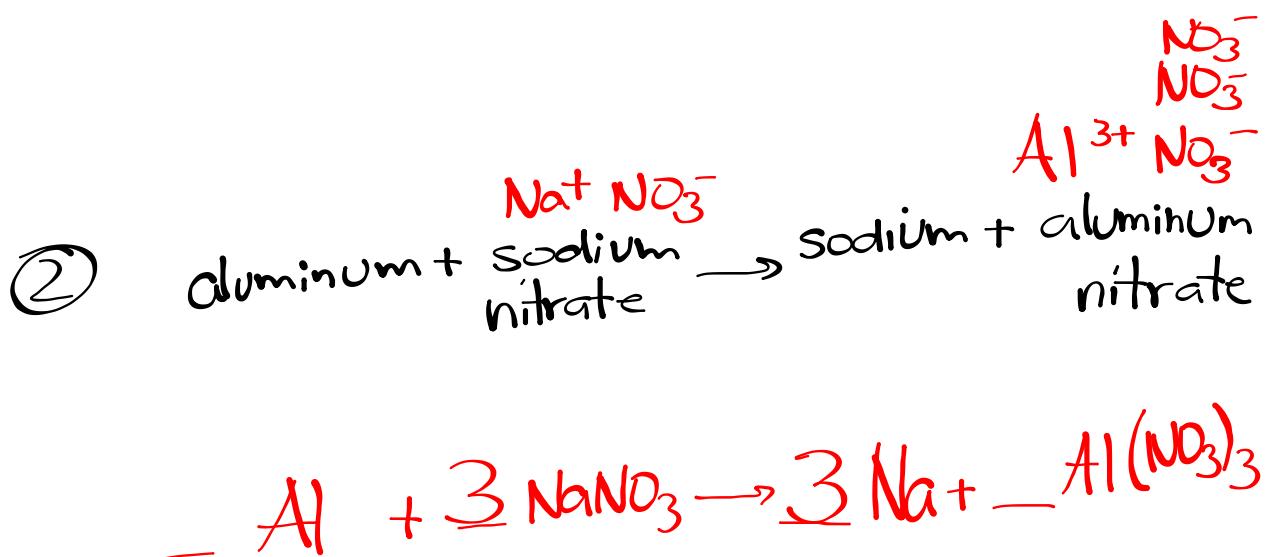


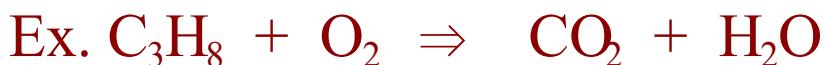
## Check Homework



## Reactions so far...

### Combustion

element/compound + O<sub>2</sub> ⇒ oxides + energy



### Synthesis

two smaller particles (elements) ⇒ one molecule



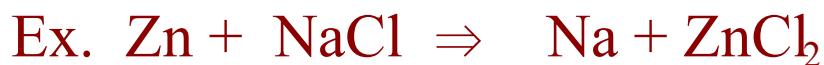
### Decomposition

one molecule ⇒ smaller particles (elements)



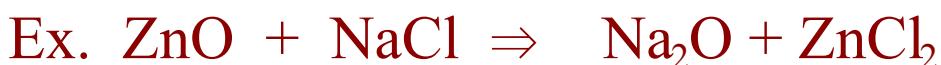
### Single Replacement

element + compound ⇒ element + compound



### Double Replacement

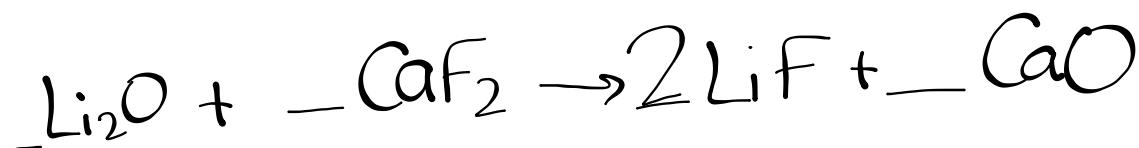
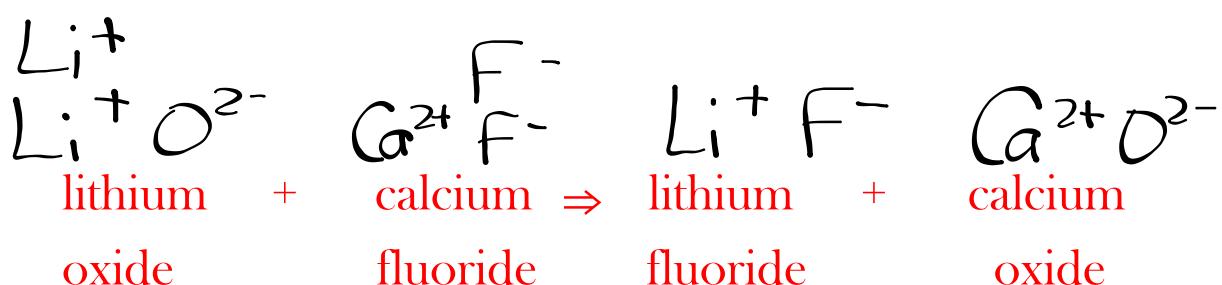
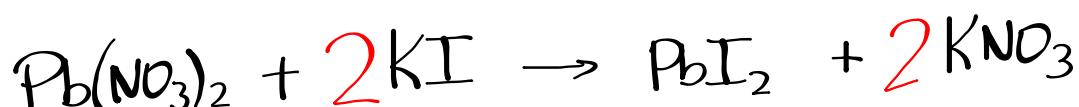
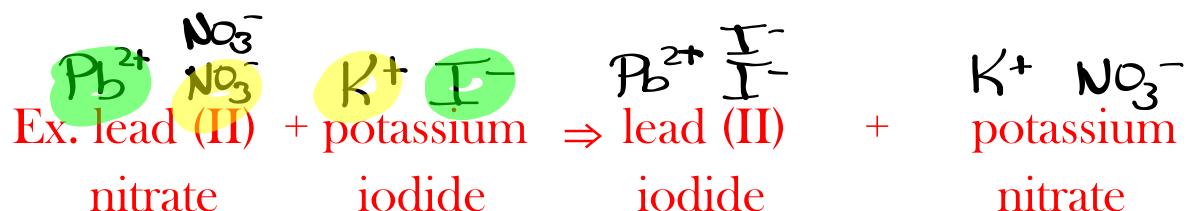
compound + compound ⇒ compound + compound



## Double Replacement Reactions

Double replacement reactions are chemical changes that involve **two compounds** as reactants.

⇒ metals (or nonmetals) will 'trade'



# **Homework**

**p. 241 #1-3**

2. a) copper + Silver nitrate  $\rightarrow$  silver + Copper(II) nitrate |
- b) zinc + hydrochloric acid  $\rightarrow$  hydrogen + Zinc chloride
- c) calcium carbonate + hydrochloric acid  $\rightarrow$  hydrogen carbonate + Calcium chloride
- d) aluminum + copper(II) fluoride  $\rightarrow$  copper + aluminum fluoride