

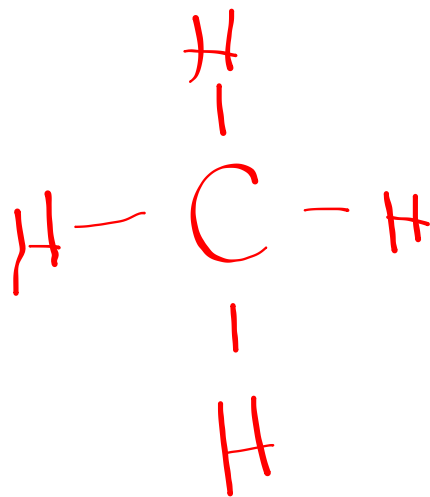
Unit 2 - Compounds

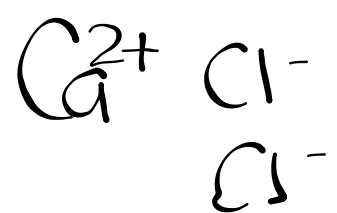
- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements
- Drawing structural diagrams
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

NaOH

Br HONCIF

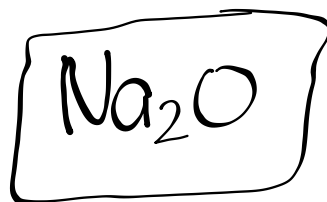
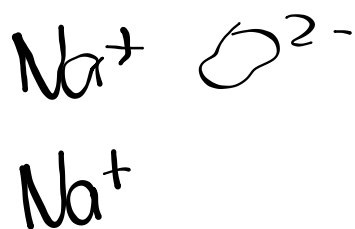
P/S
4/8

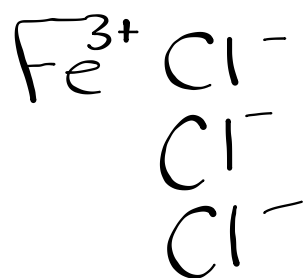




Calcium chloride

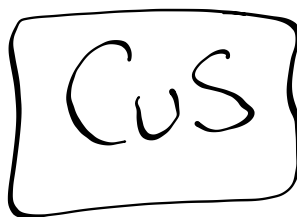
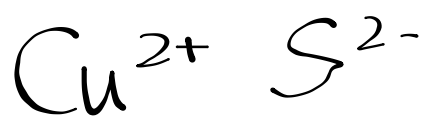
sodium oxide

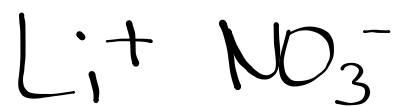




iron(III) chloride

copper (II) sulfide

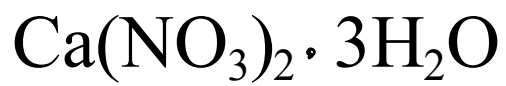




Lithium nitrate

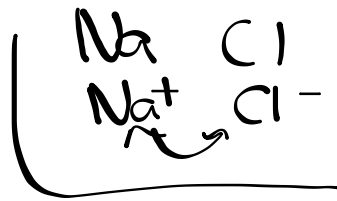
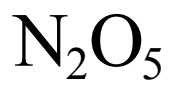
sodium sulfate





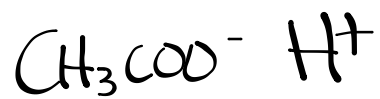
Calcium nitrate trihydrate

" " - 3-water



SHARE

dinitrogen pentoxide



aqueous hydrogen hypochlorite

hypochlorous acid

chromic acid

"-ate"



Review Questions p. 281-282

Worksheets

43-61, 65-71