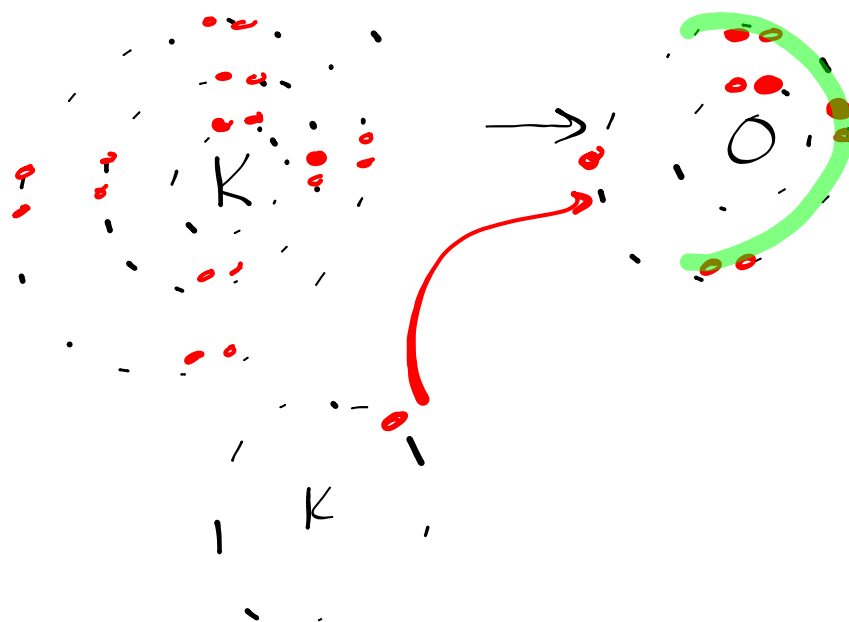


Check Homework - Worksheet



Ionic Compounds

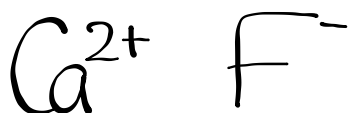
How to Write Formula and Names for Ionic Compounds

Exam
Example Calcium and Fluorine

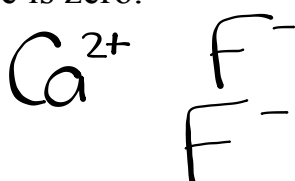
- Write the symbols, with the metal always being written first.



- Write the ionic charge above the symbol to indicate the stable ion that each element forms

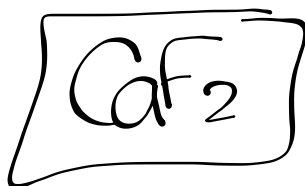


- Determine how many ions of each type you need so that the total ionic charge is zero.



$$(+2) + (-2) = 0$$

- Write the formula using subscripts to indicate the number of ions of each type.



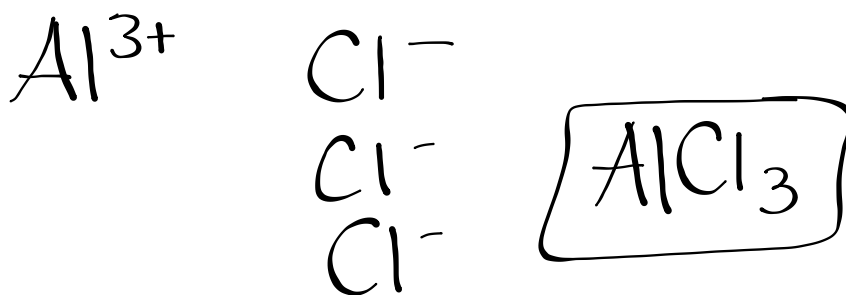
- when naming ionic compounds, the name of the metal remains the same but the name of the nonmetal changes to an -ide ending

Calcium fluoride

NOTE: If there is a common factor between the two charges, reduce!

Ex. Al^{3+} N^{3-}

Ex. aluminum and chlorine

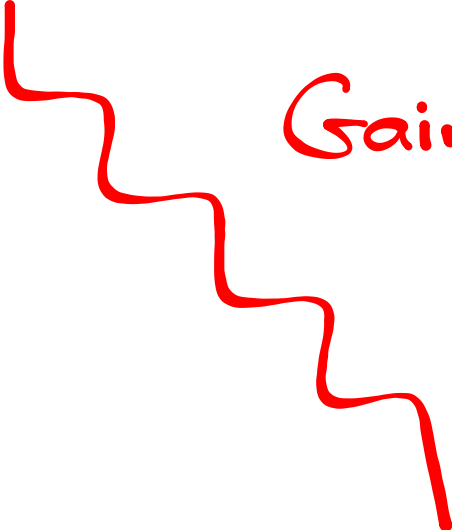


aluminum chloride

lithium and sodium



lose



Gain

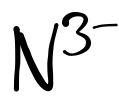
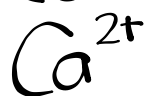
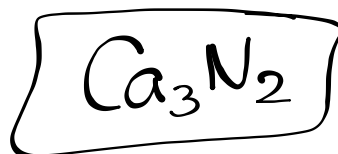
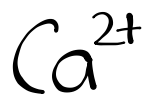
Write the formulas and the names of the compounds formed by the combination of the elements:

- a) calcium and nitrogen
- b) magnesium and oxygen

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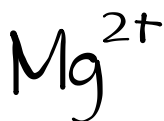
a)



Calcium nitride



b)



magnesium oxide

Homework

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