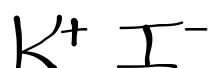


Warm Up

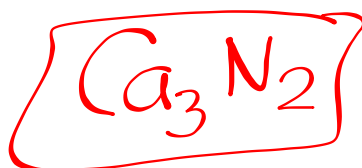
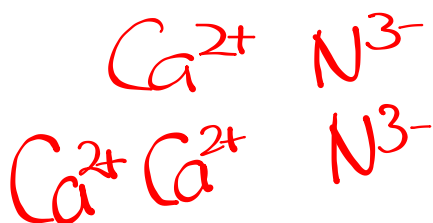
Write the chemical formula or name for the following ionic compounds:

a) KI



potassium iodide

c) calcium nitride

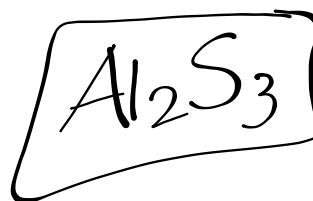
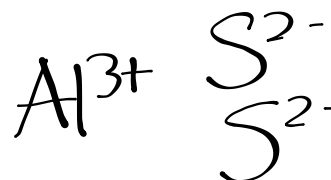
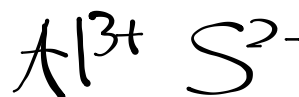


b) SrCl₂



strontium chloride

d) aluminum sulfide



Multi-Valent Metals

- can form more than one type of ion (always positive).
- include transition metals and some representative metals.

Ex. Fe^{3+} and Fe^{2+} Pb^{2+} and Pb^{4+}

In the periodic table the most common ion is usually listed in the key.

In naming multi-valent compounds (from a formula):

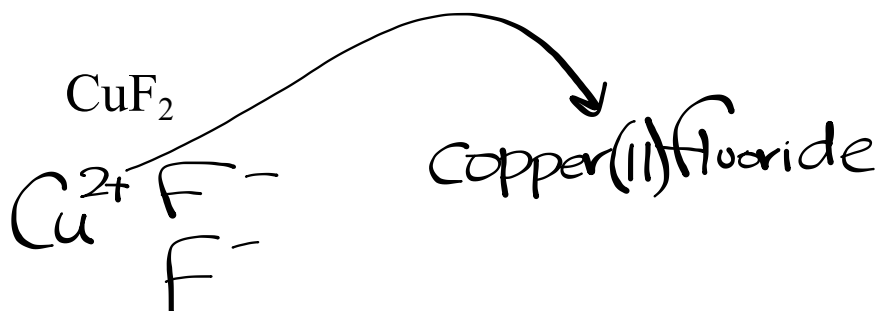
- [a] name the two ions
- [b] place the charge of the metal ion in roman numerals after the metal ion.
- [c] end the anion with an -ide suffix.

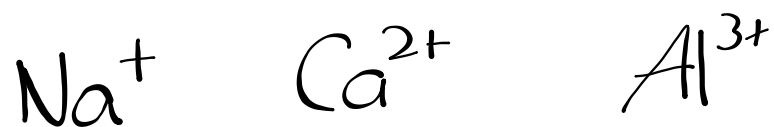
Ex. Name to Formula:

iron (III) oxide



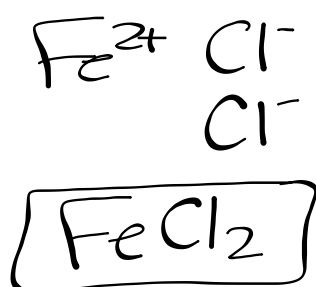
I, II, III, IV

Formula to Name :

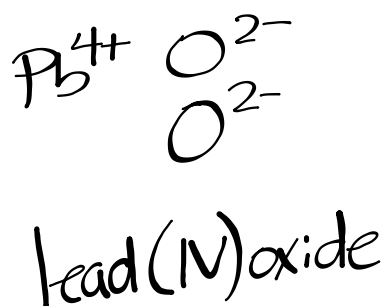


Write the chemical formula or name for the following ionic compounds:

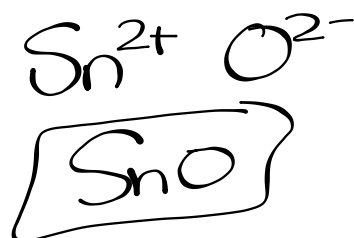
a) iron (II) chloride



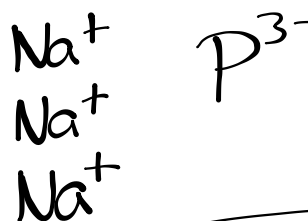
c) PbO_2



b) tin (II) oxide



d) Na_3P



Binary Ionic Compounds Type II

Worksheet #1,3

