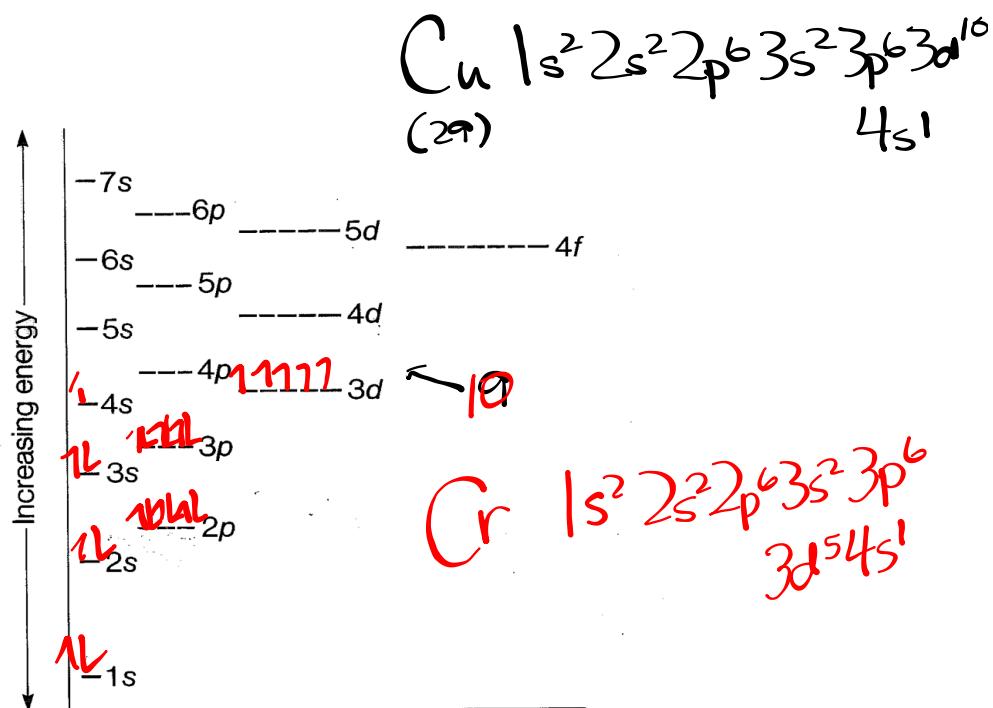


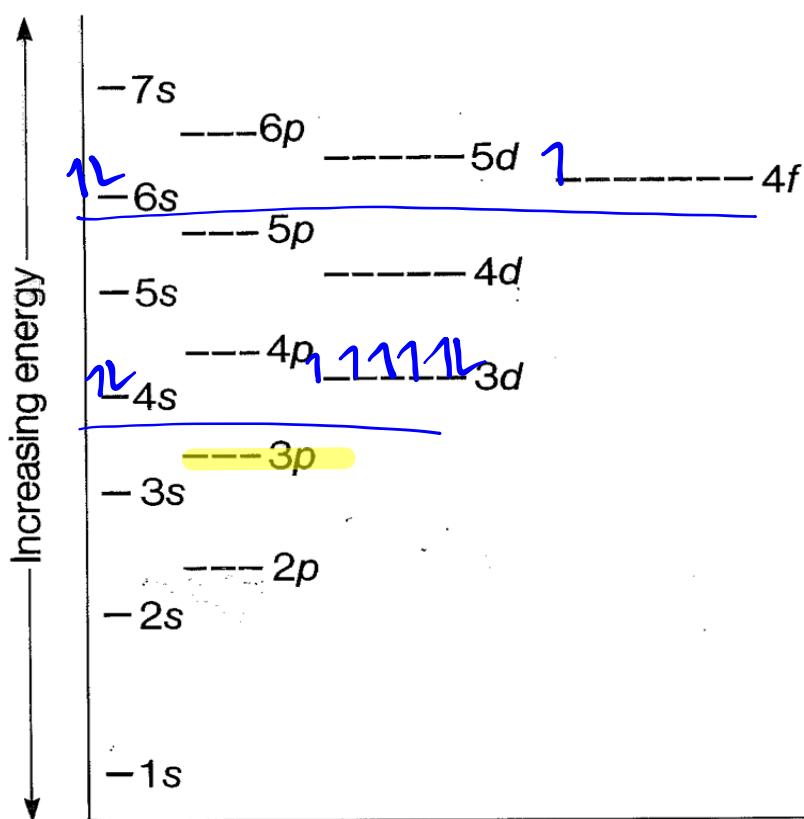
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Exceptional Electron Configurations

Although half-filled sublevels are not as stable as filled sublevels, they are more stable than other configurations.

- more likely to occur at higher principal quantum numbers, because of the small energy differences between sublevels.



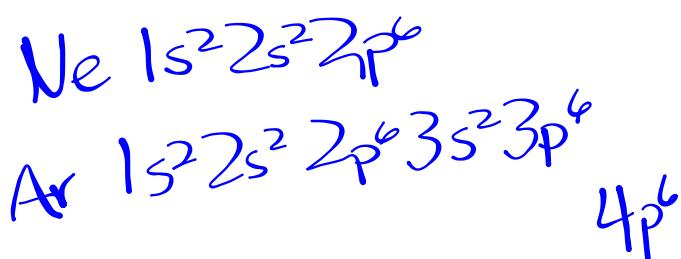
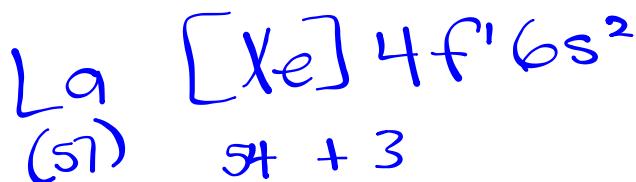
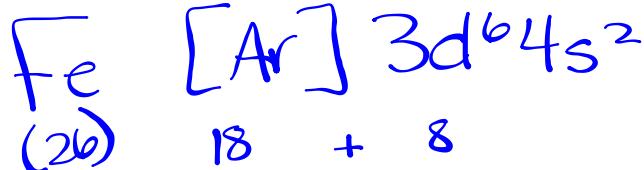


Abbreviated Electron Configurations

Handwritten annotations on the periodic table:

- 1s →** Points to Hydrogen (H) at position 1.
- 2s →** Points to Lithium (Li) at position 3.
- 3s →** Points to Sodium (Na) at position 11.
- 4s →** Points to Potassium (K) at position 19.
- 5s →** Points to Rubidium (Rb) at position 37.
- 1s²** Handwritten above the first two columns of the table.
- Na 1s² 2s² 2p⁶ 3s¹** Handwritten in blue ink.

1 H																			2 He															
3 Li	4 Be																																	
11 Na	12 Mg																																	
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr																	
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe																	
55 Cs	56 Ba	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn																		
87 Fr	88 Ra	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Uut	114 Uuo	115 Uup	116 Uuh	117 Uus	118 Uuo																		
																			57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu	
																			89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	



Homework

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Electron Configuration worksheet