Reactions in Aqueous Solutions

$$A_{5}^{+}$$
 $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$ $N_{0_{3}^{-}}$

Complete Ionic Equation

An equation that shows <u>dissolved ionic compounds</u> as dissociated free <u>ions</u>.

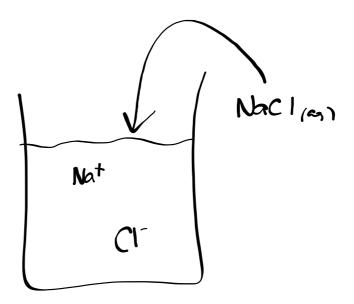
$$Ag_{\alpha i}^{+} + No_{s(\alpha i)} + No_{s(\alpha i)}^{+} + \Omega_{(\alpha i)}^{-} \rightarrow Ag(l(s)) + No_{s(\alpha i)}^{+} + No_{s(\alpha i)}^{+}$$

Spectator Ion

An ion that appears on both sides of the equation and is not directly involved in the reaction.

Net Ionic Equation

An equation for a reaction in solution that only shows the particles directly involved in the reaction.



Pb2+ ND3

Pb(s) +2AgNO_{3(aq)}
$$\Rightarrow$$
 2Ag(s) + Pb(NO₃)_{2(aq)}

Complete Tonic:

Pb(s) + 2Ag(an + 2NO_{3(aq)} \Rightarrow Ag(s) + Pb(an + 2NO_{3(aq)}

Spectator Ton(s): W3(ag)

Net Tonic: Ploss + 2Agray -> 2Agrs + Plan

Homework

Worksheet