Warm Up

What volume of 0.900M KNO_3 is needed to prepare 750. mL of a 0.450M solution?

$V_i = ?$
$C_i = 0.900M$
$V_f = 750.mL$
$C_f = 0.450M$

$V_i C_i = V_f C_f$

$V_i(0.900M) = (750mL)(0.450M)$

$V_i = \frac{(750.mL)(0.450M)}{(0.900M)}$

$V_i = 375mL$

$$C = \frac{n}{V}$$

% m/m

% v/v

$$V_1 C_1 = V_F C_F$$

Dilution Worksheet

Solutions Test

- Net Ionic Equations
- Solubility
- Concentration
- Dilutions

p. 347 #53-55

p. 499 #42, 44, 45, 48, 51-55