Reactions in Aqueous Solutions

$$Ag^{+}NO_{3}^{-}$$
 Not C(-
 $AgNO_{3(aq)} + NaCl_{(aq)} \Rightarrow AgCl_{(s)} + NaNO_{3(aq)}$

Complete Ionic Equation



An equation that shows dissolved ionic compounds as dissociated free ions.

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$$Ag(a) + Ag(b) + Ag(b) + Ag(b) + Ag(b)$$

Spectator Ion

An ion that appears on both sides of the equation and is not directly involved in the reaction.

Net Ionic Equation

An equation for a reaction in solution that only shows the particles directly involved in the reaction.

Pb: $Agt NO_3$ Pb(s) $2 AgNO_{3(aq)} \Rightarrow 2 Ag_{(s)} + Pb(NO_3)_{2(aq)}$ Complete Ionic Equation

Po(s) $+ 2 Ag_{(aq)} + 2 NO_{3(aq)} - 2 Ag_{(s)} + Pb_{(aq)}^{24} + 2 NO_{3(aq)}$ Spectator Ions: $NO_{3(aq)}$ Net Ionic Equation:

Plo(s) $+ 2 Ag_{(aq)} - 2 Ag_{(s)} + Po_{(aq)}^{24}$

Ba²⁺ CI - Agt NO₃ barium chloride and silver nitrate

Baclzan + AgNO3an -> Bacno3/2an + AgCls

Homework

Worksheet