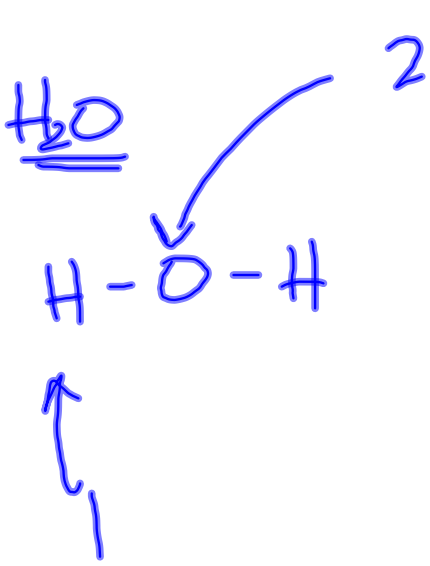
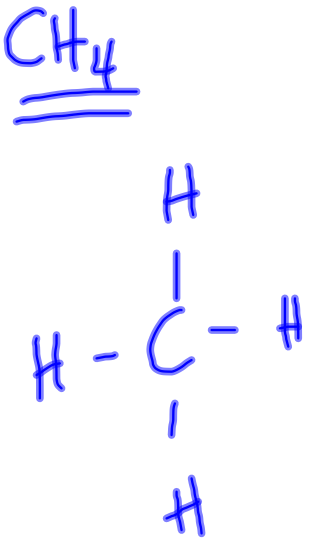
Unit 2 - Compounds

- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements
- Drawing structural diagrams
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds


H2O
2
H - O - H
1


CH4
H
H - C - H
H

Br_2 $H_2O_2N_2Cl_2I_2F_2$

P_4 S_8

$CaCl_2$

Ca^{2+} Cl^-
Cl^-

calcium chloride

sodium oxide

Na^+ O^{2-}
Na^+

Na_2O

$FeCl_3$

Fe^{3+} Cl^-
Cl^-
Cl^-

iron (III) chloride

copper (II) sulfide

Cu^{2+} S^{2-}

CuS

$LiNO_3$

Li^+ NO_3^-

lithium nitrate

sodium sulfate

Na^+ SO_4^{2-}

Na_2SO_4

Na^+

$Ca(NO_3)_2 \cdot 3H_2O$

Ca^{2+} NO_3^-

calcium nitrate - 3- water

" " trihydrate

N_2O_5

dinitrogen pentoxide

CO_2

HClO

H^+ ClO^- aqueous hydrogen hypochlorite

hypochlorous acid

chromic acid

H^+ CrO_4^{2-} H_2CrO_4

H^+

acetic acid CH_3COOH

CH_3COO^- H^+

↓

"acetate"

Review Questions p. 281-282

43-61, 65-71

Worksheet