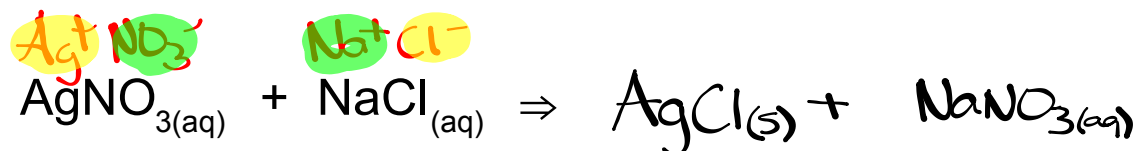
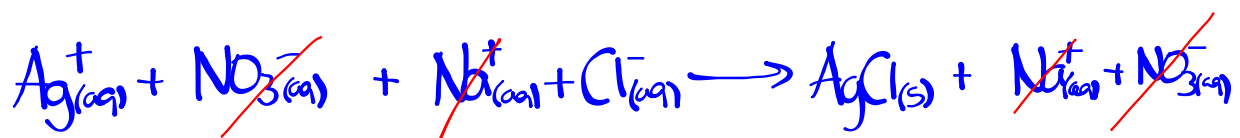


## Reactions in Aqueous Solutions



### Complete Ionic Equation

An equation that shows dissolved ionic compounds as dissociated free ions.



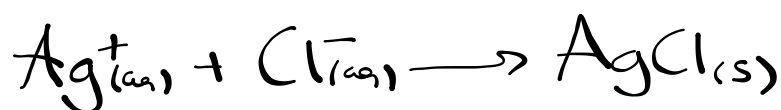
### Spectator Ion

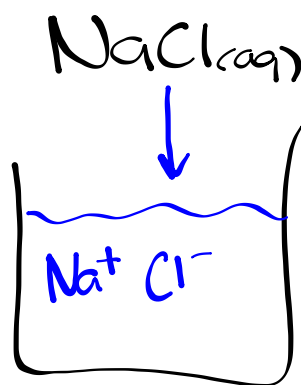
An ion that appears on both sides of the equation and is not directly involved in the reaction.



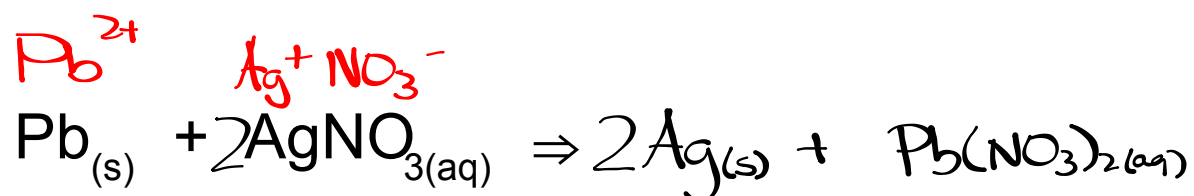
### Net Ionic Equation

An equation for a reaction in solution that only shows the particles directly involved in the reaction.

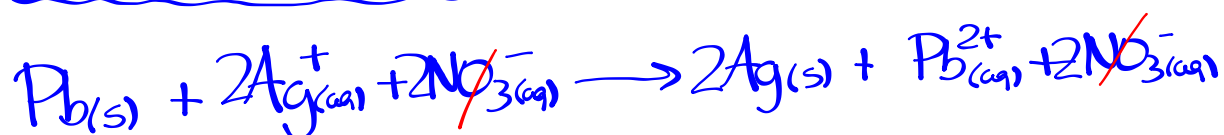




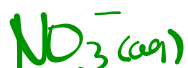
**\*All net ionic equations must be balanced with respect to both mass and charge**



COMPLETE IONIC



SPECTATOR IONS



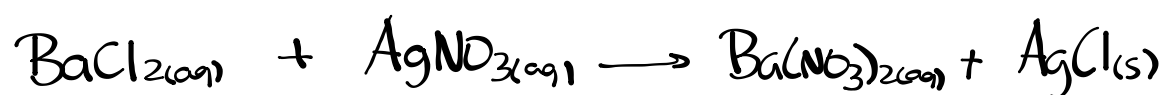
NET IONIC



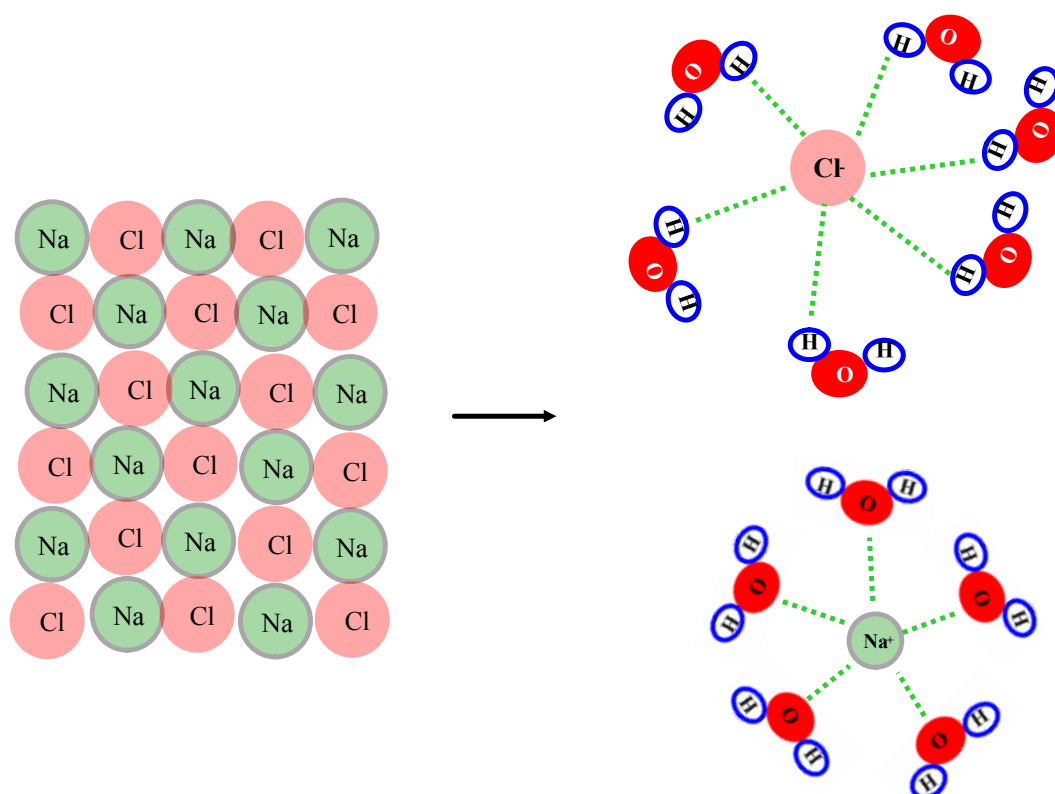
# Homework

# Worksheet

$\text{Ba}^{2+}$   $\text{Cl}^-$   $\text{Ag}^+$   $\text{NO}_3^-$   
barium chloride and silver nitrate



**What happens when an ionic compound dissolves??**



**This process is called solvation.**