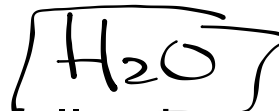
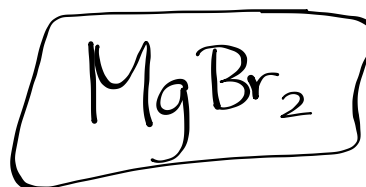
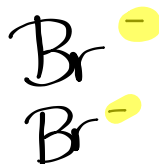


# Warm Up

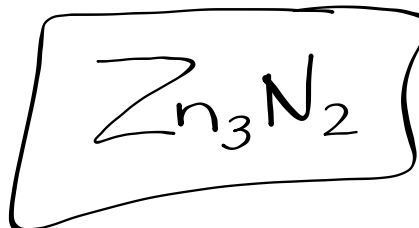
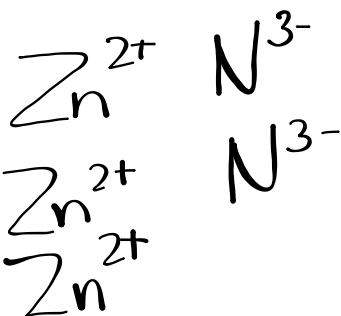


Write chemical formulas for the following ionic compounds:

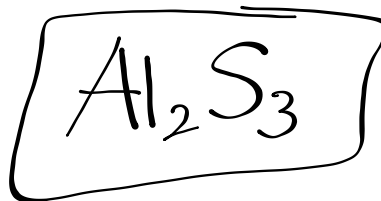
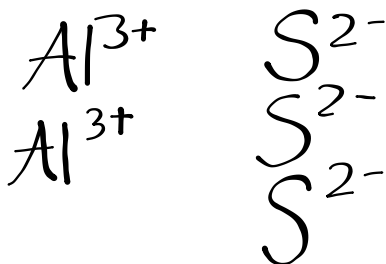
(a) magnesium bromide



(b) zinc nitride

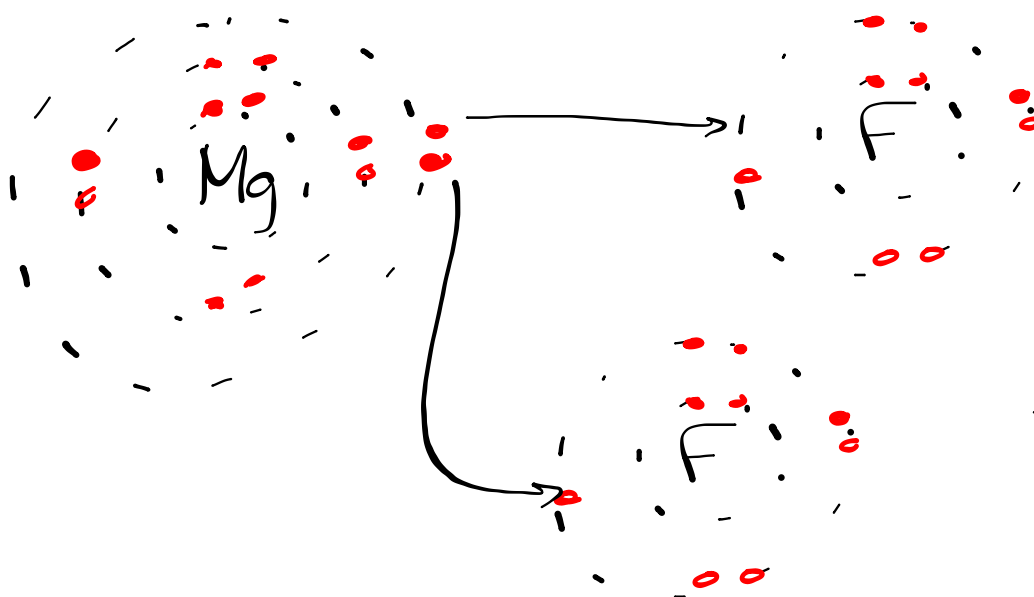


(c) aluminum sulfide



a) magnesium and bromine

## Check Homework - #2-6

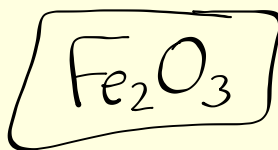
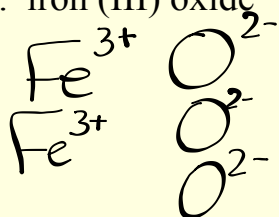


## Multivalent Metals

- some metals have more than one charge they are called **multi-valent ions**
- these elements are found in the middle block of the periodic table
- the charge that is to be used is indicated in brackets with a Roman numeral (Table 2 p. 195)



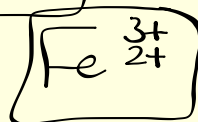
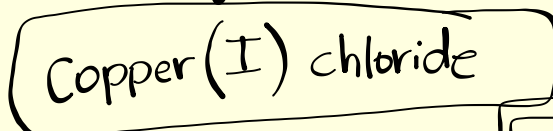
Ex. iron (III) oxide



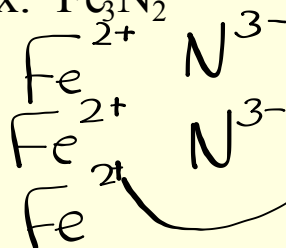
*Naming Ionic Compounds from Formula (multivalent ions):*

- Identify positive ion (metal) and negative ion (nonmetal)
- If metal is multivalent, determine its charge from the formula (balance total positives and negatives) and include in name

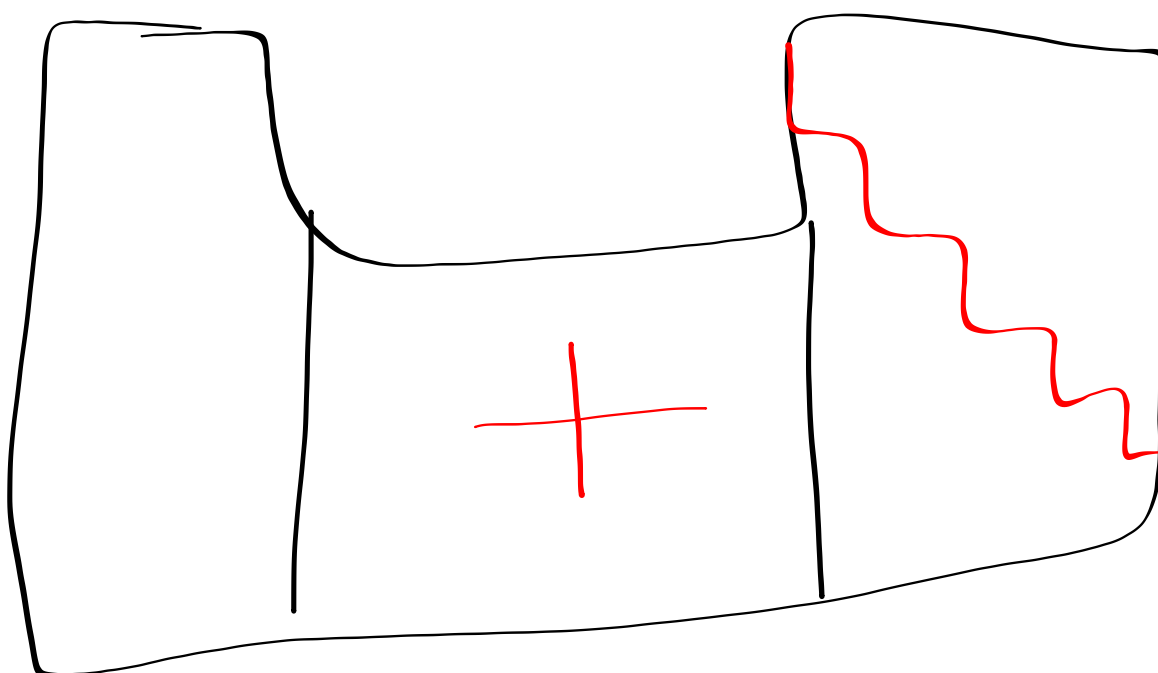
Ex. CuCl



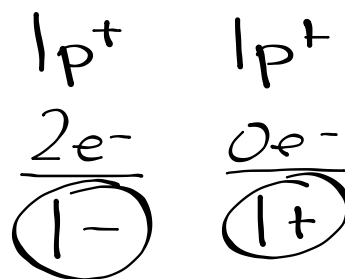
Ex. Fe<sub>3</sub>N<sub>2</sub>



iron(II)nitride



H



**Write the chemical formula of the following ionic compounds:**

**(a) iron (II) oxide**

**(b) lead (IV) chloride**

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**Write the name of the following ionic compounds:**

**(a)  $\text{Fe}_2\text{O}_3$**

**(b)  $\text{PbO}$**