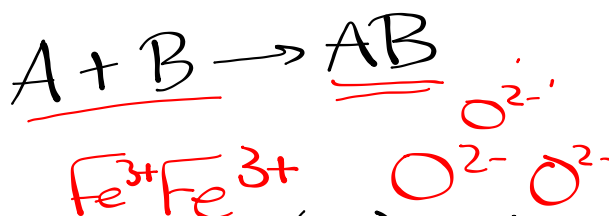
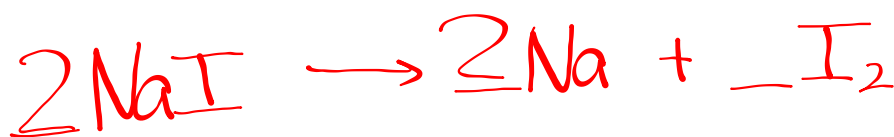
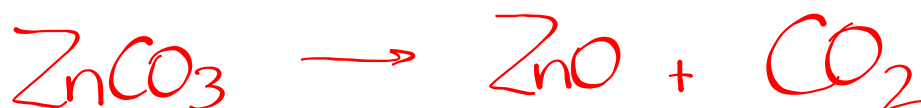


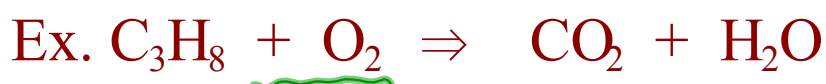
Check Homework #1-4

DECOMP.SYNTHESIS2.a) iron + oxygen \rightarrow iron (III) oxideSYNTHESIS $\text{Na}^+ \quad \text{I}^-$ b) sodium iodide \rightarrow sodium + iodineDECOMPOSITIONd) $\text{Zn}^{2+} \text{CO}_3^{2-} \rightarrow \text{Zn}^{2+} \text{O}^{2-}$ + carbon dioxide
zinc carbonate \rightarrow zinc oxide +

Reactions so far...

Combustion

element/compound + O₂ ⇒ oxides + energy



Synthesis

two smaller particles (elements) ⇒ one molecule



Decomposition

one molecule ⇒ smaller particles (elements)

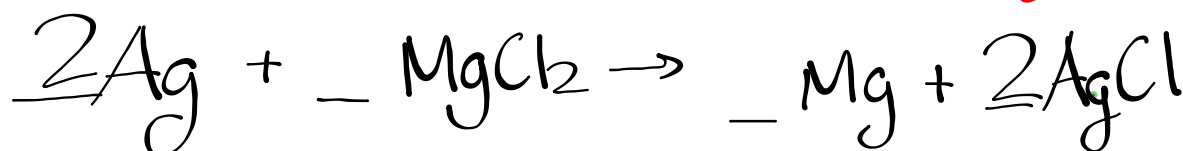


Single Replacement Reactions

Single replacement reactions are chemical changes that involve an **element** and a **compound** as reactants.

⇒ a metal displaces a metal, or a nonmetal displaces a nonmetal.

Ex. silver + magnesium chloride ⇒ magnesium + silver chloride



Ex. bromine + calcium iodide ⇒ iodine + calcium bromide

