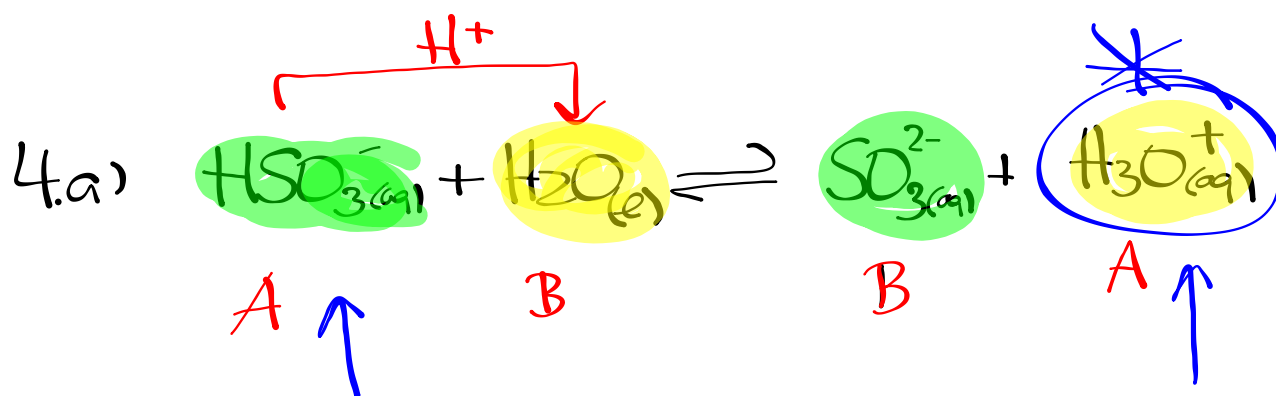
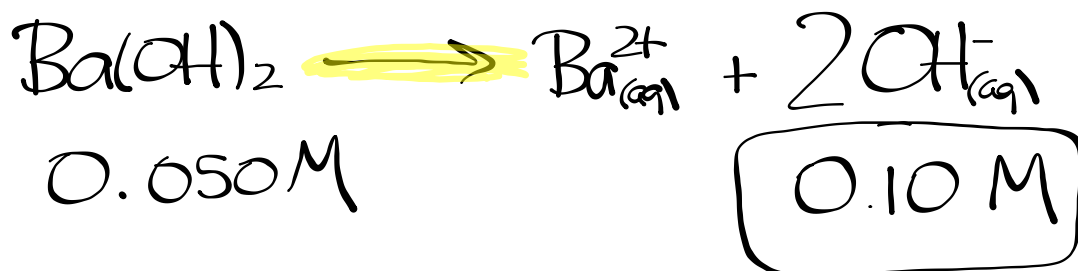


Bronsted-Lowry Worksheet



$$K_a = \frac{[\text{SO}_3^{2-}][\text{H}_3\text{O}^+]}{[\text{HSO}_3^-][\text{H}_2\text{O}]} = 1.23 \times 10^{-7}$$

Strong Bases



pH $[\text{H}^+]$
 pOH $[\text{OH}^-]$ ✓

$$\text{pOH} = -\log[\text{OH}^-_{(aq)}]$$

$$\text{pOH} = -\log[0.10]$$

$$\text{pOH} = 1.00$$

$$\text{pH} = -\log[\text{H}^+_{(aq)}]$$

$$\text{pH} = -\log[1.0 \times 10^{-13}]$$

$$\text{pH} = 13.00$$

$$K_w = [\text{H}^+_{(aq)}][\text{OH}^-_{(aq)}]$$

$$[\text{H}^+_{(aq)}] = \frac{K_w}{[\text{OH}^-_{(aq)}]}$$

$$[\text{H}^+_{(aq)}] = \frac{1.0 \times 10^{-14}}{[0.10]}$$

$$[\text{H}^+_{(aq)}] = 1.0 \times 10^{-13} \text{ M}$$

Worksheet