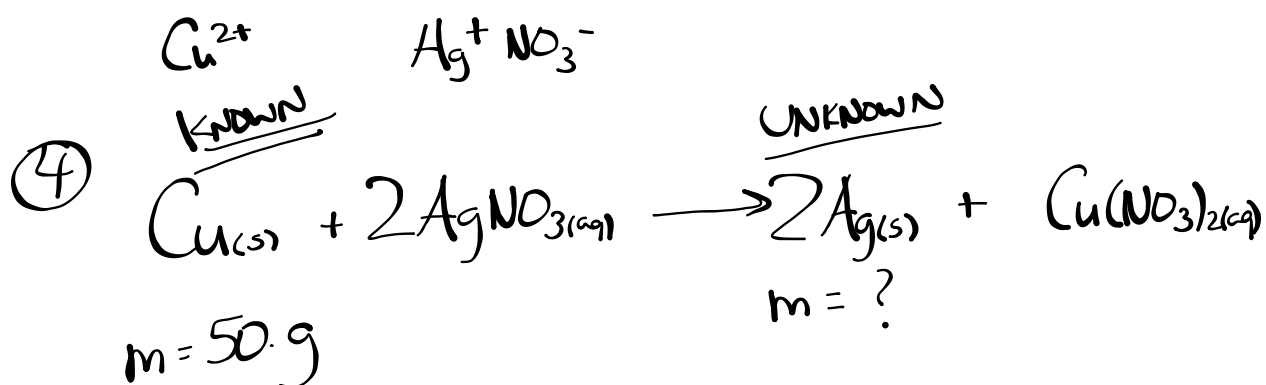


Homework - Worksheet



$$50.9 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.54 \text{ g Cu}} = 0.787 \text{ mol Cu}$$

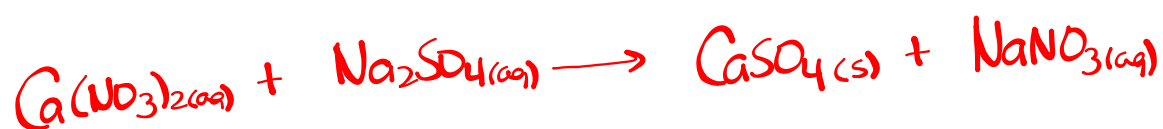
$$0.787 \text{ mol Cu} \times \frac{2 \text{ mol Ag}}{1 \text{ mol Cu}} = 1.574 \text{ mol Ag}$$

$$1.574 \text{ mol Ag} \times \frac{107.87 \text{ g Ag}}{1 \text{ mol Ag}} = \boxed{170 \text{ g Ag}}$$

Solutions Test

- Net Ionic Equations
- Properties of Solutions
Solute/solvent, factors affecting rate of dissolving
- Solubility
- Concentration $\rightarrow C = \frac{n}{V}$
- Dilutions $v_i C_i = v_f C_f$
- Stoichiometry

Ca^{2+} NO_3^- Na^+ SO_4^{2-}
 Calcium nitrate and sodium sulfate

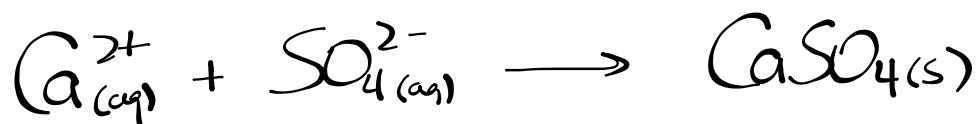


Complete Ionic:



Spectator Ions: $\text{NO}_3^-_{(\text{aq})}$, $\text{Na}^+_{(\text{aq})}$

Net Ionic:



What mass of calcium fluoride is found in 2.45 L of a 0.847M solution?

$$m = ?$$



$$V = 2.45 \text{ L}$$

$$C = 0.847 \text{ mol/L}$$

$$C = \frac{n}{V}$$

$$0.847 \text{ mol/L} = \frac{n}{2.45 \text{ L}}$$

$$n = (0.847 \text{ mol/L})(2.45 \text{ L})$$

$$n = 2.075 \text{ mol}$$

$$2.075 \text{ mol CaF}_2 \times \frac{78.08 \text{ g CaF}_2}{1 \text{ mol CaF}_2} = \boxed{162 \text{ g CaF}_2}$$

Test Review

Worksheet