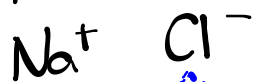


## Unit 2 - Compounds

- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements Br HONClIF / P<sub>4</sub> S<sub>8</sub>
- Drawing structural diagrams
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

IONIC

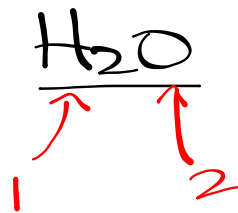
- lose metal / gain nonmetal
- transfer electrons

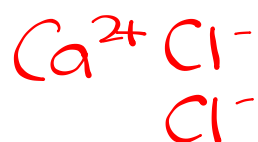


attraction

COVALENT

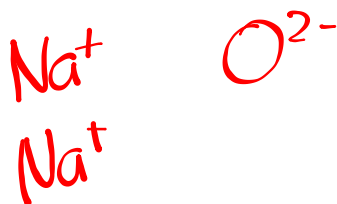
- nonmetals
- share electrons
- ~~two negatives~~

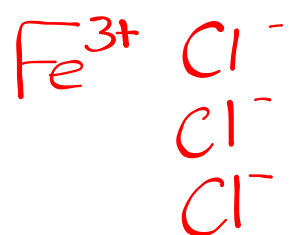




Calcium chloride

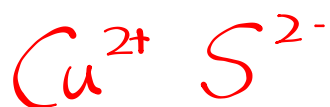
sodium oxide





iron (III) chloride

copper (II) sulfide

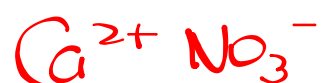
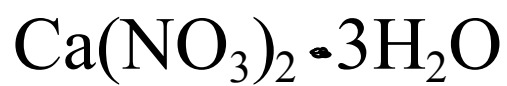




lithium nitrate

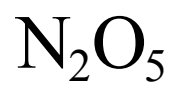
sodium sulfate





calcium nitrate - 3-water

calcium nitrate trihydrate



dinitrogen pentoxide



HClO

$H^+$   $ClO^-$  aqueous hydrogen hypochlorite

hypochlorous acid

"ate"

chromic acid

$H^+$   $CrO_4^{2-}$

$H^+$



**Review Questions p. 281-282**

**Worksheets**

**# 43-61, 65-71**