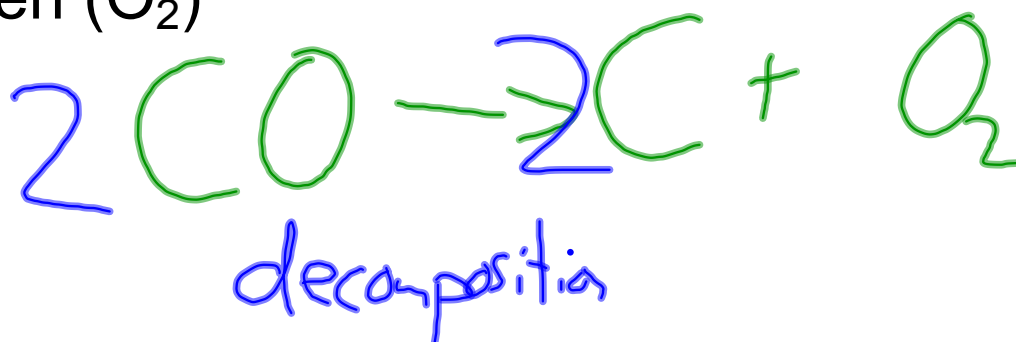
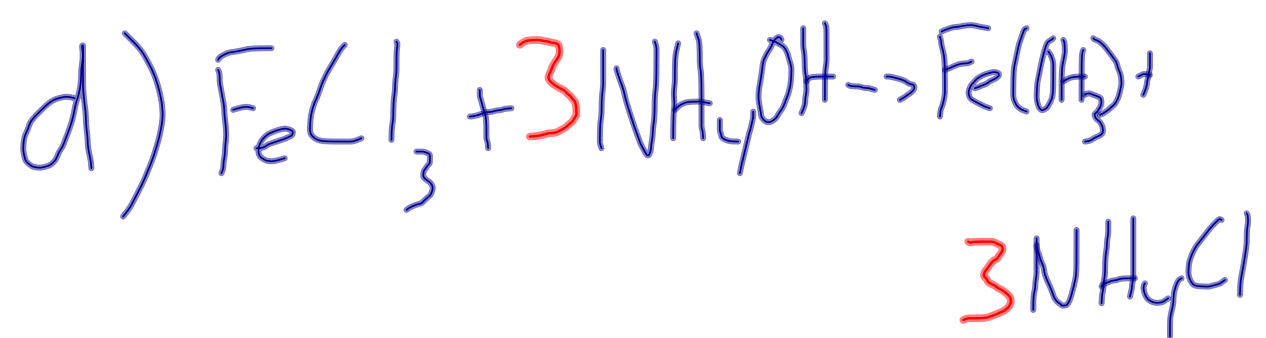
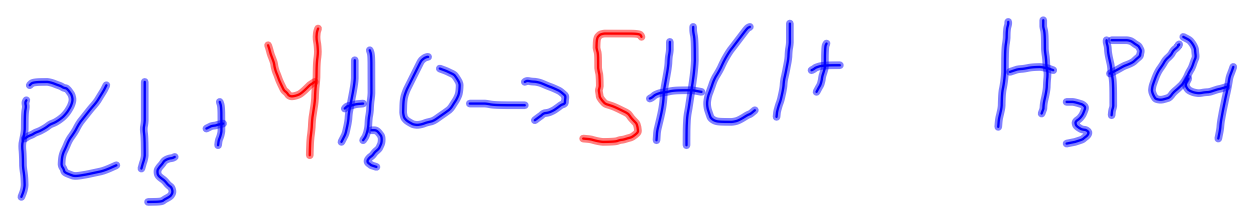


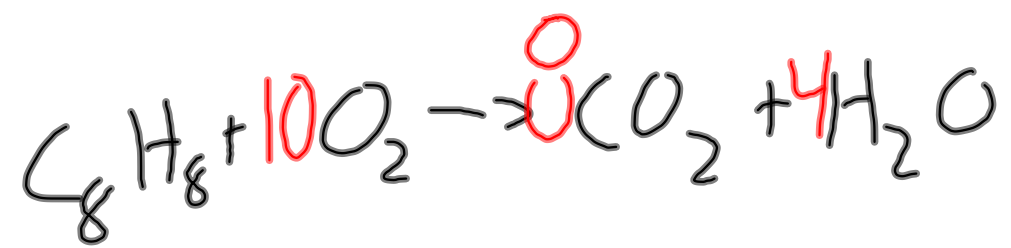
Warm-Up!

Write and balance the reaction for carbon monoxide **breaking down** into carbon and oxygen (O₂)











b)

Check Homework

Reactions so far...

Combustion

element/compound + O₂ ⇒ oxides + energy



Synthesis

two smaller particles (elements) ⇒ one molecule



Decomposition

one molecule ⇒ smaller particles (elements)

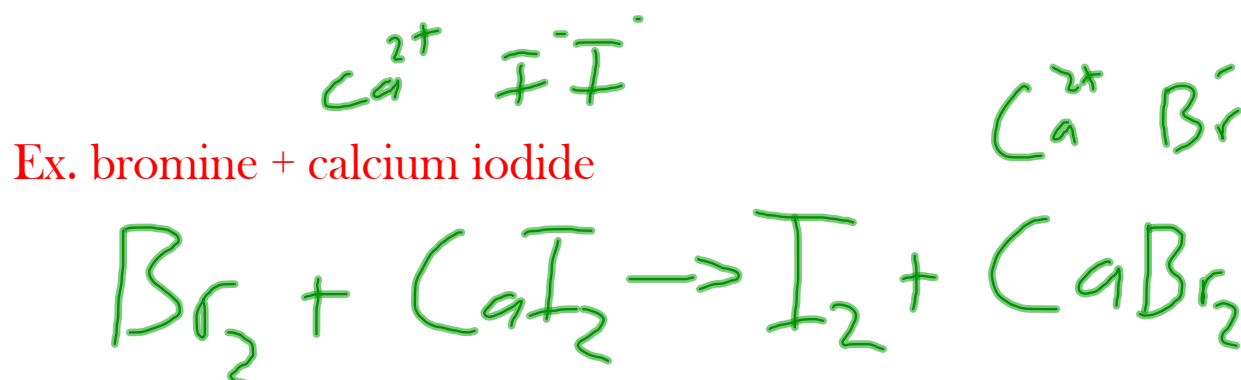
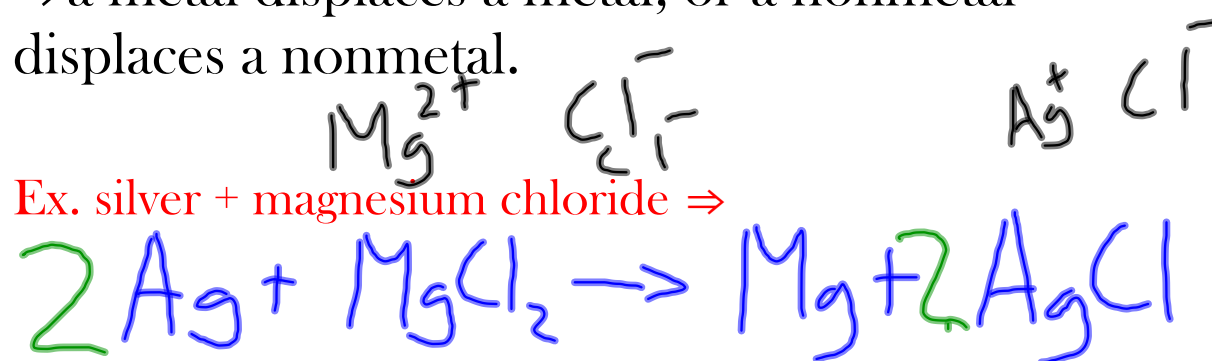


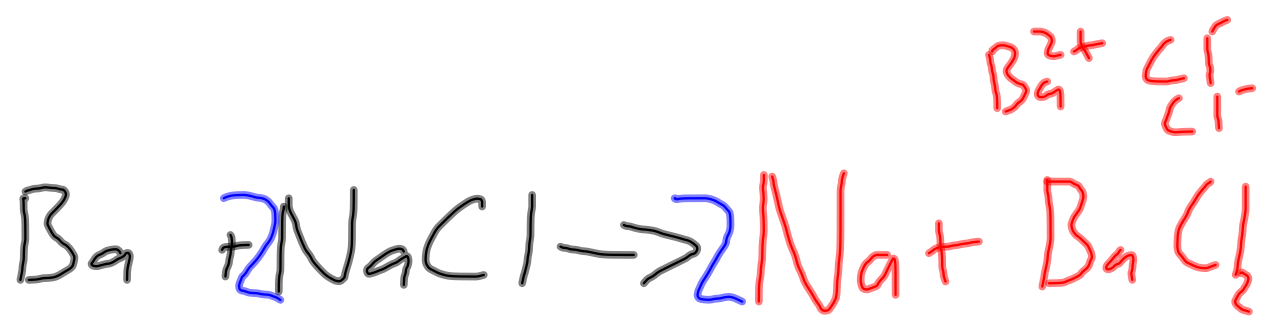
Single Replacement Reactions

IV.

Single replacement reactions are chemical changes that involve an **element and a compound** as reactants.

⇒ a metal displaces a metal, or a nonmetal displaces a nonmetal.





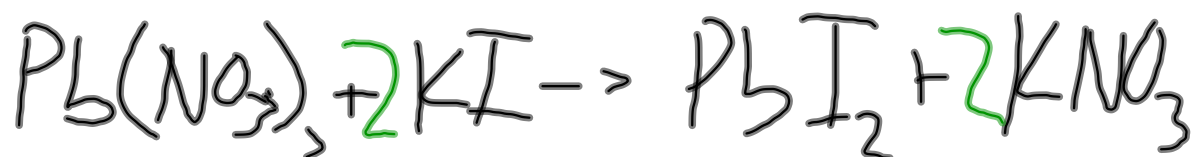
Double Replacement Reactions

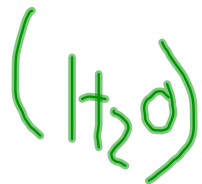
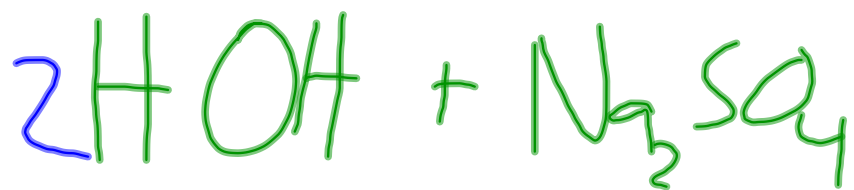
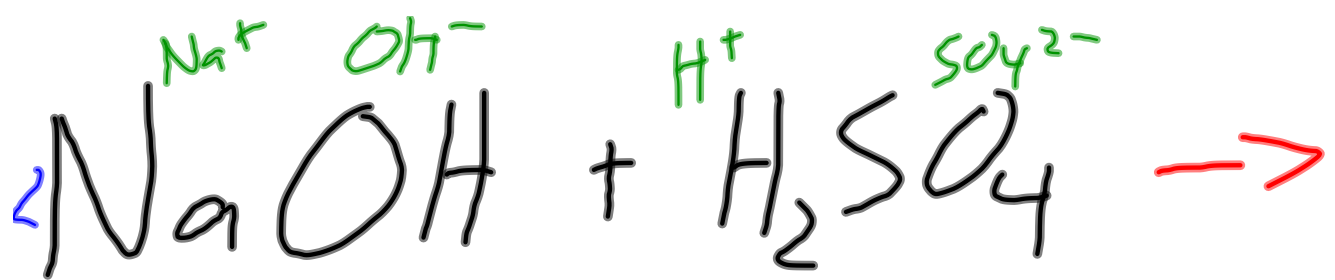
Double replacement reactions are chemical changes that involve **two compounds** as reactants.

⇒ metals (or nonmetals) will 'trade'

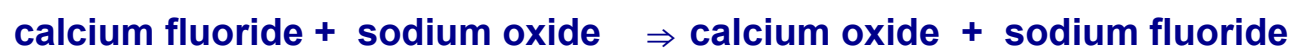


Ex. lead (II) nitrate + potassium iodide ⇒





Write a balanced chemical equation for the following word equation:



Homework

p. 241 #1-3