Warm Up

How many moles are in 2.14×10^{24} molecules of NO_2?

2.14×10^{24} molecules NO_2 x $\frac{1 \text{ mol } NO_2}{6.02 \times 10^{23} \text{ molecules } NO_2}$ =

$3.55 \text{ mol } NO_2$

How many atoms are in 8.08 moles of C_3H_8?

$8.08 \text{ mol } C_3H_8 \times \frac{6.02 \times 10^{23} \text{ molecules } C_3H_8}{1 \text{ mol } C_3H_8} \times \frac{11 \text{ atoms}}{1 \text{ molecule } C_3H_8}$

$= 5.35 \times 10^{25}$ atoms

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③ 4.65 mol Si

④ 0.360 mol Br_2

⑤ 2.75×10^{24} atoms

⑥ 7.72 mol NO_2

Homework

Worksheet - Molar Calculations

1 dozen = 12

1 mol = 6.02×10^{23} rep particles

1 mol C = 6.02×10^{23} atoms C

1 mol CO_2 = 6.02×10^{23} molecules CO_2