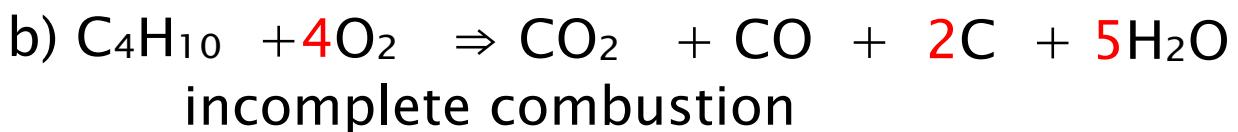
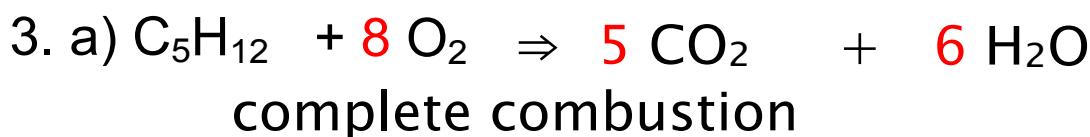


g) sodium hydroxide + calcium carbonate \Rightarrow sodium carbonate + calcium hydroxide
Type: Double Replacement



h) potassium chloride \Rightarrow potassium + chlorine
Type: Decomposition





4. a) calcium + acid \Rightarrow product + gas
 $4.5 \text{ g} + 103.8 \text{ g} \Rightarrow 109.4 \text{ g} + ?$
 $108.3 \text{ g} \Rightarrow 108 \text{ g} + ?$
 $108.3 \text{ g} - 108 \text{ g} = 0.3 \text{ g}$

b) The law of conservation of mass states that the mass of the reactants must equal the mass of the products in a reaction. The mass that is missing is the mass of the gas because it has evaporated into the air and cannot be measured in the beaker.