

Apology

Unit 2 - Compounds

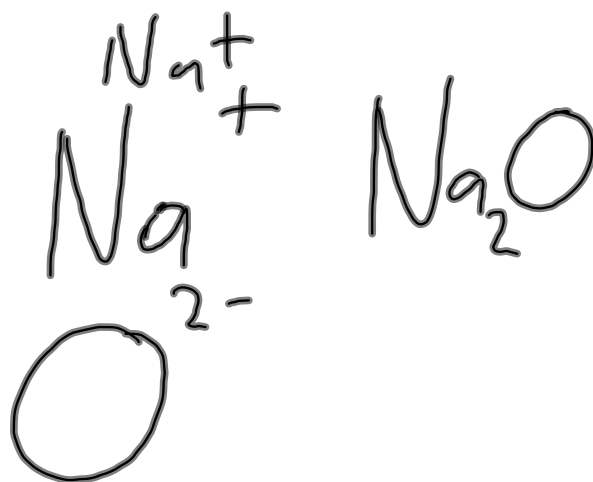
- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Element ← diatomic elements
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

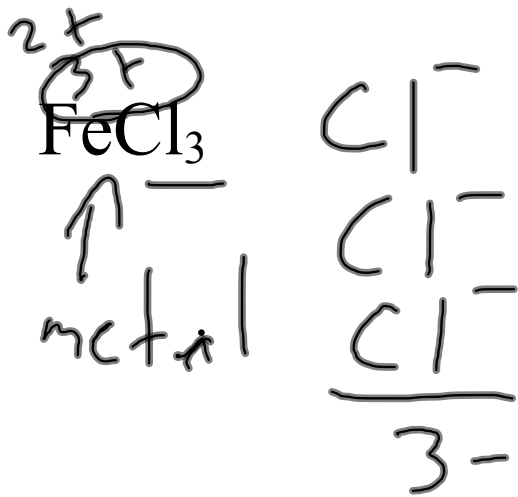


metal

Calcium Chloride

sodium oxide

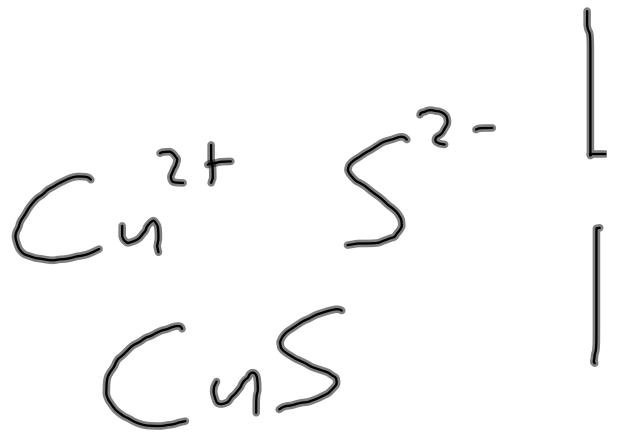




iron(III) chloride

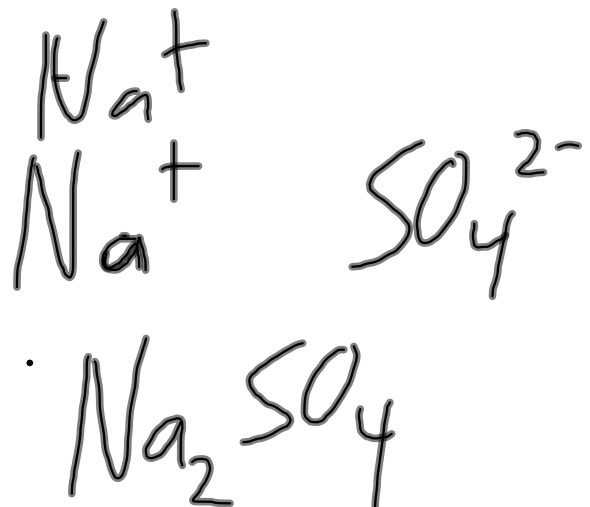
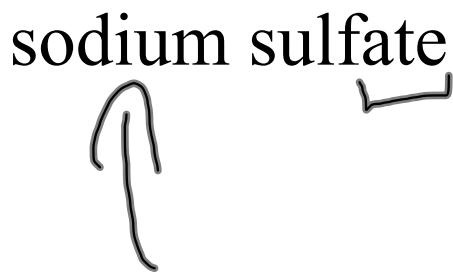
copper (II) sulfide

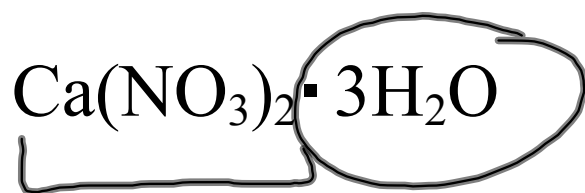
not sulfate
(SO_4^{2-})





lithium nitrate

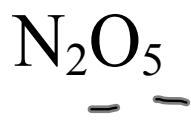




ionic

calcium nitrate

trihydrate



dinitrogen pentoxide



acid

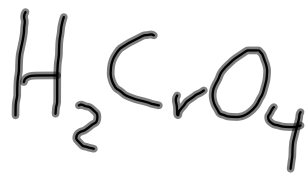


hypochlorous acid

chromic acid
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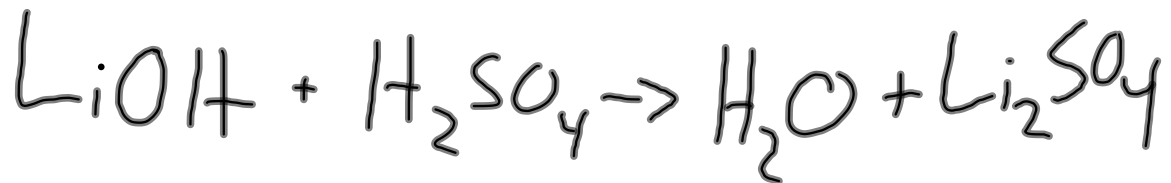
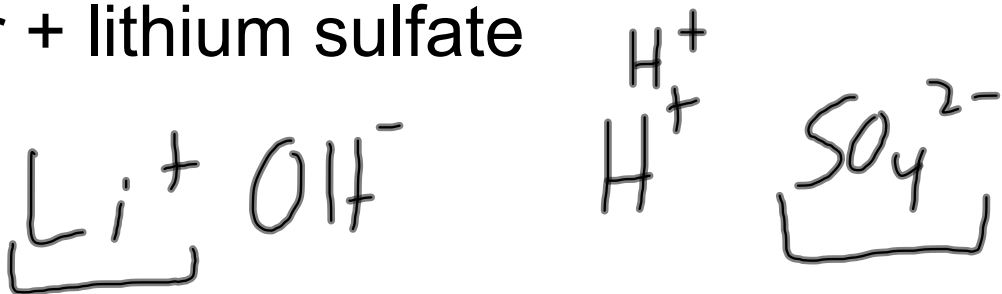
-ate

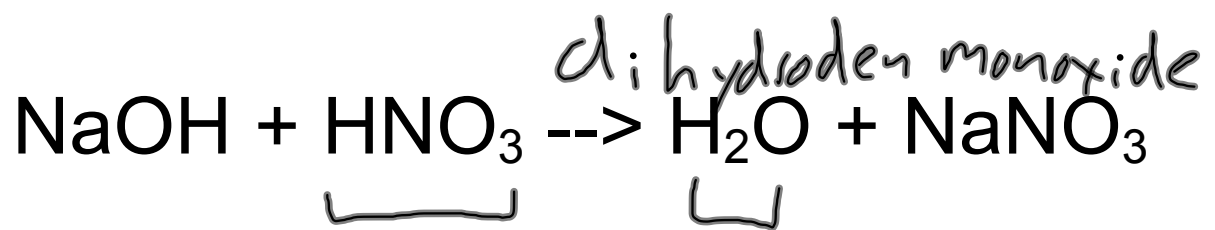
chromate



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lithium hydroxide + hydrogen sulfate -->
water + lithium sulfate





Sodium hydroxide + nitric acid →
water + sodium nitrate

Review Questions p. 281-282



43-61, 65-71