

May 31, 2013

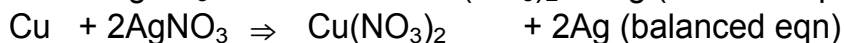
Time to complete yesterdays work/ go over answers

Quiz Monday!!

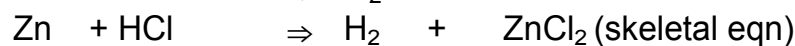
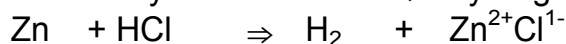
Answers pg 241 #1-3

- two compounds as reactants = double replacement
 - one element and one compound as reactants = single displacement
 - two elements as reactants = synthesis
 - one compound as a reactant = decomposition
- copper + silver nitrate \longrightarrow silver + copper (II) nitrate
single displacement
 - zinc + hydrochloric acid \longrightarrow hydrogen + zinc chloride
single displacement
 - calcium carbonate + hydrochloric acid \longrightarrow carbonic acid + calcium chloride
double replacement
 - aluminum + copper (II) chloride \longrightarrow copper + aluminum chloride
single replacement

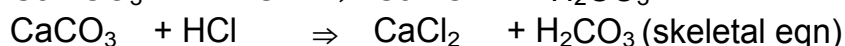
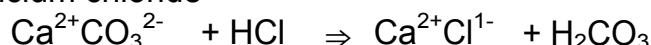
3. a) copper + silver nitrate \Rightarrow silver + copper (II) nitrate



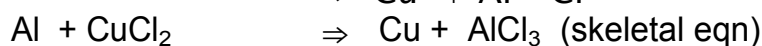
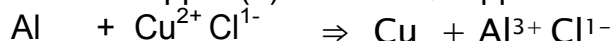
b) zinc + hydrochloric acid \Rightarrow hydrogen + zinc chloride



c) calcium carbonate + hydrochloric acid \Rightarrow carbonic acid + calcium chloride



d) aluminum + copper (II) chloride \Rightarrow copper + aluminum chloride



Answers pg 247 #1-5

1. In both synthesis and single displacement reactions, at least one reactant is an element and one new compound is produced.

2. In both decompositions and double displacement reactions, the reactant compounds are broken down.

3. Synthesis and decomposition reactions appear opposite. In synthesis a compound is formed and in decomposition a compound is broken down.

4. a) synthesis

b) single replacement

c) double replacement

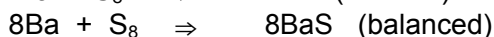
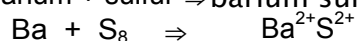
d) decomposition

e) decomposition

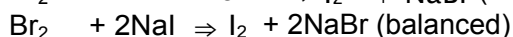
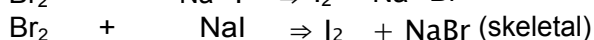
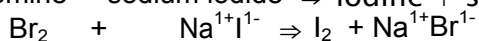
f) single replacement

g) synthesis

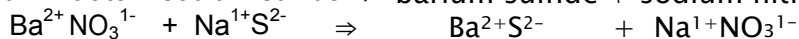
5. a) barium + sulfur \Rightarrow barium sulfide



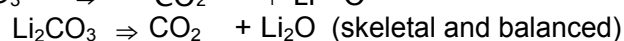
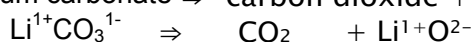
b) bromine + sodium iodide \Rightarrow iodine + sodium bromide



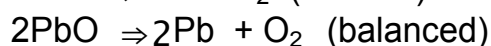
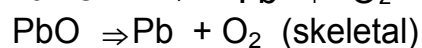
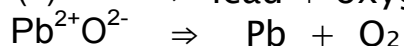
c) barium nitrate + sodium sulfide \Rightarrow barium sulfide + sodium nitrate



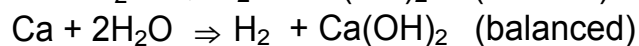
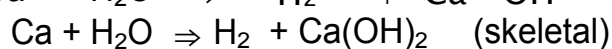
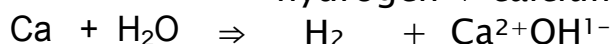
d) lithium carbonate \Rightarrow carbon dioxide + lithium oxide



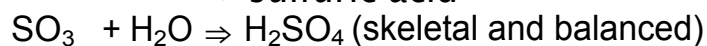
e) lead (II) oxide \Rightarrow lead + oxygen



f) calcium + water \Rightarrow hydrogen + calcium hydroxide



g) sulfur trioxide + water \Rightarrow sulfuric acid



- 1) $2 \text{NaBr} + 1 \text{Ca(OH)}_2 \rightarrow 1 \text{CaBr}_2 + 2 \text{NaOH}$ Type of reaction: **double displacement**
- 2) $2 \text{NH}_3 + 1 \text{H}_2\text{SO}_4 \rightarrow 1 (\text{NH}_4)_2\text{SO}_4$ Type of reaction: **synthesis**
- 3) $4 \text{C}_5\text{H}_9\text{O} + 27 \text{O}_2 \rightarrow 20 \text{CO}_2 + 18 \text{H}_2\text{O}$ Type of reaction: **combustion**
- 4) $3 \text{Pb} + 2 \text{H}_3\text{PO}_4 \rightarrow 3 \text{H}_2 + 1 \text{Pb}_3(\text{PO}_4)_2$ Type of reaction: **single replacement**
- 5) $1 \text{Li}_3\text{N} + 3 \text{NH}_4\text{NO}_3 \rightarrow 3 \text{LiNO}_3 + 1 (\text{NH}_4)_3\text{N}$ Type of reaction: **double replacement**
- 6) $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ Type of reaction : **decomposition**
- 7) $2 \text{KBr} + \text{F}_2 \rightarrow 2 \text{KF} + \text{Br}_2$ Type of reaction: **single replacement**
- 8) $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$ Type of reaction: **double replacement**
- 9) $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$ Type of reaction: **double replacement**
- 10) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ Type of reaction: **decomposition**
- 11) $2 \text{C}_5\text{H}_5 + \text{Fe} \rightarrow \text{Fe}(\text{C}_5\text{H}_5)_2$ Type of reaction: **synthesis**
- 12) $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$ Type of reaction: **single replacement**
- 13) $\text{C}_3\text{H}_6\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ Type of reaction: **combustion**
- 14) $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$ Type of reaction : **decomposition**
- 15) $1 \text{AlCl}_3 + 3 \text{Cs} \rightarrow 3 \text{CsCl} + 1 \text{Al}$ Type of reaction: **Single Displacement**