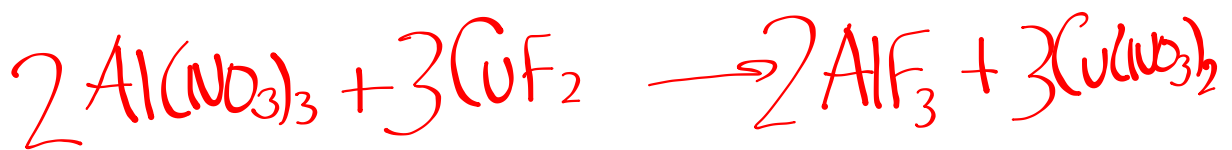
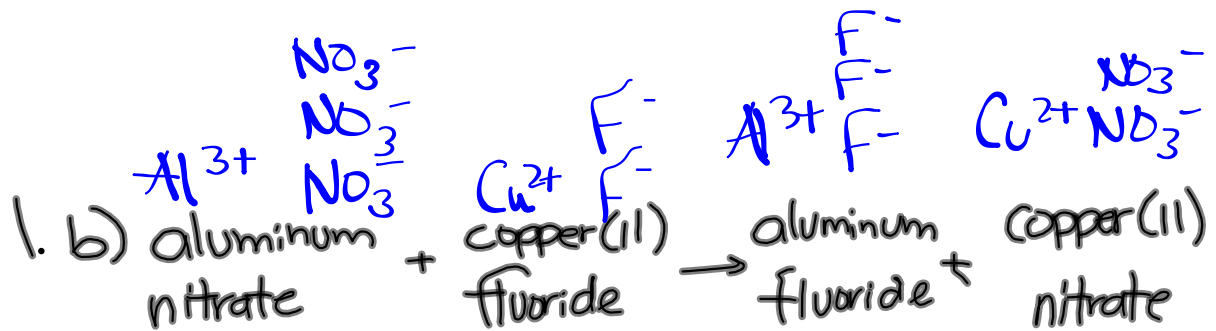
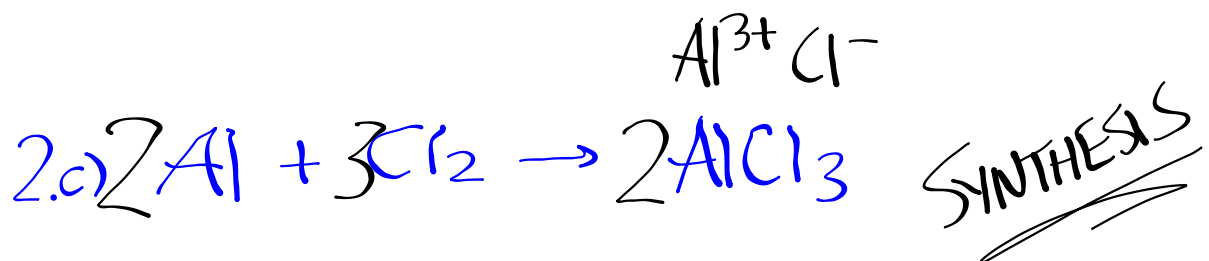
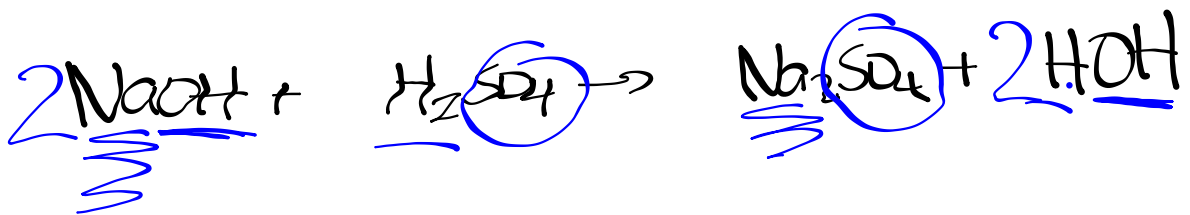


Check Thursday's work



DOUBLE
REPLACEMENT



aluminum + chlorine \rightarrow aluminum chloride

Kinetic Molecular Theory

Principles of KMT:

- smallest particles of a substance are in constant, random motion
- these particles can be atoms, ions or molecules
- these particles collide with one another
- in order for a reaction to occur, a collision must take place
- the faster they are moving, the more likely a reaction can occur

(some move faster than others)

Factors affecting a Reaction

Rate of reaction - the speed at which a reaction occurs.
The greater the rate of reaction, the less time the reaction will take.

There are four factors affecting the rate of a chemical reaction:

1. Temperature

more collisions

2. Concentration of reactants

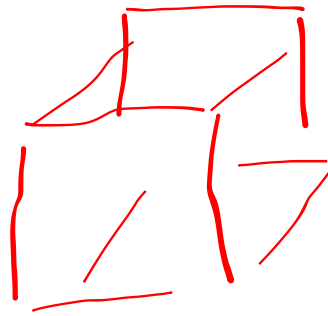
more collisions

3. Surface area (smaller particles)

more collisions

4. Catalysts

less energy



Homework

p. 264 #1-5

	How ?	Example
Temperature	Increased temp. increases rate or rxn	Ice melts quicker at higher temp.

 <http://explodingsink.com/?p=96>

 http://www.mpcfacyty.net/mark_bishop/KMT.htm