

Thursday Jan 12, 2012

Answers pg 241 #1-3

Review Reactions Types

Review Questions

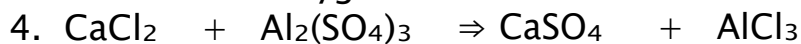
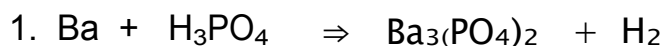
Reminder:

Test Monday Jan 16!!!

Warm- Up

For each equation state the reaction type:

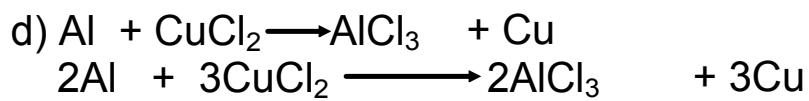
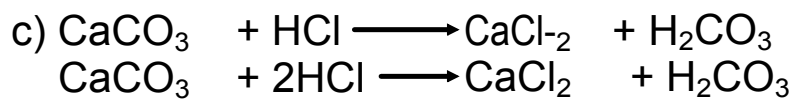
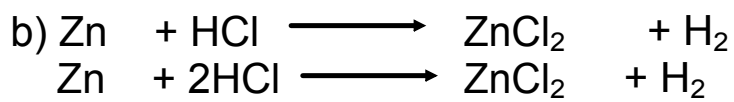
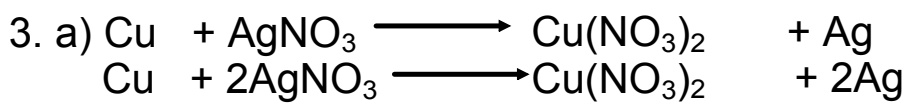
Combustion (C) , Synthesis (S), Decomposition (D), Single Replacement (SR), double replacement (DR)



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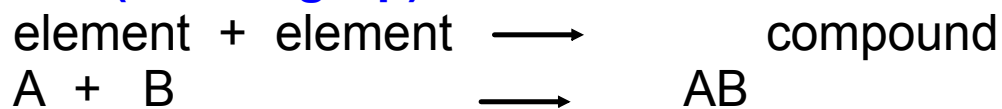
1. a) two compounds as reactants = double replacement
b) one element and one compound as reactants = single displacement
c) two elements as reactants = synthesis
d) one compound as a reactant = decomposition

2. a) copper + silver nitrate \longrightarrow silver + copper (II) nitrate
single displacement
b) zinc + hydrochloric acid \longrightarrow hydrogen + zinc chloride
single displacement
c) calcium carbonate + hydrochloric acid \longrightarrow carbonic acid + calcium chloride
double replacement
d) aluminum + copper (II) chloride \longrightarrow copper + aluminum chloride
single replacement

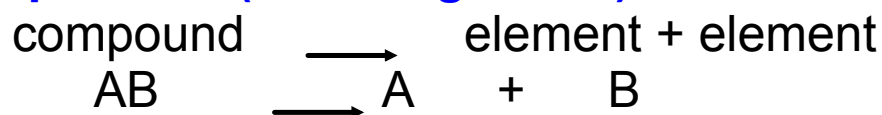


Types of Reactions

Synthesis (Building Up)



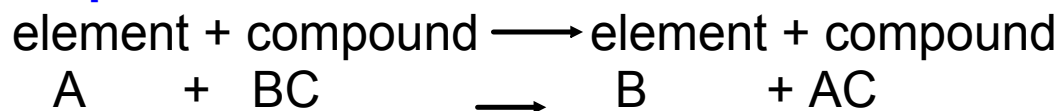
Decomposition (Breaking Down)



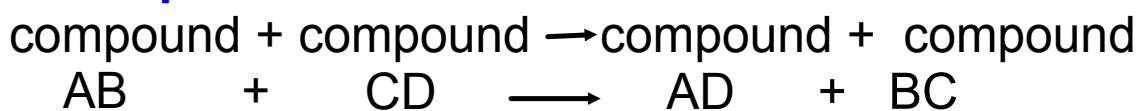
Combustion (burning)



Single Replacement



Double Replacement



Homework

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Reactions Worksheet