

Happy World
Fisheries Day, World
Television Day, and
World Hello Day!

(It's NOT Jake John James birthday :|)

Homework Questions?

Chemical Reactions

or displacement

IV. Single Replacement Reaction

Reaction of an element with a compound to produce a new element and an ionic compound.

⇒ usually occurs in aqueous solution

⇒ reaction will only occur if the element is replacing a less reactive element (see table 11.2)

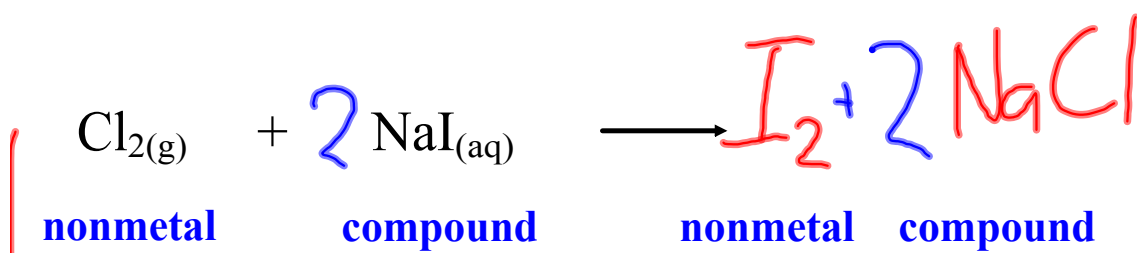
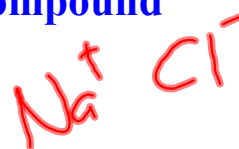
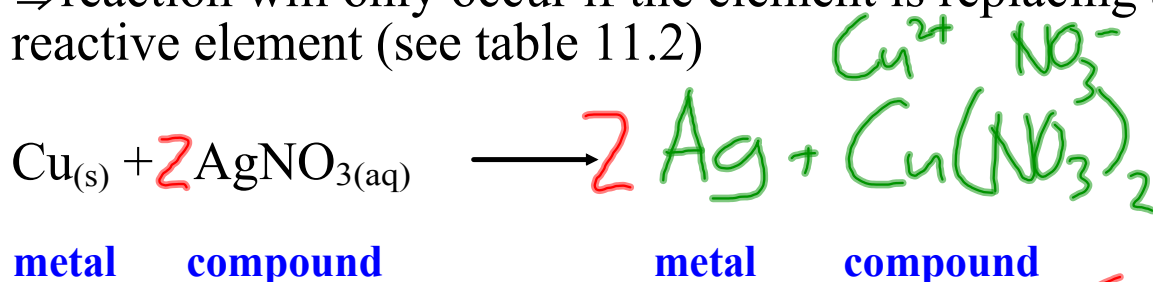
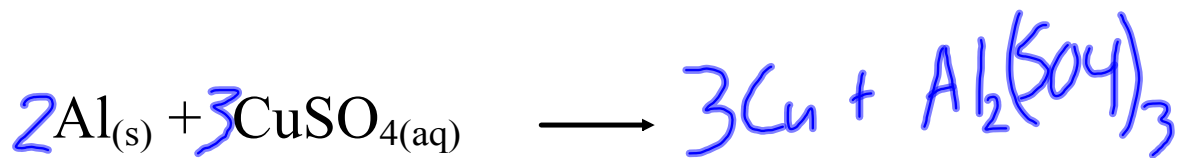
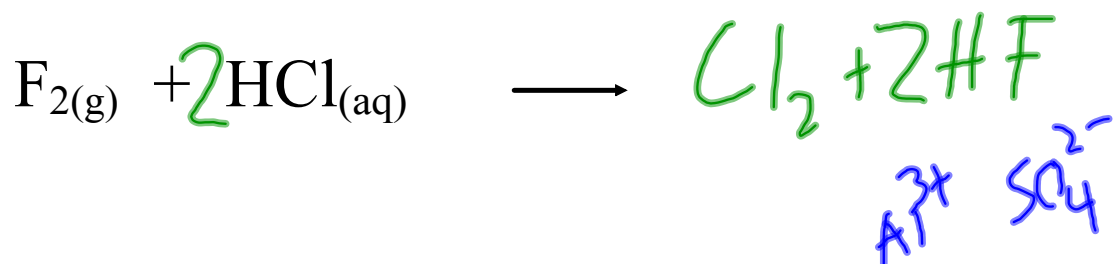
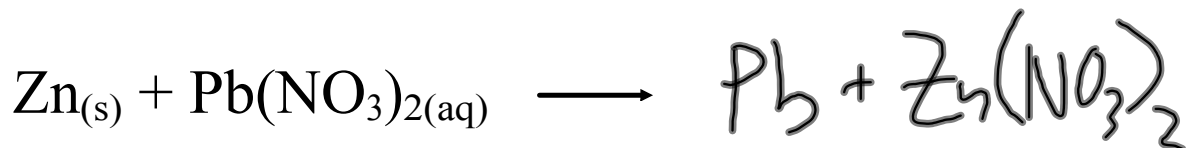


Table 11.2**Activity Series of Metals**

	Name	Symbol
Decreasing reactivity ↓	Lithium	Li
	Potassium	K
	Calcium	Ca
	Sodium	Na
	Magnesium	Mg
	Aluminum	Al
	Zinc	Zn
	Iron	Fe
	Lead	Pb
	(Hydrogen)	(H) ⁺
	Copper	Cu
	Mercury	Hg
	Silver	Ag

Practice Problems



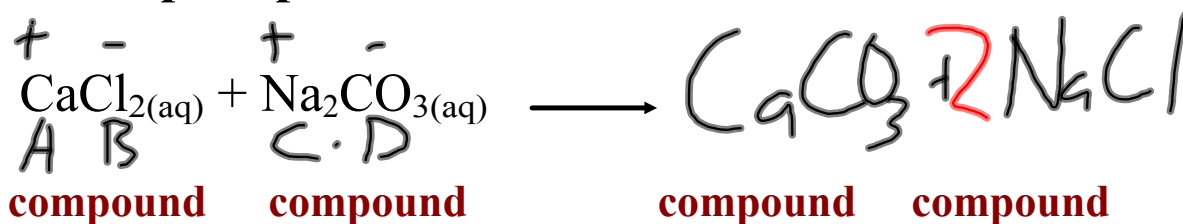
Chemical Reactions

displacement

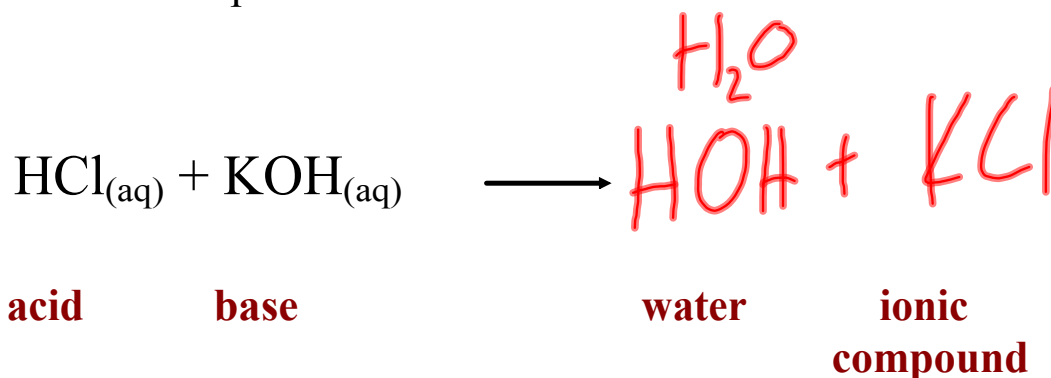
V. Double Replacement Reaction

Reaction that occurs between two ionic compounds in solution. Ions will "change partners".

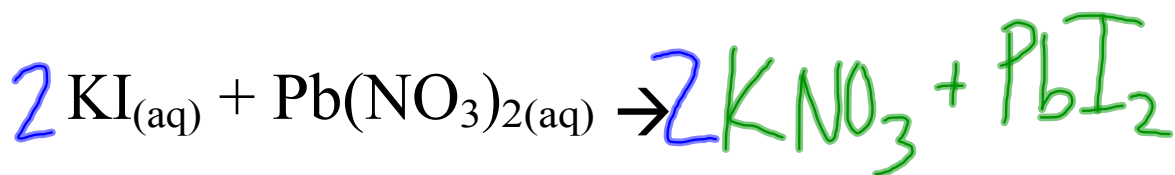
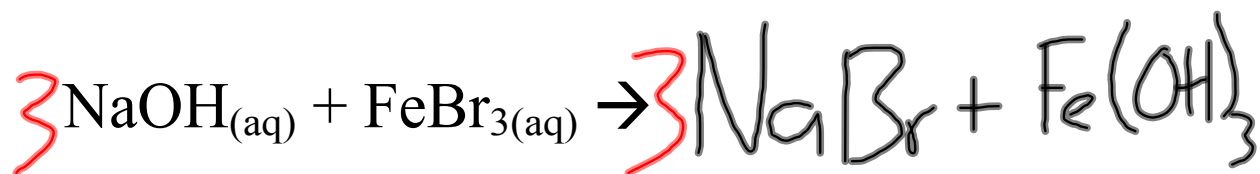
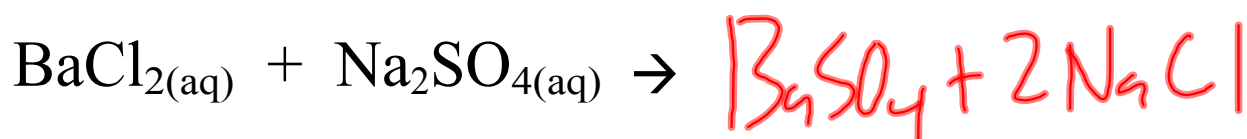
⇒ if one of the products has low solubility, it may form a precipitate (solid). This double replacement reaction is called **precipitation**.



A second type of double replacement reaction is a **neutralization** reaction, which is a reaction between an acid and a base, to form water and an ionic compound.



Practice Problems



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p. 335 #18,19
p. 339 #22-27