Warm Up

$$3Br_{2(1)} + 2AlCl_{3(aq)} \longrightarrow 3Cl_{2(aq)} + 2AlBr_{3(aq)}$$

$$\frac{\text{nonmotal}}{\text{dement} + compound}$$

$$2Al_{(s)} + 3Pb(NO_3)_{2(aq)} \longrightarrow 3Pb_{(s)} + 2A(NO_3)_{3(aq)}$$

$$\frac{\text{nonmotal}}{\text{model}}$$

Homework - #17

Chemical Reactions

V. Double Replacement Reaction

Reaction that occurs between two ionic compounds in solution. Ions will "change partners".

⇒if one of the products has low solubility, it may form a precipitate (solid). This double replacement reaction is called **precipitation**.

A second type of double replacement reaction is aneutralization reaction, which is a reaction between an acid and a base, to form water and an ionic compound.

DOUBLE REPLACEMENT

Compound + compound -> compound + compound

Practice Problems

$$BaCl_{2(aq)} + Na_2SO_{4(aq)} \rightarrow BaSO_{4(s)} + 2NaCl_{(aq)}$$

$$3\text{NaOH}_{(aq)} + \text{FeBr}_{3(aq)} \rightarrow 3\text{NaBr}_{(aq)} + \text{Fe}OH_{3(3q)}$$

$$2KI_{(aq)} + Pb(NO_3)_{2(aq)} \rightarrow 2KN_{3an} + PbI_{2(5)}$$

p. 335 #18, 19p. 339 #24-26