

# Warm Up

How many **moles** are in  $2.14 \times 10^{24}$  molecules of  $\text{NO}_2$ ?

$$2.14 \times 10^{24} \text{ molecules } \text{NO}_2 \times \frac{1 \text{ mol } \text{NO}_2}{6.02 \times 10^{23} \text{ molecules } \text{NO}_2} =$$

$$3.55 \text{ mol } \text{NO}_2$$

How many **atoms** are in 6.08 moles of  $\text{C}_4\text{H}_8$ ?

$$6.08 \text{ mol } \text{C}_4\text{H}_8 \times \frac{6.02 \times 10^{23} \text{ molecules } \text{C}_4\text{H}_8}{1 \text{ mol } \text{C}_4\text{H}_8} \times \frac{12 \text{ atoms}}{1 \text{ molecule } \text{C}_4\text{H}_8}$$

$$= 4.39 \times 10^{25} \text{ atoms}$$

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(3) 4.65 mol Si

(5)  $2.75 \times 10^{24}$  atoms

(4) 0.360 mol Br<sub>2</sub>

(6) 7.72 mol NO<sub>2</sub>

$$2.17 \times 10^{23} \text{ molecules Br}_2 \times \frac{1 \text{ mol Br}_2}{6.02 \times 10^{23} \text{ molecules Br}_2} = \boxed{0.360 \text{ mol Br}_2}$$

$$1.14 \text{ mol SO}_3 \times \frac{6.02 \times 10^{23} \text{ molecules SO}_3}{1 \text{ mol SO}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule SO}_3} = \boxed{2.75 \times 10^{24} \text{ atoms}}$$

Determine the number of moles found in  $7.44 \times 10^{26}$  atoms of iron?

Determine the number of atoms found in 9.08 moles of carbon dioxide?

# Molar Mass

- the **molar mass** of a substance represents the mass of one mole of the substance
  - it is expressed in **grams per mol (g/mol)**

To determine the molar mass of a substance:

- make sure the formula is written properly
- determine the number of atoms of each element
- use the atomic molar masses of each atom from the periodic table and multiply this by the number of atoms
- add the mass of the atoms together so as to represent the total mass of the substance in grams per mole

Ex. What is the molar mass of  $(\text{NH}_4)_3\text{PO}_4$ ?

$$\text{N} \rightarrow 3 \times 14.01 = 42.03$$

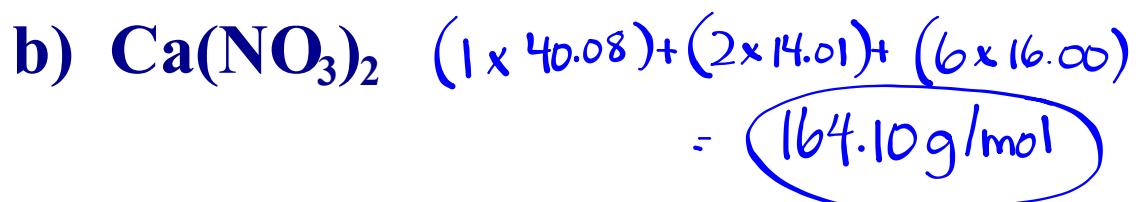
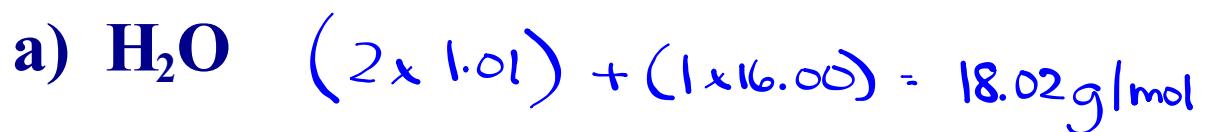
$$\text{H} \rightarrow 12 \times 1.01 = 12.12$$

$$\text{P} \rightarrow 1 \times 30.97 = 30.97$$

$$\text{O} \rightarrow 4 \times 16.00 = 64.00$$

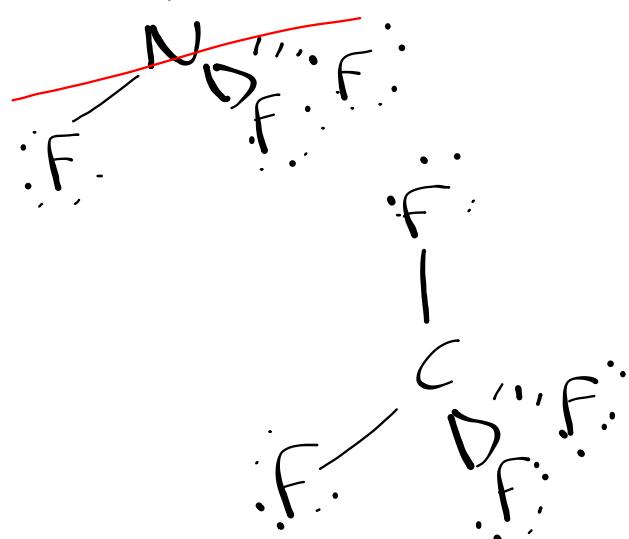
$$\boxed{149.12 \text{ g/mol}}$$

## **Find the molar mass of:**

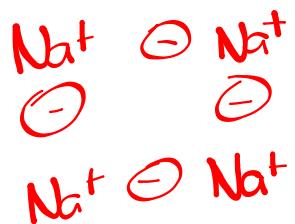


# Homework

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$$P \equiv C - \ddot{\text{F}}$$



$$C_1 \equiv C_2 = C_3$$

