

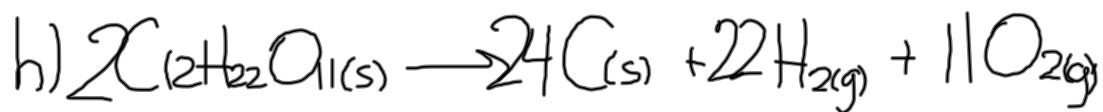
Check Homework - Worksheet



g) zinc + oxygen \rightarrow zinc oxide



DECOMP.



Sucrose \rightarrow carbon + hydrogen + oxygen



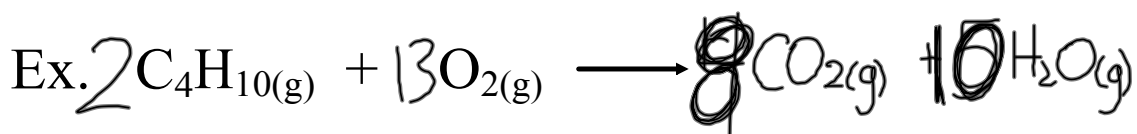
Chemical Reactions

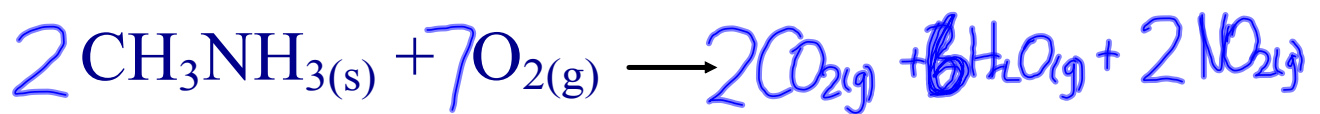
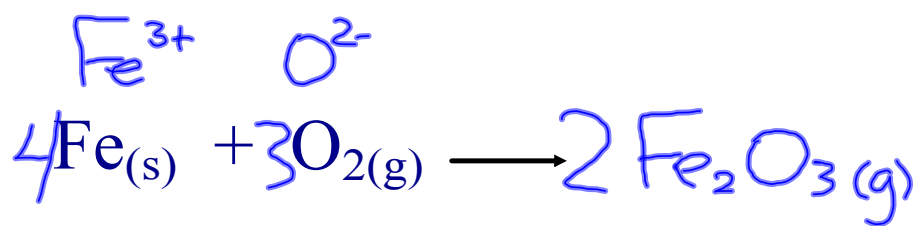
III. Combustion Reaction

A complete combustion reaction is the ^{+ O_{2(g)}} burning of a substance with oxygen to produce the most common oxides of the elements in the substance being burned.

Most Common Oxides:

- Carbon : CO_{2(g)}
- Hydrogen: H₂O_(g)
- Sulfur: SO_{2(g)}
- Nitrogen: NO_{2(g)}
- A metal: Oxide of metal with most common ion charge





$$1.12 \text{ mol NH}_3 \times \frac{6.02 \times 10^{23} \text{ molecules NH}_3}{1 \text{ mol NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 2.70 \times 10^{24}$$

③ 92.12 g/mol

$$0.780 \text{ mol Ca(CN)}_2 \times \frac{92.12 \text{ g}}{1 \text{ mol}} = 71.9 \text{ g}$$

Combustion Reactions

Write a balanced chemical equation for the following combustion reactions:



Homework

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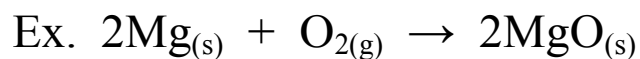
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Chemical Reactions

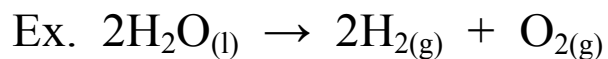
I. Formation Reactions

elements \longrightarrow compound



II. Decomposition Reactions

compound \longrightarrow elements



III. Combustion Reaction

substance + oxygen $\xrightarrow{\text{O}_2}$ most common oxides

